

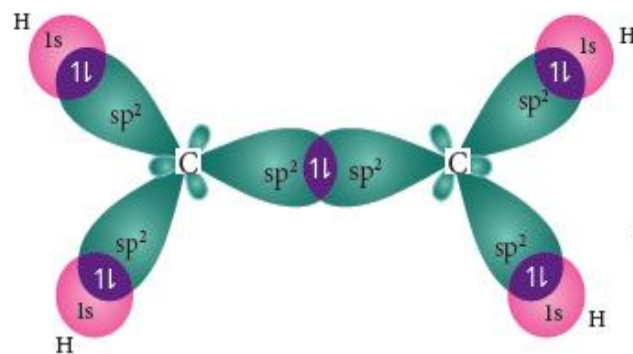
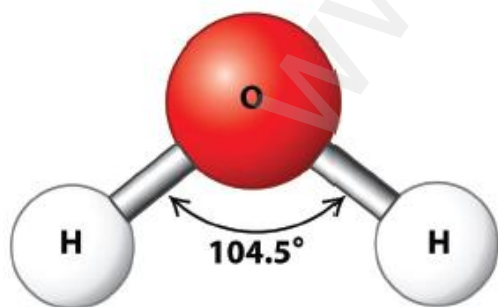
XI - Chemistry

Volume – I & II

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UNITWISE

EVALUATION and ADDITIONAL ONE MARK QUESTIONS with ANSWER KEY



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1. Basic Concepts of Chemistry and Chemical Calculations

EVALUATION:

I. Choose the correct answer:

1. 40 ml of methane is completely burnt using 80 ml of oxygen at room temperature. The volume of gas left after cooling to room temperature is
 (a) 40 ml CO₂ gas (b) 40 ml CO₂ gas and 80 ml H₂O gas
 (c) 60 ml CO₂ gas and 60 ml H₂O gas (d) 120 ml CO₂ gas
2. An element X has the following isotopic composition ²⁰⁰X = 90 %, ¹⁹⁹X = 8 % and ²⁰²X = 2 %. The weighted average atomic mass of the element X is closest to
 (a) 201 u (b) 202 u (c) 199 u (d) 200 u
3. Assertion : Two mole of glucose contains 12.044×10^{23} molecules of glucose
 Reason : Total number of entities present in one mole of any substance is equal to 6.02×10^{22}
 (a) both assertion and reason are true and the reason is the correct explanation of assertion
 (b) both assertion and reason are true but reason is not the correct explanation of assertion
 (c) assertion is true but reason is false
 (d) both assertion and reason are false
4. Carbon forms two oxides, namely carbon monoxide and carbon dioxide. The equivalent mass of which element remains constant?
 (a) Carbon (b) oxygen (c) both carbon and oxygen (d) neither carbon nor oxygen
5. The equivalent mass of a trivalent metal element is 9 g eq⁻¹ the molar mass of its anhydrous oxide is
 (a) 102 g (b) 27 g (c) 270 g (d) 78 g
6. The number of water molecules in a drop of water weighing 0.018 g is
 (a) 6.022×10^{26} (b) 6.022×10^{23} (c) 6.022×10^{20} (d) 9.9×10^{22}
7. 1 g of an impure sample of magnesium carbonate (containing no thermally decomposable impurities) on complete thermal decomposition gave 0.44 g of carbon dioxide gas. The percentage of impurity in the sample is
 (a) 0 % (b) 4.4 % (c) 16 % (d) 8.4 %
8. When 6.3 g of sodium bicarbonate is added to 30 g of acetic acid solution, the residual solution is found to weigh 33 g. The number of moles of carbon dioxide released in the reaction is
 (a) 3 (b) 0.75 (c) 0.075 (d) 0.3
9. When 22.4 litres of H₂ (g) is mixed with 11.2 litres of Cl₂ (g), each at 273 K at 1 atm the moles of HCl (g), formed is equal to
 (a) 2 moles of HCl (g) (b) 0.5 moles of HCl (g)
 (c) 1.5 moles of HCl (g) (d) 1 moles of HCl (g)
10. Hot concentrated sulphuric acid is a moderately strong oxidising agent. Which of the following reactions does not show oxidising behaviour?
 (a) $\text{Cu} + 2\text{H}_2\text{SO}_4 \rightarrow \text{CuSO}_4 + \text{SO}_2 + 2\text{H}_2\text{O}$ (b) $\text{C} + 2\text{H}_2\text{SO}_4 \rightarrow \text{CO}_2 + 2\text{SO}_2 + 2\text{H}_2\text{O}$
 (c) $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{HCl}$ (d) none of the above
11. Choose the disproportionation reaction among the following redox reactions.
 (a) $3\text{Mg}(\text{s}) + \text{N}_2(\text{g}) \rightarrow \text{Mg}_3\text{N}_2(\text{s})$ (b) $\text{P}_4(\text{s}) + 3\text{NaOH} + 3\text{H}_2\text{O} \rightarrow \text{PH}_3(\text{g}) + 3\text{NaH}_2\text{PO}_2(\text{aq})$
 (c) $\text{Cl}_2(\text{g}) + 2\text{KI}(\text{aq}) \rightarrow 2\text{KCl}(\text{aq}) + \text{I}_2$ (d) $\text{Cr}_2\text{O}_3(\text{s}) + 2\text{Al}(\text{s}) \rightarrow \text{Al}_2\text{O}_3(\text{s}) + 2\text{Cr}(\text{s})$
12. The equivalent mass of potassium permanganate in alkaline medium is $\text{MnO}_4^- + 2\text{H}_2\text{O} + 3\text{e}^- \rightarrow \text{MnO}_2 + 4\text{OH}^-$
 (a) 31.6 (b) 52.7 (c) 79 (d) None of these

13. Which one of the following represents 180g of water?
 (a) 5 Moles of water (b) 90 moles of water
 (c) $(6.022 \times 10^{23}) / 180$ molecules of water (d) 6.022×10^{24} molecules of water
14. 7.5 g of a gas occupies a volume of 5.6 litres at 0°C and 1 atm pressure. The gas is
 (a) NO (b) N_2O (c) CO (d) CO_2
15. Total number of electrons present in 1.7 g of ammonia is
 (a) 6.022×10^{23} (b) $\frac{6.022 \times 10^{22}}{1.7}$
 (c) $\frac{6.022 \times 10^{24}}{1.7}$ (d) $\frac{6.022 \times 10^{23}}{1.7}$
16. The correct increasing order of the oxidation state of sulphur in the anions
 SO_4^{2-} , SO_3^{2-} , $\text{S}_2\text{O}_4^{2-}$, $\text{S}_2\text{O}_6^{2-}$ is
 (a) $\text{SO}_3^{2-} < \text{SO}_4^{2-} < \text{S}_2\text{O}_4^{2-} < \text{S}_2\text{O}_6^{2-}$ (b) $\text{SO}_4^{2-} < \text{S}_2\text{O}_4^{2-} < \text{S}_2\text{O}_6^{2-} < \text{SO}_3^{2-}$
 (c) $\text{S}_2\text{O}_4^{2-} < \text{SO}_3^{2-} < \text{S}_2\text{O}_6^{2-} < \text{SO}_4^{2-}$ (d) $\text{S}_2\text{O}_6^{2-} < \text{S}_2\text{O}_4^{2-} < \text{SO}_4^{2-} < \text{SO}_3^{2-}$
17. The equivalent mass of ferrous oxalate is
 a) Molar mass of ferrous oxalate / 1 b) Molar mass of ferrous oxalate / 2
 c) Molar mass of ferrous oxalate / 3 d) None of these
18. If Avogadro number were changed from 6.022×10^{23} to 6.022×10^{20} , this would change
 (a) the ratio of chemical species to each other in a balanced equation
 (b) the ratio of elements to each other in a compound
 (c) the definition of mass in units of grams
 (d) the mass of one mole of carbon
19. Two 22.4 litre containers A and B contains 8 g of O_2 and 8 g of SO_2 respectively at 273 K and 1 atm pressure, then
 (a) Number of molecules in A and B are same
 (b) Number of molecules in B is more than that in A.
 (c) The ratio between the number of molecules in A to number of molecules in B is 2:1
 (d) Number of molecules in B is three times greater than the number of molecules in A.
20. What is the mass of precipitate formed when 50 ml of 8.5 % solution of AgNO_3 is mixed with 100 ml of 1.865 % potassium chloride solution? (a) 3.59 g (b) 7 g (c) 14 g (d) 28 g
21. The mass of a gas that occupies a volume of 612.5 ml at room temperature and pressure (25°C and 1 atm pressure) is 1.1g. The molar mass of the gas is
 (a) 66.25 g mol^{-1} (b) 44 g mol^{-1} (c) 24.5 g mol^{-1} (d) 662.5 g mol^{-1}
22. Which of the following contain same number of carbon atoms as in 6 g of carbon-12.
 (a) 7.5 g ethane (b) 8 g methane (c) both (a) and (b) (d) none of these
23. Which of the following compound(s) has /have percentage of carbon same as that in ethylene (C_2H_4)
 (a) propene (b) ethyne (c) benzene (d) ethane
24. Which of the following is/are true with respect to carbon -12.
 (a) relative atomic mass is 12 u (b) oxidation number of carbon is +4 in all its compounds.
 (c) 1 mole of carbon-12 contain 6.022×10^{22} carbon atoms (d) all of these
25. Which one of the following is used as a standard for atomic mass.
 (a) ${}^{12}_6\text{C}$ (b) ${}^{12}_7\text{C}$ (c) ${}^{13}_6\text{C}$ (d) ${}^{14}_6\text{C}$

ADDITIONAL :

1. One mole of CO_2 contains (a) 6.023×10^{23} atoms of C (b) 6.023×10^{23} atoms of O
 (c) 18.1×10^{23} molecules of CO_2 (d) 3g atoms of CO_2
2. The number of moles of H_2 in 0.224 liter of hydrogen gas at STP is (a) 1 (b) 0.1 (c) 0.01 (d) 0.001

3. The number of molecules in 16 g of methane is(a) 3.023×10^{23} (b) 6.023×10^{23}
(c) $16 / 6.023 \times 10^{23}$ (d) $6.023 / 3 \times 10^{23}$
4. The equivalent mass of KMnO_4 when it is converted to MnSO_4 is equal to molar mass divided by
(a) 6 (b) 4 (c) 5 (d) 2
5. The empirical formula of hydrogen peroxide is (a) HO (b) H_2O (c) H_3O (d) H_2O_2
6. Molecular mass = (a) Vapour Density $\times 2$ (b) Vapour Density $\div 2$ (c) Vapour Density $\times 3$ (d) Vapour Density
7. When 22 L of hydrogen gas is mixed with 11.2 L of chlorine gas, each at STP, the moles of HCl gas formed is equal to(a) 2 (b) 0.5 (c) 1.5 (d) 1
8. 5.6 L of a gas at STP are found to have mass of 11 g. The molecular mass of the gas is
(a) 36 (b) 48 (c) 40 (d) 44
9. Oxidation number of Fluorine in all compounds is(a) +1 (b) -1 (c) 0 (d) -2
10. The mass of one mole of CaCl_2 is(a) 55.5 g mol^{-1} (b) 111 g mol^{-1} (c) 222 g mol^{-1} (d) 77.5 g mol^{-1}
11. The formula weight of ethanol ($\text{C}_2\text{H}_5\text{OH}$) is(a) 56.5 amu (b) 16 amu (c) 60 amu (d) 46 amu
12. The equivalent mass of Aluminium is(a) 27 (b) 13.5 (c) 54 (d) 9
13. How many molecules of hydrogen is required to produce 4 moles of ammonia?
(a) 15 moles (b) 20 moles (c) 6 moles (d) 4 moles
14. The number of moles of oxygen required to prepare 1 mole of water is
(a) 1 mole (b) 0.5 mole (c) 2 moles (d) 0.4 mole
15. How much volume of CO_2 is produced when 50 g of CaCO_3 is heated strongly?
(a) $2.24 \times 10^{-2} \text{ m}^3$ (b) 22.4 (c) 11.2 L (d) 22400 cm^3
16. Identify the compound formed during the rusting of iron.
(a) Fe_2O_3 (b) $\text{Fe}_2\text{O}_3 \cdot x \text{H}_2\text{O}$ (c) $\text{FeO} \cdot x \text{H}_2\text{O}$ (d) FeO
17. The oxidation number of Cr in $\text{K}_2\text{Cr}_2\text{O}_7$ is (a) +4 (b) +6 (c) 0 (d) +7
18. The oxidation number of N in NH_4^+ ion is(a) +4 (b) +3 (c) -3 (d) -4
19. $\text{Zn}_{(s)} + \text{Cu}^{2+}_{(aq)} \rightarrow \text{Zn}^{2+}_{(aq)} + \text{Cu}_{(s)}$. In this reaction, which gets oxidised?
(a) Cu^{2+} (b) Zn^{2+} (c) Zn (d) Zn, Cu^{2+}
20. Which one of the following is an example of disproportionation reaction?
(a) $\text{CuSO}_4 + \text{Zn} \rightarrow \text{ZnSO}_4 + \text{Cu}$ (b) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
(c) $\text{PCl}_5 \rightarrow \text{PCl}_3 + \text{Cl}_2$ (d) $4\text{H}_3\text{PO}_3 \rightarrow 3\text{H}_3\text{PO}_4 + \text{PH}_3$
21. The molar mass of Na_2SO_4 is(a) 129 (b) 142 (c) 110 (d) 70
22. The oxidation number of Fe in $\text{Fe}_2(\text{SO}_4)_3$ is(a) +2 (b) +3 (c) +2, +3 (d) 0
23. Among the following molecules in which Chlorine shows maximum oxidation state?
(a) Cl_2 (b) KCl (c) KClO_3 (d) Cl_2O_7

2. Quantum Mechanical Model of Atom

1. Electronic configuration of species M^{2+} is $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6$ and its atomic weight is 56. The number of neutrons in the nucleus of species M is (a) 26 (b) 22 (c) 30 (d) 24
2. The energy of light of wavelength 45 nm is (a) $6.67 \times 10^{15} \text{ J}$ (b) $6.67 \times 10^{11} \text{ J}$ (c) $4.42 \times 10^{-18} \text{ J}$ (d) $4.42 \times 10^{-15} \text{ J}$
3. The energies E_1 and E_2 of two radiations are 25 eV and 50 eV respectively. The relation between their wavelengths λ_1 and λ_2 will be
(a) $\frac{\lambda_1}{\lambda_2} = 1$ (b) $\lambda_1 = 2\lambda_2$ (c) $\lambda_1 = \sqrt{25 \times 50} \lambda_2$ (d) $2\lambda_1 = \lambda_2$
4. Splitting of spectral lines in an electric field is called
(a) Zeeman effect (b) Shielding effect (c) Compton effect (d) Stark effect
5. Based on equation $E = -2.178 \times 10^{-18} \text{ J } z^2/n^2$, certain conclusions are written. Which of them is not correct?
(a) Equation can be used to calculate the change in energy when the electron changes orbit

- b) For $n = 1$, the electron has a more negative energy than it does for $n = 6$ which means that the electron is more loosely bound in the smallest allowed orbit
- c) The negative sign in equation simply means that the energy of electron bound to the nucleus is lower than it would be if the electrons were at the infinite distance from the nucleus.
- d) Larger the value of n , the larger is the orbit radius.

6. According to the Bohr Theory, which of the following transitions in the hydrogen atom will give rise to the least energetic photon ? a) $n = 6$ to $n = 1$ b) $n = 5$ to $n = 4$ c) $n = 5$ to $n = 3$ d) $n = 6$ to $n = 5$

7. Assertion : The spectrum of He^+ is expected to be similar to that of hydrogen

Reason : He^+ is also one electron system.

- (a) If both assertion and reason are true and reason is the correct explanation of assertion.
- (b) If both assertion and reason are true but reason is not the correct explanation of assertion.
- (c) If assertion is true but reason is false
- (d) If both assertion and reason are false

8. Which of the following pairs of d-orbitals will have electron density along the axes ?

- a) dz^2, dxz b) dxz, dyz c) dz^2, dx^2-y^2 d) dxy, dx^2-y^2

9. Two electrons occupying the same orbital are distinguished by

- a) azimuthal quantum number b) spin quantum number
- c) magnetic quantum number d) orbital quantum number

10. The electronic configuration of Eu (Atomic no. 63) Gd (Atomic no. 64) and Tb (Atomic no. 65) are

- a) $[\text{Xe}] 4f^6 5d^1 6s^2$, $[\text{Xe}] 4f^7 5d^1 6s^2$ and $[\text{Xe}] 4f^8 5d^1 6s^2$
- b) $[\text{Xe}] 4f^7, 6s^2$, $[\text{Xe}] 4f^7 5d^1 6s^2$ and $[\text{Xe}] 4f^9 6s^2$
- c) $[\text{Xe}] 4f^7, 6s^2$, $[\text{Xe}] 4f^8 6s^2$ and $[\text{Xe}] 4f^8 5d^1 6s^2$
- d) $[\text{Xe}] 4f^6 5d^1 6s^2$, $[\text{Xe}] 4f^7 5d^1 6s^2$ and $[\text{Xe}] 4f^9 6s^2$

11. The maximum number of electrons in a sub shell is given by the expression

- a) $2n^2$ b) $2l + 1$ c) $4l + 2$ d) none of these

12. For d-electron, the orbital angular momentum is

- a) $\frac{\sqrt{2}h}{2\pi}$ b) $\frac{\sqrt{2}h}{2\pi}$ c) $\frac{\sqrt{2 \times 4} h}{2\pi}$ d) $\frac{\sqrt{6} h}{2\pi}$

13. What is the maximum numbers of electrons that can be associated with the following set of quantum numbers ? $n = 3, l = 1$ and $m = -1$ a) 4 b) 6 c) 2 d) = 10

14. Assertion : Number of radial and angular nodes for 3p orbital are 1, 1 respectively.

Reason : Number of radial and angular nodes depends only on principal quantum number.

- (a) both assertion and reason are true and reason is the correct explanation of assertion.
- (b) both assertion and reason are true but reason is not the correct explanation of assertion.
- (c) assertion is true but reason is false
- (d) both assertion and reason are false

15. The total number of orbitals associated with the principal quantum number $n = 3$ is a) 9 b) 8 c) 5 d) 7

16. If $n = 6$, the correct sequence for filling of electrons will be,

- a) $ns \rightarrow (n - 2) f \rightarrow (n - 1) d \rightarrow np$ b) $ns \rightarrow (n - 1) d \rightarrow (n - 2) f \rightarrow np$
- c) $ns \rightarrow (n - 2) f \rightarrow np \rightarrow (n - 1) d$ d) none of these are correct

17. Consider the following sets of quantum numbers :
- | | n | l | m | s |
|-------|---|---|----|----------------|
| (i) | 3 | 0 | 0 | $+\frac{1}{2}$ |
| (ii) | 2 | 2 | 1 | $-\frac{1}{2}$ |
| (iii) | 4 | 3 | -2 | $+\frac{1}{2}$ |
| (iv) | 1 | 0 | -1 | $+\frac{1}{2}$ |
| (v) | 3 | 4 | 3 | $-\frac{1}{2}$ |
- Which of the following sets of quantum number is not possible ?
 a) (i), (ii), (iii) and (iv) b) (ii), (iv) and (v)
 c) (i) and (iii) d) (ii), (iii) and (iv)
18. How many electrons in an atom with atomic number 105 can have $(n + l) = 8$?
 a) 30 b) 17 c) 15 d) unpredictable
19. Electron density in the yz plane of 3dxy orbital is
 a) zero b) 0.50 c) 0.75 d) 0.90
20. If uncertainty in position and momentum are equal, then minimum uncertainty in velocity is
 a) $\frac{1}{m} \sqrt{\frac{h}{\pi}}$ b) $\sqrt{\frac{h}{\pi}}$ c) $\frac{1}{2m} \sqrt{\frac{h}{\pi}}$ d) $\frac{h}{4\pi}$
21. A macroscopic particle of mass 100 g and moving at a velocity of 100 cm s⁻¹ will have a de Broglie wavelength of
 a) 6.6×10^{-29} cm b) 6.6×10^{-30} cm c) 6.6×10^{-31} cm d) 6.6×10^{-32} cm
22. The ratio of de Broglie wavelengths of a deuterium atom to that of an α - particle, when the velocity of the former is five times greater than that of later, is
 a) 4 b) 0.2 c) 2.5 d) 0.4
23. The energy of an electron in the 3rd orbit of hydrogen atom is -E. The energy of an electron in the first orbit will be
 a) -3E b) -E3 c) -E9 d) -9E
24. Time independent Schrodinger wave equation is
 a) $\hat{H}\psi = E\psi$ b) $\nabla^2\psi + \frac{8\pi^2m}{h^2}(E+V)\psi = 0$
 c) $\frac{\partial^2\psi}{\partial x^2} + \frac{\partial^2\psi}{\partial y^2} + \frac{\partial^2\psi}{\partial z^2} + \frac{2m}{h^2}(E-V)\psi = 0$ d) all of these
25. Which of the following does not represent the mathematical expression for the Heisenberg uncertainty principle ?
 a) $\Delta x \cdot \Delta p \geq \frac{h}{4\pi}$ b) $\Delta x \cdot \Delta v \geq \frac{h}{4\pi m}$
 c) $\Delta E \cdot \Delta t \geq \frac{h}{4\pi}$ d) $\Delta E \cdot \Delta x \geq \frac{h}{4\pi}$

Additional :

1. The energy of an electron of hydrogen atom in 2nd main shell is equal to
 (a) - 13.6 eV atom⁻¹ (b) - 6.8 eV atom⁻¹ (c) - 0.34 eV atom⁻¹ (d) - 3.4 eV atom⁻¹
2. The energy of an electron of Li²⁺ in the 3rd main shell is
 (a) - 1.51 eV atom⁻¹ (b) - 6.8 eV atom⁻¹ (c) + 1.51 eV atom⁻¹ (d) - 3.4 eV atom⁻¹
3. The maximum number of electrons that can be accommodated in N shell is
 (a) 8 (b) 18 (c) 32 (d) 36
4. When $l = 0$, the number of electrons that can be accommodated in the subshell is
 (a) 0 (b) 2 (c) 6 (d) 8
5. Which of the following provides the experimental justification of magnetic quantum number?
 (a) Zeeman effect (b) Stark effect (c) Uncertainty principle (d) Quantum condition
6. What are the values of n, l, m, and s for 3p_x electron?
 (a) 3, 2, 1, 0 (b) 3, 1, -1, $+\frac{1}{2}$ (c) 3, 2, +1, $-\frac{1}{2}$ (d) 3, 0, 0, $+\frac{1}{2}$
7. The region where the probability density function of electron reduces to zero is called
 (a) orbit (b) orbital (c) nodal surface (d) subshell

8. Which one of the following is the correct increasing order of effective nuclear charge felt by an electron?
 (a) $s > p > d > f$ (b) $s < p < d < f$ (c) $s > p > f > d$ (d) $f < p < d < s$
9. The value of n , l , m and s of 8th electron in an oxygen atom are respectively
 (a) 1, 0, 0, $+\frac{1}{2}$ (b) 2, 1, $+1$, $-\frac{1}{2}$ (c) 2, 1, -1 , $-\frac{1}{2}$ (d) 2, 1, 0, $+\frac{1}{2}$
10. Which of the following is the actual configuration of Cr ($Z = 24$)?
 (a) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4 4s^2$ (b) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1$
 (c) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6$ (d) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^3$
11. The orbital with $n = 3$ and $l = 2$ is
 (a) 3s (b) 3p (c) 3d (d) 3f

3. Periodic Classification of Elements

EVALUATION:

I. Choose the correct answer:

1. What would be the IUPAC name for an element with atomic number 222?
 a) bibibium b) bididium c) didibium d) bibibium
2. The electronic configuration of the elements A and B are $1s^2, 2s^2, 2p^6, 3s^2$ and $1s^2, 2s^2, 2p^5$ respectively. The formula of the ionic compound that can be formed between these elements is
 a) AB b) AB_2 c) A_2B d) none of the above.
3. The group of elements in which the differentiating electron enters the anti penultimate shell of atoms are called
 a) p-block elements b) d-block elements c) s-block elements d) f-block elements
4. In which of the following options the order of arrangement does not agree with the variation of property indicated against it?
 a) $I < Br < Cl < F$ (increasing electron gain enthalpy) b) $Li < Na < K < Rb$ (increasing metallic radius)
 c) $Al^{3+} < Mg^{2+} < Na^+ < F^-$ (increasing ionic size) d) $B < C < O < N$ (increasing first ionisation enthalpy)
5. Which of the following elements will have the highest electronegativity?
 a) Chlorine b) Nitrogen c) Cesium d) Fluorine
6. Various successive ionisation enthalpies (in kJ mol^{-1}) of an element are given below.

The element is
 a) phosphorus b) Sodium
 c) Aluminium d) Silicon

| IE_1 | IE_2 | IE_3 | IE_4 | IE_5 |
|--------|--------|--------|--------|--------|
| 577.5 | 1,810 | 2,750 | 11,580 | 14,820 |

7. In the third period the first ionization potential is of the order.
 a) $Na > Al > Mg > Si > P$ b) $Na < Al < Mg < Si < P$
 c) $Mg > Na > Si > P > Al$ d) $Na < Al < Mg < P < Si$
8. Identify the wrong statement.
 a) Amongst the isoelectronic species, smaller the positive charge on cation, smaller is the ionic radius
 b) Amongst isoelectronic species greater the negative charge on the anion, larger is the ionic radius
 c) Atomic radius of the elements increases as one moves down the first group of the periodic table
 d) Atomic radius of the elements decreases as one moves across from left to right in the 2nd period of the periodic table.
9. Which one of the following arrangements represent the correct order of least negative to most negative electron gain enthalpy
 a) $Al < O < C < Ca < F$ b) $Al < Ca < O < C < F$ c) $C < F < O < Al < Ca$ d) $Ca < Al < C < O < F$
10. The correct order of electron gain enthalpy with negative sign of F, Cl, Br and I having atomic number 9, 17, 35 and 53 respectively is
 a) $I > Br > Cl > F$ b) $F > Cl > Br > I$ c) $Cl > F > Br > I$ d) $Br > I > Cl > F$

11. Which one of the following is the least electronegative element?
 a) Bromine b) Chlorine c) Iodine d) Hydrogen
12. The element with positive electron gain enthalpy is
 a) Hydrogen b) Sodium c) Argon d) Fluorine
13. The correct order of decreasing electronegativity values among the elements X, Y, Z and A with atomic numbers 4, 8, 7 and 12 respectively
 a) $Y > Z > X > A$ b) $Z > A > Y > X$ c) $X > Y > Z > A$ d) $X > Y > A > Z$
14. Assertion: Helium has the highest value of ionisation energy among all the elements known
 Reason: Helium has the highest value of electron affinity among all the elements known
 a) Both assertion and reason are true and reason is correct explanation for the assertion
 b) Both assertion and reason are true but the reason is not the correct explanation for the assertion
 c) Assertion is true and the reason is false
 d) Both assertion and the reason are false
15. The electronic configuration of the atom having maximum difference in first and second ionisation energies is
 a) $1s^2, 2s^2, 2p^6, 3s^1$ b) $1s^2, 2s^2, 2p^6, 3s^2$
 c) $1s^2, 2s^2, 2p^6, 3s^2, 3s^2, 3p^6, 4s^1$ d) $1s^2, 2s^2, 2p^6, 3s^2, 3p^1$
16. Which of the following is second most electronegative element?
 a) Chlorine b) Fluorine c) Oxygen d) Sulphur
17. IE_1 and IE_2 of Mg are 179 and 348 kcal mol⁻¹ respectively. The energy required for the reaction
 $Mg \rightarrow Mg^{2+} + 2e^-$ is a) +169 kcal mol⁻¹ b) -169 kcal mol⁻¹ c) +527 kcal mol⁻¹ d) -527 kcal mol⁻¹
18. In a given shell the order of screening effect is
 a) $s > p > d > f$ b) $s > p > f > d$ c) $f > d > p > s$ d) $f > p > s > d$
19. Which of the following orders of ionic radii is correct?
 a) $H^- > H^+ > H$ b) $Na^+ > F^- > O^{2-}$ c) $F > O^{2-} > Na^+$ d) None of these
20. The First ionisation potential of Na, Mg and Si are 496, 737 and 786 kJ mol⁻¹ respectively. The ionisation potential of Al will be closer to
 a) 760 kJ mol⁻¹ b) 575 kJ mol⁻¹ c) 801 kJ mol⁻¹ d) 419 kJ mol⁻¹
21. Which one of the following is true about metallic character when we move from left to right in a period and top to bottom in a group?
 a) Decreases in a period and increases along the group b) Increases in a period and decreases in a group
 c) Increases both in the period and the group d) Decreases both in the period and in the group
22. How does electron affinity change when we move from left to right in a period in the periodic table?
 a) Generally increases b) Generally decreases
 c) Remains unchanged d) First increases and then decreases
23. Which of the following pairs of elements exhibit diagonal relationship?
 a) Be and Mg b) Li and Be c) Be and B d) Be and Al

ADDITIONAL :

1. The law of triads is not obeyed by (a) Ca, Sr, Ba (b) Cl, Br, I (c) Li, Na, K (d) Be, B, C
2. Which of the following elements were unknown at that time of Mendeleev?
 (a) Na, Mg (b) Fe, Co (c) K, Cu (d) Ga, Ge
3. According to modern periodic law, the physical and chemical properties of the elements are periodic functions of their (a) atomic volume (b) atomic numbers (c) atomic weights (d) valency
4. Which one of the following is called halogen family?
 (a) Group 17 (b) Group 16 (c) Group 1 (d) Group 2

5. Which one of the following is a metalloid? (a) N (b) P (c) Bi (d) Sb
6. The general electronic configuration of d-block elements is
(a) $ns^2 nd^{1-10}$ (b) $(n-1)d^{1-10} ns^{0-2}$ (c) $(n-2)d^{1-10} (n-1)^{0-2}$ (d) $ns^2 nd^5$
7. All the s-block and p-block elements excluding 18 group are called elements.
(a) representative (b) transition (c) inner-transition (d) trans uranium
8. Which of the following is the correct electronic configuration of noble gases?
(a) $ns^2 np^6 nd^{10}$ (b) $ns^2 np^5$ (c) $ns^2 np^6$ (d) $ns^2 np^3$
9. Which of the following is not a periodic property?
(a) Atomic radius (b) Ionization enthalpy (c) Electron affinity (d) Oxidation number
10. Which of the following property increases as we go down the group in the periodic property?
(a) ionization energy (b) Electro negativity (c) Atomic radius (d) Electron affinity
11. Which one of the following is not an iso electronic ion?
(a) Na^+ (b) Mg^{2+} (c) Cl^- (d) O^{2-}
12. Which of the following possess almost same properties due to lanthanide contraction?
(a) Zr, Hf (b) Na, K (c) Zn, Cd (d) Ag, Au
13. Which of the following have zero electron gain enthalpy?
(a) Halogens (b) Noble gases (c) Chalcogens (d) Gold
14. Among all the elements which one has the highest value of electronegativity?
(a) Chlorine (b) Bromine (c) Fluorine (d) Iodine
15. Among the alkali metals which one form compounds with more covalent character?
(a) Sodium (b) Potassium (c) Rubidium (d) Lithium
16. Which of the following pair is not diagonally related? (a) Li, Mg (b) Li, Na (c) Be, Al (d) B, Si
17. Considering the elements B, Al, Mg and K, the correct order of their metallic character is:
(a) $B > Al > Mg > K$ (b) $Al > Mg > B > K$ (c) $Mg > Al > K > B$ (d) $K > Mg > Al > B$
18. Which one of the following is isoelectronic with Ne?
(a) N^{3-} (b) Mg^{2+} (c) Al^{3+} (d) All the above

4. Hydrogen

EVALUATION:

I. Choose the correct answer:

1. Which of the following statements about hydrogen is incorrect ?
a) Hydrogen ion, H_3O^+ exists freely in solution.
b) Dihydrogen acts as a reducing agent.
c) Hydrogen has three isotopes of which tritium is the most common.
d) Hydrogen never acts as cation in ionic salts.
2. Water gas is (a) $H_2O(g)$ (b) $CO + H_2O$ (c) $CO + H_2$ (d) $CO + N_2$
3. Which one of the following statements is incorrect with regard to ortho and para dihydrogen ?
a) They are nuclear spin isomers
b) Ortho isomer has zero nuclear spin whereas the para isomer has one nuclear spin
c) The para isomer is favoured at low temperatures
d) The thermal conductivity of the para isomer is 50% greater than that of the ortho isomer.
4. Ionic hydrides are formed by (a) halogens (b) chalcogens (c) inert gases (d) group one elements
5. Tritium nucleus contains (a) $1p + 0n$ (b) $2p + 1n$ (c) $1p + 2n$ (d) none of these
6. Non-stoichiometric hydrides are formed by (a) palladium, vanadium (b) carbon, nickel
(c) manganese, lithium (d) nitrogen, chlorine

7. Assertion : Permanent hardness of water is removed by treatment with washing soda.
Reason : Washing soda reacts with soluble calcium and magnesium chlorides and sulphates in hard water to form insoluble carbonates
- Both assertion and reason are true and reason is the correct explanation of assertion.
 - Both assertion and reason are true but reason is not the correct explanation of assertion.
 - Assertion is true but reason is false
 - Both assertion and reason are false
8. If a body of a fish contains 1.2 g hydrogen in its total body mass, if all the hydrogen is replaced with deuterium then the increase in body weight of the fish will be
- 1.2 g
 - 2.4 g
 - 3.6 g
 48. g
9. The hardness of water can be determined by volumetrically using the reagent
- sodium thio sulphate
 - potassium permanganate
 - hydrogen peroxide
 - EDTA
10. The cause of permanent hardness of water is due to
- $\text{Ca}(\text{HCO}_3)_2$
 - $\text{Mg}(\text{HCO}_3)_2$
 - CaCl_2
 - MgCO_3
11. Zeolite used to soften hardness of water is, hydrated
- Sodium aluminium silicate
 - Calcium aluminium silicate
 - Zinc aluminium borate
 - Lithium aluminium hydride
12. A commercial sample of hydrogen peroxide marked as 100 volume H_2O_2 , it means that
- 1 ml of H_2O_2 will give 100 ml O_2 at STP
 - 1 L of H_2O_2 will give 100 ml O_2 at STP
 - 1 L of H_2O_2 will give 22.4 L O_2
 - 1 ml of H_2O_2 will give 1 mole of O_2 at STP
13. When hydrogen peroxide is shaken with an acidified solution of potassium dichromate in presence of ether, the ethereal layer turns blue due to the formation of
- Cr_2O_3
 - CrO_4^{2-}
 - $\text{CrO}(\text{O}_2)_2$
 - none of these
14. For decolourisation of 1 mole of acidified KMnO_4 , the moles of H_2O_2 required is
- 1/2
 - 3/2
 - 5/2
 - 7/2
15. Volume strength of 1.5 N H_2O_2 is
- 1.5
 - 4.5
 - 16.8
 - 8.4
16. The hybridisation of oxygen atom in H_2O and H_2O_2 are, respectively
- sp and sp^3
 - sp and sp
 - sp and sp^2
 - sp^3 and sp^3
17. The reaction $\text{H}_3\text{PO}_2 + \text{D}_2\text{O} \rightarrow \text{H}_2\text{DPO}_2 + \text{HDO}$ indicates that hypo-phosphorus acid is
- tribasic acid
 - dibasic acid
 - mono basic acid
 - none of these
18. In solid ice, oxygen atom is surrounded
- tetrahedrally by 4 hydrogen atoms
 - octahedrally by 2 oxygen and 4 hydrogen atoms
 - tetrahedrally by 2 hydrogen and 2 oxygen atoms
 - octahedrally by 6 hydrogen atoms
19. The type of H-bonding present in ortho nitro phenol and p-nitro phenol are respectively
- inter molecular H-bonding and intra molecular H-bonding
 - intra molecular H-bonding and inter molecular H-bonding
 - intra molecular H - bonding and no H - bonding
 - intra molecular H - bonding and intra molecular H - bonding
20. Heavy water is used as
- moderator in nuclear reactions
 - coolant in nuclear reactions
 - both (a) and (b)
 - none of these
21. Water is a
- basic oxide
 - acidic oxide
 - amphoteric oxide
 - none of these

ADDITIONAL QUESTIONS :

1. Which one of the metal is used to convert para-hydrogen into ortho hydrogen?
- Copper
 - Aluminium
 - Sodium
 - Platinum

2. The radioactive isotope of hydrogen is
 (a) protium (b) deuterium (c) tritium (d) heavy hydrogen
3. The half-life period of tritium is ____ (in years). (a) 13.2 (b) 10.5 (c) 12.3 (d) 15.8
4. The composition of syngas is (a) CO + N₂ (b) CO + H₂O (c) CO + H₂ (d) CO₂ + H₂
5. The most common metal ions present in hard water are
 (a) Magnesium and Iron (b) Calcium and Aluminium
 (c) Magnesium and Calcium (d) Manganese and Calcium
6. The permanent hardness of water is due to the presence of soluble salts of ____ and ____ of magnesium and calcium.
 (a) carbonates and bicarbonates (b) chlorides and carbonates
 (c) bicarbonates and sulphates (d) chlorides and sulphates
7. The general formula of zeolites is (a) NaOAl₂O₃. xSiO₂. yH₂O (b) Na₂O.Al₂O₃. ySiO₂. xH₂O
 (c) NaOH.Al₂O₃. xSiO₂. yH₂O (d) NaO.Al (OH)₃. xSiO₂. yH₂O
8. ____ is used as a moderator and coolant in nuclear reactors.
 (a) Heavy hydrogen (b) Ortho hydrogen (c) Hydrogen peroxide (d) Heavy water
9. In chelating method of softening of hard water is used.
 (a) magnesia (b) lime (c) EDTA (d) washing soda
10. The percentage of hydrogen peroxide in '100 volume' is (a) 40 (b) 30 (c) 50 (d) 20
11. Which of the following molecule shows an intramolecular hydrogen bond?
 (a) Water (b) Ammonia (c) Salicylaldehyde (d) Para-nitrophenol
12. Metallic hydrides are otherwise called
 (a) Salt hydrides (b) Saline hydrides (c) molecular hydrides (d) Interstitial hydrides
13. Which of the following is the correct order of stability of bonds?
 (a) Hydrogen bond < Covalent bond < Vanderwaals bond
 (b) Vanderwaals bond < Hydrogen bond < Covalent bond
 (c) Vanderwaals bond > Hydrogen bond > Covalent bond
 (d) Covalent bond < Hydrogen bond < Vanderwaals bond
14. Which of the following contains intramolecular hydrogen bonding?
 (a) Acetic acid (b) o-nitrophenol (c) Hydrogen fluoride (d) water
15. Which one of the following is an example for Clathrate hydrate?
 (a) CuSO₄.5H₂O (b) Na₂CO₃. 10H₂O (c) CH₄. 20 H₂O (d) FeSO₃.7H₂O

5. Alkali and Alkaline Earth Metals

EVALUATION:

I. Choose the correct answer:

1. For alkali metals, which one of the following trends is incorrect ?
 a) Hydration energy : Li > Na > K > Rb
 b) Ionisation energy : Li > Na > K > Rb
 c) Density : Li < Na < K < Rb
 d) Atomic size : Li < Na < K < Rb
2. Which of the following statements is incorrect ?
 a) Li⁺ has minimum degree of hydration among alkali metal cations.
 b) The oxidation state of K in KO₂ is +1
 c) Sodium is used to make Na / Pb alloy
 d) MgSO₄ is readily soluble in water
3. Which of the following compounds will not evolve H₂ gas on reaction with alkali metals ?
 a) ethanoic acid (b) ethanol (c) phenol (d) none of these
4. Which of the following has the highest tendency to give the reaction

$$M^+(g) \xrightarrow[\text{Medium}]{\text{Aqueous}} M^+(aq)$$

 a) Na (b) Li (c) Rb (d) K

5. Sodium is stored in a) alcohol b) water c) kerosene d) none of these
6. RbO_2 is a) superoxide and paramagnetic b) peroxide and diamagnetic
c) superoxide and diamagnetic d) peroxide and paramagnetic
7. Find the wrong statement
a) sodium metal is used in organic qualitative analysis
b) sodium carbonate is soluble in water and it is used in inorganic qualitative analysis
c) potassium carbonate can be prepared by solvay process
d) potassium bicarbonate is acidic salt
8. Lithium shows diagonal relationship with a) sodium b) magnesium c) calcium d) aluminium
9. In case of alkali metal halides, the ionic character increases in the order
a) $\text{MF} < \text{MCl} < \text{MBr} < \text{MI}$ b) $\text{MI} < \text{MBr} < \text{MCl} < \text{MF}$ c) $\text{MI} < \text{MBr} < \text{MF} < \text{MCl}$ d) none of these
10. In which process, fused sodium hydroxide is electrolysed for extraction of sodium?
a) Castner's process b) Cyanide process c) Down process d) All of these
11. The product obtained as a result of a reaction of nitrogen with CaC_2 is
a) $\text{Ca}(\text{CN})_3$ b) CaN_2 c) $\text{Ca}(\text{CN})_2$ d) Ca_3N_2
12. Which of the following has highest hydration energy?
a) MgCl_2 b) CaCl_2 c) BaCl_2 d) SrCl_2
13. Match the flame colours of the alkali and alkaline earth metal salts in the bunsen burner
(p) Sodium (1) Brick red
(q) Calcium (2) Yellow
(r) Barium (3) Lilac (violet)
(s) Strontium (4) Apple green
(t) Cesium (5) Crimson red
(u) Potassium (6) Blue
a) p - 2, q - 1, r - 4, s - 5, t - 6, u - 3
b) p - 1, q - 2, r - 4, s - 5, t - 6, u - 3
c) p - 4, q - 1, r - 2, s - 3, t - 5, u - 6
d) p - 6, q - 5, r - 4, s - 3, t - 1, u - 2
14. Assertion : Generally alkali and alkaline earth metals form superoxides
Reason : There is a single bond between O and O in superoxides.
a) both assertion and reason are true and reason is the correct explanation of assertion
b) both assertion and reason are true but reason is not the correct explanation of assertion
c) assertion is true but reason is false
d) both assertion and reason are false
15. Assertion : BeSO_4 is soluble in water while BaSO_4 is not
Reason : Hydration energy decreases down the group from Be to Ba and lattice energy remains almost constant.
a) both assertion and reason are true and reason is the correct explanation of assertion
b) both assertion and reason are true but reason is not the correct explanation of assertion
c) assertion is true but reason is false
d) both assertion and reason are false
16. Which is the correct sequence of solubility of carbonates of alkaline earth metals?
a) $\text{BaCO}_3 > \text{SrCO}_3 > \text{CaCO}_3 > \text{MgCO}_3$ b) $\text{MgCO}_3 > \text{CaCO}_3 > \text{SrCO}_3 > \text{BaCO}_3$
c) $\text{CaCO}_3 > \text{BaCO}_3 > \text{SrCO}_3 > \text{MgCO}_3$ d) $\text{BaCO}_3 > \text{CaCO}_3 > \text{SrCO}_3 > \text{MgCO}_3$

17. In context with beryllium, which one of the following statements is incorrect ?
 a) It is rendered passive by nitric acid b) It forms Be_2C
 c) Its salts are rarely hydrolysed d) Its hydride is electron deficient and polymeric
18. The suspension of slaked lime in water is known as
 a) lime water b) quick lime c) milk of lime d) aqueous solution of slaked lime
19. A colourless solid substance (A) on heating evolved CO_2 and also gave a white residue, soluble in water. Residue also gave CO_2 when treated with dilute HCl.
 a) Na_2CO_3 b) NaHCO_3 c) CaCO_3 d) $\text{Ca}(\text{HCO}_3)_2$
20. The compound (X) on heating gives a colourless gas and a residue that is dissolved in water to obtain (B). Excess of CO_2 is bubbled through aqueous solution of B, C is formed. Solid (C) on heating gives back X. (B) is
 a) CaCO_3 b) $\text{Ca}(\text{OH})_2$ c) Na_2CO_3 d) NaHCO_3
21. Which of the following statement is false ?
 a) Ca^{2+} ions are not important in maintaining the regular beating of the heart
 b) Mg^{2+} ions are important in the green parts of the plants
 c) Mg^{2+} ions form a complex with ATP
 d) Ca^{2+} ions are important in blood clotting
22. The name 'Blue John' is given to which of the following compounds ?
 a) CaH_2 b) CaF_2 c) $\text{Ca}_3(\text{PO}_4)_2$ d) CaO
23. Formula of Gypsum is a) $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ b) $\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$ c) $3\text{CaSO}_4 \cdot \text{H}_2\text{O}$ d) $2\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$
24. When CaC_2 is heated in atmospheric nitrogen in an electric furnace the compound formed is
 a) $\text{Ca}(\text{CN})_2$ b) CaNCN c) CaC_2N_2 d) CaNC_2
25. Among the following the least thermally stable is
 (a) K_2CO_3 b) Na_2CO_3 (c) BaCO_3 d) Li_2CO_3

ADDITIONAL QUESTIONS :

1. The alkali metal used in photoelectric cells is (a) Na (b) Cs (c) Rb (d) Fr
2. The most common oxidation state of alkali metals is(a) +1 (b) +2 (c) +3 (d) +5
3. The colour produced by potassium when burnt in Bunsen flame is
 (a) red (b) blue (c) green (d) lilac
4. Which of the following ions are more responsible for transmission of nerve signal?
 (a) Li^+ (b) Rb^+ (c) Cs^+ (d) K^+
5. Celestite and strontianite are the ores of(a) cesium (b) strontium (c) magnesium (d) barium
6. The ammonia used in the Solvay process- recovered by using
 (a) calcium chloride (b) Calcium hydroxide (c) calcium carbonate (d) calcium oxide
7. Hydroxides of beryllium are in nature. (a) neutral (b) basic (c) acidic (d) amphoteric
8. Consider the following statements. (i) BeO is basic. (ii) MgO is weakly basic. (iii) BaO is strongly acidic. Which of the above statements is/are not correct?
 (a) (i) only (b) (ii) only (c) (ii) and (iii) (d) (i) and (iii)
9. Which one of the following is named as bleaching powder?
 (a) CaCl_2 (b) CaOCl (c) $\text{Ca}(\text{OCl})_2$ (d) $\text{Ca}(\text{HCO}_3)_2$
10. About 393K, when Plaster of Paris is heated, it forms
 (a) burnt alum (b) dead burnt plaster (c) gypsum plaster (d) alabaster
11. The reducing property of alkali metals follows the order
 (a) $\text{Na} < \text{K} < \text{Rb} < \text{Cs} < \text{Li}$ (b) $\text{K} < \text{Na} < \text{Rb} < \text{Cs} < \text{Li}$ (c) $\text{Li} < \text{Cs} < \text{Rb} < \text{K} < \text{Na}$ (a) $\text{Rb} < \text{Cs} < \text{K} < \text{Na} < \text{Li}$
12. Which of the following is not a peroxide? (a) KO_2 (b) CrO_5 (c) Na_2O_2 (d) BaO_2

13. Which of the following is used as a coolant in fast breeder nuclear reactor?

- (a) Liquid ammonia (b) Liquid helium (c) Liquid Na metal (d) Solid CO₂

6. Gaseous State

EVALUATION:

1. Gases deviate from ideal behavior at high pressure. Which of the following statement(s) is correct for non-ideality?

- a) at high pressure the collision between the gas molecule become enormous
b) at high pressure the gas molecules move only in one direction
c) at high pressure, the volume of gas become insignificant
d) at high pressure the intermolecular interactions become significant

2. Rate of diffusion of a gas is

- a) directly proportional to its density
b) directly proportional to its molecular weight
c) directly proportional to its square root of its molecular weight
d) inversely proportional to the square root of its molecular weight

3. Which of the following is the correct expression for the equation of state of van der Waals gas?

(a) $\left(P + \frac{a}{n^2V^2}\right)(V - nb) = nRT$

(b) $\left(P + \frac{na}{n^2V^2}\right)(V - nb) = nRT$

(c) $\left(P + \frac{an^2}{V^2}\right)(V - nb) = nRT$

(d) $\left(P + \frac{n^2a^2}{V^2}\right)(V - nb) = nRT$

4. When an ideal gas undergoes unrestrained expansion, no cooling occurs because the molecules

- a) are above inversion temperature (b) exert no attractive forces on each other
c) do work equal to the loss in kinetic energy (d) collide without loss of energy

5. Equal weights of methane and oxygen are mixed in an empty container at 298 K. The fraction of total pressure exerted by oxygen is

- (a) $\frac{1}{3}$ (b) $\frac{1}{2}$ (c) $\frac{2}{3}$ (d) $\frac{1}{3} \times 273 \times 298$

6. The temperature at which real gases obey the ideal gas laws over a wide range of pressure is called

- a) Critical temperature (b) Boyle temperature (c) Inversion temperature (d) Reduced temperature

7. In a closed room of 1000 m³ a perfume bottle is opened up. The room develops a smell. This is due to which property of gases? a) Viscosity (b) Density (c) Diffusion (d) None

8. A bottle of ammonia and a bottle of HCl connected through a long tube are opened simultaneously at both ends. The white ammonium chloride ring first formed will be

- a) At the center of the tube (b) Near the hydrogen chloride bottle
c) Near the ammonia bottle (d) Throughout the length of the tube

9. The value of universal gas constant depends upon

- a) Temperature of the gas (b) Volume of the gas
c) Number of moles of the gas (d) units of Pressure and volume.

10. The value of the gas constant R is

- a) 0.082 dm³atm. (b) 0.987 cal mol⁻¹ K⁻¹ (c) 8.3 J mol⁻¹K⁻¹ (d) 8 erg mol⁻¹K⁻¹

11. Use of hot air balloon in sports and meteorological observation is an application of

- a) Boyle's law (b) Newton's law (c) Kelvin's law (d) Brown's law

12. The table indicates the value of van der Waals constant 'a' in $(\text{dm}^3)^2 \text{atm. mol}^{-2}$

| Gas | O ₂ | N ₂ | NH ₃ | CH ₄ |
|-----|----------------|----------------|-----------------|-----------------|
| a | 1.360 | 1.390 | 4.170 | 2.253 |

The gas which can be most easily liquefied is a) O₂ b) N₂ c) NH₃ d) CH₄

13. Consider the following statements

- i) Atmospheric pressure is less at the top of a mountain than at sea level
 - ii) Gases are much more compressible than solids or liquids
 - iii) When the atmospheric pressure increases the height of the mercury column rises
- Select the correct statement a) I and II b) II and III c) I and III d) I, II and III

14. Compressibility factor for CO₂ at 400 K and 71.0 bar is 0.8697. The molar volume of CO₂ under these conditions is a) 22.04 dm³ b) 2.24 dm³ c) 0.41 dm³ d) 19.5 dm³

15. If temperature and volume of an ideal gas is increased to twice its values, the initial pressure P becomes

- a) 4P b) 2P c) P d) 3P

16. At identical temperature and pressure, the rate of diffusion of hydrogen gas is 33 times that of a hydrocarbon having molecular formula C_nH_{2n-2}. What is the value of n? a) 8 b) 4 c) 3 d) 1

17. Equal moles of hydrogen and oxygen gases are placed in a container, with a pin-hole through which both can escape what fraction of oxygen escapes in the time required for one-half of the hydrogen to escape.

- a) 3/8 b) 1/2 c) 1/8 d) 1/4

18. The variation of volume V, with temperature T, $\alpha = \frac{1}{V} \left(\frac{\partial V}{\partial T} \right)_P$ keeping pressure constant is called the coefficient of thermal expansion i.e. For an ideal gas α is equal to

- a) T b) 1/T c) P d) none of these

19. Four gases P, Q, R and S have almost same values of 'b' but their 'a' values (a, b are Vander Waals Constants) are in the order Q < R < S < P. At a particular temperature, among the four gases the most easily liquefiable one is a) P b) Q c) R d) S

20. Maximum deviation from ideal gas is expected from a) CH₄ (g) b) NH₃ (g) c) H₂ (g) d) N₂ (g)

21. The units of Vander Waals constants 'b' and 'a' respectively

- a) mol L⁻¹ and L atm² mol⁻¹ b) mol L and L atm mol² c) mol⁻¹L and L² atm mol⁻² d) none of these

22. Assertion : Critical temperature of CO₂ is 304K, it can be liquefied above 304K.

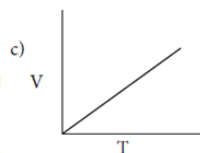
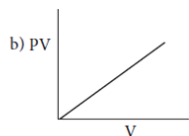
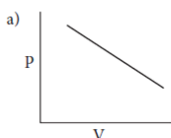
Reason : For a given mass of gas, volume is directly proportional to pressure at constant temperature

- a) both assertion and reason are true and reason is the correct explanation of assertion
- b) both assertion and reason are true but reason is not the correct explanation of assertion
- c) assertion is true but reason is false
- d) both assertion and reason are false

23. What is the density of N₂ gas at 227°C and 5.00 atm pressure? (R = 0.082 L atm K⁻¹ mol⁻¹)

- a) 1.40 g/L b) 2.81 g/L c) 3.41 g/L d) 0.29 g/L

24. Which of the following diagrams correctly describes the behaviour of a fixed mass of an ideal gas? (T is measured in K)



d) All of these

25. 25g of each of the following gases are taken at 27°C and 600 mm Hg pressure. Which of these will have the least volume ? a) HBr b) HCl c) HF d) HI

Additional :

- At constant temperature, the pressure of the gas is reduced to one-third, the volume
(a) reduce to one-third (b) increases by three times (c) remaining the same (d) cannot be predicted
- At constant temperature for a given mass, for each degree rise in temperature, all gases expand by of their volume at 0°C. (a) 273 (b) 298 (c) 1/273 (d) 1/298
- The rate of diffusion of a gas is inversely proportional to the
(a) square of molar mass (b) square root of density (c) square root of molar mass (d) square of density
- The gas used in the pressure-volume isotherm study of Andrew's experiment is
(a) N₂ (b) H₂S (c) NH₃ (d) CO₂
- Which one of the following is absolute zero? (a) 293 K (b) 273 K (c) - 273.15°C (d) 0°C
- The value of critical temperature of carbon dioxide is (a) 273 K (b) 303.98 K (c) 373 K (d) - 80°C
- The rate of diffusion of gases A and B of molecular weight 36 and 64 are in the ratio
(a) 9 : 16 (b) 4 : 3 (c) 3 : 4 (d) 16 : 9

7. Thermodynamics

I. Choose the correct answer:

- The amount of heat exchanged with the surrounding at constant pressure is given by the quantity
a) ΔE b) ΔH c) ΔS d) ΔG
- All the naturally occurring processes proceed spontaneously in a direction which leads to
a) decrease in entropy b) increase in enthalpy
c) increase in free energy d) decrease in free energy
- In an adiabatic process, which of the following is true ?
a) $q = w$ b) $q = 0$ c) $\Delta E = q$ d) $P \Delta V = 0$
- In a reversible process, the change in entropy of the universe is
a) > 0 b) ≥ 0 c) < 0 d) $= 0$
- In an adiabatic expansion of an ideal gas
a) $w = -\Delta u$ b) $w = \Delta u + \Delta H$ c) $\Delta u = 0$ d) $w = 0$
- The intensive property among the quantities below is
a) mass b) volume c) enthalpy d) mass/volume
- An ideal gas expands from the volume of $1 \times 10^{-3} \text{ m}^3$ to $1 \times 10^{-2} \text{ m}^3$ at 300 K against a constant pressure at $1 \times 10^5 \text{ Nm}^{-2}$. The work done is
a) - 900 J b) 900 kJ c) 270 kJ d) - 900 kJ
- Heat of combustion is always
a) positive b) negative c) zero d) either positive or negative
- The heat of formation of CO and CO₂ are - 26.4 kCal and - 94 kCal, respectively. Heat of combustion of carbon monoxide will be
a) + 26.4 kcal b) - 67.6 kcal c) - 120.6 kcal d) + 52.8 kcal
- $\text{C(diamond)} \rightarrow \text{C(graphite)}$, $\Delta H = -\text{ve}$, this indicates that
a) graphite is more stable than diamond b) graphite has more energy than diamond
c) both are equally stable d) stability cannot be predicted
- The enthalpies of formation of Al₂O₃ and Cr₂O₃ are - 1596 kJ and - 1134 kJ, respectively. ΔH for the reaction $2\text{Al} + \text{Cr}_2\text{O}_3 \rightarrow 2\text{Cr} + \text{Al}_2\text{O}_3$ is
a) - 1365 kJ b) 2730 kJ c) - 2730 kJ d) - 462 kJ
- Which of the following is not a thermodynamic function ?
a) internal energy b) enthalpy c) entropy d) frictional energy
- If one mole of ammonia and one mole of hydrogen chloride are mixed in a closed container to form ammonium chloride gas, then
a) $\Delta H > \Delta U$ b) $\Delta H - \Delta U = 0$ c) $\Delta H + \Delta U = 0$ d) $\Delta H < \Delta U$

14. Change in internal energy, when 4 kJ of work is done on the system and 1 kJ of heat is given out by the system is
a) +1 kJ b) - 5 kJ c) +3 kJ d) - 3 kJ
15. The work done by the liberated gas when 55.85 g of iron (molar mass 55.85 g mol⁻¹) reacts with hydrochloric acid in an open beaker at 250 C
a) - 2.48 kJ b) - 2.22 kJ c) + 2.22 kJ d) + 2.48 kJ
16. The value of ΔH for cooling 2 moles of an ideal monatomic gas from 1250 C to 250 C at constant pressure will be
a) - 250 R b) - 500 R c) 500 R d) + 250 R
[given $C_p = \frac{5}{2} R$]
17. Given that $C(g) + O_2(g) \rightarrow CO_2(g) \Delta H^0 = -a \text{ kJ}$; $2 CO(g) + O_2(g) \rightarrow 2CO_2(g) \Delta H^0 = -b \text{ kJ}$; Calculate the ΔH^0 for the reaction $C(g) + \frac{1}{2}O_2(g) \rightarrow CO(g)$
a) $\frac{b+2a}{2}$ b) $2a-b$ c) $\frac{2a-b}{2}$ d) $\frac{b-2a}{2}$
18. When 15.68 litres of a gas mixture of methane and propane are fully combusted at 0° C and 1 atmosphere, 32 litres of oxygen at the same temperature and pressure are consumed. The amount of heat released from this combustion in kJ is ($\Delta H_{C(CH_4)} = -890 \text{ kJ mol}^{-1}$ and $\Delta H_{C(C_3H_8)} = -2220 \text{ kJ mol}^{-1}$)
a) - 889 kJ b) - 1390 kJ c) - 3180 kJ d) - 632.68 kJ
19. The bond dissociation energy of methane and ethane are 360 kJ mol⁻¹ and 620 kJ mol⁻¹ respectively. Then, the bond dissociation energy of C-C bond is
a) 170 kJ mol⁻¹ b) 50 kJ mol⁻¹ c) 80 kJ mol⁻¹ d) 220 kJ mol⁻¹
20. The correct thermodynamic conditions for the spontaneous reaction at all temperature is
a) $\Delta H < 0$ and $\Delta S > 0$ b) $\Delta H < 0$ and $\Delta S < 0$
c) $\Delta H > 0$ and $\Delta S = 0$ d) $\Delta H > 0$ and $\Delta S > 0$
21. The temperature of the system, decreases in an _____
a) Isothermal expansion b) Isothermal Compression
c) adiabatic expansion d) adiabatic compression
22. In an isothermal reversible compression of an ideal gas the sign of q, ΔS and w are respectively
a) +, -, - b) -, +, - c) +, -, + d) -, -, +
23. Molar heat of vapourisation of a liquid is 4.8 kJ mol⁻¹. If the entropy change is 16 J mol⁻¹ K⁻¹, the boiling point of the liquid is
a) 323 K b) 27° C c) 164 K d) 0.3 K
24. ΔS is expected to be maximum for the reaction
a) $Ca(S) + \frac{1}{2} O_2(g) \rightarrow CaO(S)$ b) $C(S) + O_2(g) \rightarrow CO_2(g)$
c) $N_2(g) + O_2(g) \rightarrow 2NO(g)$ d) $CaCO_3(S) \rightarrow CaO(S) + CO_2(g)$
25. The values of ΔH and ΔS for a reaction are respectively 30 kJ mol⁻¹ and 100 JK⁻¹ mol⁻¹. Then the temperature above which the reaction will become spontaneous is
a) 300 K b) 30 K c) 100 K d) 200 C

ADDITIONAL :

1. The temperature below which a gas obey Joule Thomson effect is called(a) critical temperature
(b) standard temperature (c) inversion temperature (d) normal temperature
2. The temperature produced in adiabatic process of liquefaction is
(a) zero Kelvin (b) -273 K (c) 10⁻⁴ K (d) 10⁴ K
3. Which of the following is/are extensive properties?
1. Volume 2. Surface tension 3. mass 4. internal energy
(a) 1, 2 and 4 (b) 1, 3 and 4 (c) 1 and 3 (d) 1, 2 and 3
4. Which of the following is/are path functions?
1. Pressure 2. Work 3. internal energy 4. Free energy 5. heat
(a) 1, 2 and 4 (b) 1, 3 and 4 (c) 2 and 5 (d) 2, 3 and 4

5. The unit of entropy is(a) $\text{J K}^{-1} \text{mol}^{-1}$ (b) J mol^{-1} (c) J K mol^{-1} (d) $\text{J}^{-1} \text{K}^{-1}$
6. The efficiency of engine working between 100 to 400 K (a) 25% (b) 75% (c) 100% (d) 50%
7. In a reversible process, the entropy of Universe is
(a) greater than zero (b) less than zero (c) equal to zero (d) remains constant
8. The net work done by the system (a) $w - P\Delta V$ (b) $w + P\Delta V$ (c) $-w + P\Delta V$ (d) $-w - P\Delta V$
9. In a reversible process $\Delta S_{\text{sys}} + \Delta S_{\text{surr}}$ is(a) >0 (b) <0 (c) ≥ 0 (d) $=0$
10. The standard free energy change (ΔG°) is related to equilibrium constant (K) as
(a) $\Delta G^\circ = -1303 RT \ln K$ (b) $\Delta G^\circ = 2.303 RT \log K$ (c) $\Delta G^\circ = RT \ln K$ (d) $\Delta G^\circ = -2.303 RT \log K$
11. Which of the following is not an intensive property?
(a) Pressure (b) Density (c) Volume (d) Surface tension
12. The relation between C_p and C_v is.....(a) $C_p - C_v = R$ (b) $C_p + C_v = R$ (c) -285 KJ (d) $R - C_v = C_p$
13. Which of the following always has a negative value? (a) heat of reaction (b) heat of solution
(c) heat of combustion (d) heat of formation
14. Which one of the following is not an intensive property? (a) Density (b) Molaritv (c) Molality (d) Mole
15. Which one of the following is the quantity of heat required to raise the temperature of 1 gm of water by 1°C ?
(a) 1 Joule (b) 1 Calorie (c) 1 Kelvin (d) 1 Kilo joule
16. The standard value of enthalpy of combustion of benzoic acid is
(a) $-3227 \text{ kJ mol}^{-1}$ (b) $+3227 \text{ kJ mol}^{-1}$ (c) $-32.27 \text{ Ici mol}^{-1}$ (d) $+32.27 \text{ kJ mol}^{-1}$
17. The heat of neutralization of a strong acid and strong base is around
(a) $+57.32 \text{ kJ}$ (b) -57.32 kJ (c) $-3227 \text{ kJ mol}^{-1}$ (d) $+3227 \text{ kJ mol}^{-1}$

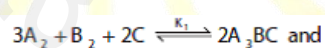
8. Physical and Chemical Equilibrium

EVALUATION:

I. Choose the best answer.

1. If K_b and K_f for a reversible reaction are 0.8×10^{-5} and 1.6×10^{-4} respectively, the value of the equilibrium constant is, a) 20 b) 0.2×10^{-1} c) 0.05 d) none of these

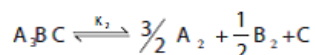
2. At a given temperature and pressure, the equilibrium constant values for the equilibria



$$a) K_1 = \frac{1}{\sqrt{K_2}}$$

$$b) K_2 = K_1^{-1/2}$$

The relation between K_1 and K_2 is



$$c) K_1^2 = 2K_2$$

$$d) \frac{K_1}{2} = K_2$$

3. The equilibrium constant for a reaction at room temperature is K_1 and that at 700 K is K_2 . If $K_1 > K_2$, then

- a) The forward reaction is exothermic b) The forward reaction is endothermic
c) The reaction does not attain equilibrium d) The reverse reaction is exothermic

4. The formation of ammonia from $N_2(g)$ and $H_2(g)$ is a reversible reaction $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) + \text{Heat}$

What is the effect of increase of temperature on this equilibrium reaction

- a) equilibrium is unaltered b) formation of ammonia is favoured
c) equilibrium is shifted to the left d) reaction rate does not change

5. Solubility of carbon dioxide gas in cold water can be increased by

- a) increase in pressure b) decrease in pressure c) increase in volume d) none of these

6. Which one of the following is incorrect statement?

- a) for a system at equilibrium, Q is always less than the equilibrium constant
b) equilibrium can be attained from either side of the reaction
c) presence of catalyst affects both the forward reaction and reverse reaction to the same extent
d) Equilibrium constant varied with temperature

7. K_1 and K_2 are the equilibrium constants for the reactions respectively.
- $$\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \xrightleftharpoons{K_1} 2\text{NO}(\text{g})$$
- $$2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \xrightleftharpoons{K_2} 2\text{NO}_2(\text{g})$$
- What is the equilibrium constant for the reaction $\text{NO}_2(\text{g}) \rightleftharpoons \frac{1}{2}\text{N}_2(\text{g}) + \text{O}_2(\text{g})$
- a) $\frac{1}{\sqrt{K_1 K_2}}$ b) $(K_1 = K_2)^{1/2}$
 c) $\frac{1}{2K_1 K_2}$ d) $\left(\frac{1}{K_1 K_2}\right)^{3/2}$
8. In the equilibrium, $2\text{A}(\text{g}) \rightleftharpoons 2\text{B}(\text{g}) + \text{C}_2(\text{g})$
 the equilibrium concentrations of A, B and C_2 at 400 K are 1×10^{-4} M, 2.0×10^{-3} M, 1.5×10^{-4} M respectively. The value of K_c for the equilibrium at 400 K is
- a) 0.06 b) 0.09 c) 0.62 d) 3×10^{-2}
9. An equilibrium constant of 3.2×10^{-6} for a reaction means, the equilibrium is
- a) largely towards forward direction b) largely towards reverse direction
 c) never established d) none of these
10. $K_c K_p$ for the reaction, $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ is
- a) 1 RT b) RT c) RT d) $(\text{RT})^2$
11. For the reaction $\text{AB}(\text{g}) \rightleftharpoons \text{A}(\text{g}) + \text{B}(\text{g})$, at equilibrium, AB is 20% dissociated at a total pressure of P, The equilibrium constant K_p is related to the total pressure by the expression
- a) $P = 24 K_p$ b) $P = 8 K_p$ c) $24 P = K_p$ d) none of these
12. In which of the following equilibrium, K_p and K_c are not equal?
- a) $2\text{NO}(\text{g}) \rightleftharpoons \text{N}_2(\text{g}) + \text{O}_2(\text{g})$ b) $\text{SO}_2(\text{g}) + \text{NO}_2 \rightleftharpoons \text{SO}_3(\text{g}) + \text{NO}(\text{g})$
 c) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$ d) $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
13. If x is the fraction of PCl_5 dissociated at equilibrium in the reaction $\text{PCl}_5 \rightleftharpoons \text{PCl}_3 + \text{Cl}_2$
 then starting with 0.5 mole of PCl_5 , the total number of moles of reactants and products at equilibrium is
- a) $0.5 - x$ b) $x + 0.5$ c) $2x + 0.5$ d) $x + 1$
14. The values of K_{p1} and K_{p2} for the reactions $\text{X} \rightleftharpoons \text{Y} + \text{Z}$
 $\text{A} \rightleftharpoons 2\text{B}$ are in the ratio 9 : 1 if degree of dissociation and initial concentration of X and A be equal then total pressure at equilibrium P_1 , and P_2 are in the ratio
- a) 36 : 1 b) 1 : 1 c) 3 : 1 d) 1 : 9
15. In the reaction, $\text{Fe}(\text{OH})_3(\text{s}) \rightleftharpoons \text{Fe}^{3+}(\text{aq}) + 3\text{OH}^{-}(\text{aq})$, if the concentration of OH^{-} ions is decreased by $\frac{1}{4}$ times, then the equilibrium concentration of Fe^{3+} will
- a) not changed b) also decreased by $\frac{1}{4}$ times
 c) increase by 4 times d) increase by 64 times
16. Consider the reaction where $K_p = 0.5$ at a particular temperature $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
 if the three gases are mixed in a container so that the partial pressure of each gas is initially 1 atm, then which one of the following is true
- a) more PCl_3 will be produced b) more Cl_2 will be produced
 c) more PCl_5 will be produced d) none of these
17. Equimolar concentrations of H_2 and I_2 are heated to equilibrium in a 1 litre flask. What percentage of initial concentration of H_2 has reacted at equilibrium if rate constant for both forward and reverse reactions are equal
- a) 33% b) 66% c) $(33)^2\%$ d) 16.5 %
18. In a chemical equilibrium, the rate constant for the forward reaction is 2.5×10^2 and the equilibrium constant is 50. The rate constant for the reverse reaction is,
- a) 11.5 b) 5 c) 2×10^2 d) 2×10^{-3}
19. Which of the following is not a general characteristic of equilibrium involving physical process
- a) Equilibrium is possible only in a closed system at a given temperature
 b) The opposing processes occur at the same rate and there is a dynamic but stable condition
 c) All the physical processes stop at equilibrium
 d) All measurable properties of the system remains constant
20. For the formation of Two moles of $\text{SO}_3(\text{g})$ from SO_2 and O_2 , the equilibrium constant is K_1 . The equilibrium constant for the dissociation of one mole of SO_3 into SO_2 and O_2 is
- a) $\frac{1}{K_1}$ b) K_1^2
 c) $\left(\frac{1}{K_1}\right)^{1/2}$ d) $\frac{K_1}{2}$

21. Match the equilibria with the corresponding conditions,

- i) Liquid \rightleftharpoons Vapour ii) Solid \rightleftharpoons Liquid
 iii) Solid \rightleftharpoons Vapour iv) Solute (s) \rightleftharpoons Solute (Solution)
 1) melting point 2) Saturated solution
 3) Boiling point 4) Sublimation point 5) Unsaturated solution

| | (i) | (ii) | (iii) | (iv) |
|-----|-----|------|-------|------|
| (a) | 1 | 2 | 3 | 4 |
| (b) | 3 | 1 | 4 | 2 |
| (c) | 2 | 1 | 3 | 4 |
| (d) | 3 | 2 | 4 | 5 |

22. Consider the following reversible reaction at equilibrium, $A + B \rightleftharpoons C$, If the concentration of the reactants A and B are doubled, then the equilibrium constant will

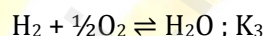
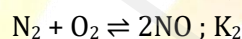
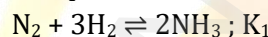
- a) be doubled b) become one fourth c) be halved d) remain the same

23. $[\text{Co}(\text{H}_2\text{O})_6]^{2+} (\text{aq}) (\text{pink}) + 4\text{Cl}^- (\text{aq}) \rightleftharpoons [\text{CoCl}_4]^{2-} (\text{aq}) (\text{blue}) + 6 \text{H}_2\text{O} (\text{l})$

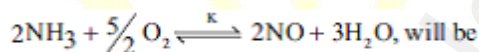
In the above reaction at equilibrium, the reaction mixture is blue in colour at room temperature. On cooling this mixture, it becomes pink in colour. On the basis of this information, which one of the following is true ?

- a) $\Delta H > 0$ for the forward reaction b) $\Delta H = 0$ for the reverse reaction
 c) $\Delta H < 0$ for the forward reaction d) Sign of the ΔH cannot be predicted based on this information.

24. The equilibrium constants of the following reactions are :



The equilibrium constant (K) for the reaction ;



- a) $K_2^3 K_3 / K_1$ b) $K_1 K_3^3 / K_2$
 c) $K_2 K_3^3 / K_1$ d) $K_2 K_3 / K_1$

25. A 20 litre container at 400 K contains $\text{CO}_2 (\text{g})$ at pressure 0.4 atm and an excess of SrO (neglect the volume of solid SrO). The volume of the container is now decreased by moving the movable piston fitted in the container. The maximum volume of the container, when pressure of CO_2 attains its maximum value will be :

Given that : $\text{SrCO}_3 (\text{s}) \rightleftharpoons \text{SrO} (\text{s}) + \text{CO}_2 (\text{g})$

$K_p = 1.6 \text{ atm}$ (NEET 2017)

- a) 2 litre b) 5 litre c) 10 litre d) 4 litre

Additional:

1. For which of the following K_p is less than K_c ?

- a) $\text{N}_2\text{O}_4 \rightleftharpoons 2\text{NO}_2$ b) $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$ c) $\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$ d) $\text{CO} + \text{H}_2\text{O} \rightleftharpoons \text{CO}_2 + \text{H}_2$

2. A reversible reaction is one which

- a) Proceeds in one direction b) proceeds in both direction
 c) proceeds spontaneously d) All the statements are wrong

3. The equilibrium constant in a reversible reaction at a given temperature

- a) depends on the initial concentration of the reactants
 b) depends on the concentration of the products at equilibrium
 c) does not depend on the initial concentrations
 d) It is not characteristic of the reaction

4. A chemical reaction is at equilibrium when

- a) Reactants are completely transformed into products
 b) The rates of forward and backward reactions are equal
 c) Formation of products is minimized
 d) Equal amounts of reactants and products are present

5. The rate constant for forward and backward reactions of the hydrolysis of ester are 1.1×10^{-2} and 1.5×10^{-3} per minute respectively. The equilibrium constant for the reaction is a) 4.33 b) 5.33 c) 6.33 d) 7.33

6. For the reaction $\text{CO} (\text{g}) + 12\text{O}_2 (\text{g}) \rightleftharpoons \text{CO}_2 (\text{g})$, K_p / K_c is _____ a) RT b) $(RT)^{-1}$ c) $(RT)^{-1/2}$ d) $(RT)^{1/2}$

7. In which of the following, the reaction proceeds towards completion
 a) $K = 10^3$ b) $K = 10^{-2}$ c) $K = 10$ d) $K = 1$
8. Under a given set of experimental conditions, with an increase in the concentration of the reactants, the rate of a chemical reaction
 a) Decreases b) Increases c) Remains unaltered d) First decreases and then increases
9. Theory of 'active mass' indicates that the rate of chemical reaction is directly proportional to the _____.
 a) Equilibrium constant b) Properties of reactants
 c) Volume of apparatus d) Concentration of reactants
10. In an equilibrium reaction $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$, $\Delta H = -3000$ calories, which factor favours dissociation of HI
 a) Low temperature b) High pressure c) High temperature d) Low pressure
11. Which of the following factor is shifted the reaction $\text{PCl}_3 + \text{Cl}_2 \rightleftharpoons \text{PCl}_5$ at the left side?
 a) Adding PCl_5 b) increase pressure c) constant temperature d) catalyst

9. Solutions

EVALUATION:

I. Choose the best answer.

1. The molality of a solution containing 1.8g of glucose dissolved in 250g of water is
 a) 0.2 M b) 0.01 M c) 0.02 M d) 0.04 M
2. Which of the following concentration terms is / are independent of temperature
 a) molality b) molarity c) mole fraction d) (a) and (c)
3. Stomach acid, a dilute solution of HCl can be neutralised by reaction with Aluminium hydroxide
 $\text{Al}(\text{OH})_3 + 3\text{HCl}(\text{aq}) \rightarrow \text{AlCl}_3 + 3\text{H}_2\text{O}$
 How many millilitres of 0.1 M $\text{Al}(\text{OH})_3$ solution are needed to neutralise 21 mL of 0.1 M HCl ?
 a) 14 mL b) 7 mL c) 21 mL d) none of these
4. The partial pressure of nitrogen in air is 0.76 atm and its Henry's law constant is 7.6×10^4 atm at 300K. What is the molefraction of nitrogen gas in the solution obtained when air is bubbled through water at 300K ?
 a) 1×10^{-4} b) 1×10^{-6} c) 2×10^{-5} d) 1×10^{-5}
5. The Henry's law constant for the solubility of Nitrogen gas in water at 350 K is 8×10^4 atm. The mole fraction of nitrogen in air is 0.5. The number of moles of Nitrogen from air dissolved in 10 moles of water at 350K and 4 atm pressure is
 a) 4×10^{-4} b) 4×10^4 c) 2×10^{-2} d) 2.5×10^{-4}
6. Which one of the following is incorrect for ideal solution ?
 a) $\Delta H_{\text{mix}} = 0$ b) $\Delta U_{\text{mix}} = 0$ c) $\Delta P = P_{\text{observed}} - P_{\text{Calculated by Raoult's law}} = 0$ d) $\Delta G_{\text{mix}} = 0$
7. Which one of the following gases has the lowest value of Henry's law constant ?
 a) N_2 b) He c) CO_2 d) H_2
8. P_1 and P_2 are the vapour pressures of pure liquid components, 1 and 2 respectively of an ideal binary solution if x_1 represents the mole fraction of component 1, the total pressure of the solution formed by 1 and 2 will be
 a) $P_1 + x_1 (P_2 - P_1)$ b) $P_2 - x_1 (P_2 + P_1)$ c) $P_1 - x_2 (P_1 - P_2)$ d) $P_1 + x_2 (P_1 - P_2)$
9. Osmotic pressure (p) of a solution is given by the relation
 a) $p = nRT$ b) $pV = nRT$ c) $pRT = n$ d) none of these
10. Which one of the following binary liquid mixtures exhibits positive deviation from Raoult's law ?
 a) Acetone + chloroform b) Water + nitric acid c) HCl + water d) ethanol + water

11. The Henry's law constants for two gases A and B are x and y respectively. The ratio of mole fractions of A to B is 0.2. The ratio of mole fraction of B and A dissolved in water will be
 a) $2x/y$ b) $y/0.2x$ c) $0.2x/y$ d) $5x/y$
12. At 100°C the vapour pressure of a solution containing 6.5g a solute in 100g water is 732mm. If $K_b = 0.52$, the boiling point of this solution will be a) 102°C b) 100°C c) 101°C d) 100.52°C
13. According to Raoult's law, the relative lowering of vapour pressure for a solution is equal to
 a) mole fraction of solvent b) mole fraction of solute
 c) number of moles of solute d) number of moles of solvent
14. At same temperature, which pair of the following solutions are isotonic ?
 a) 0.2 M BaCl_2 and 0.2M urea b) 0.1 M glucose and 0.2 M urea
 c) 0.1 M NaCl and 0.1 M K_2SO_4 d) 0.1 M $\text{Ba}(\text{NO}_3)_2$ and 0.1 M Na_2SO_4
15. The empirical formula of a non-electrolyte(X) is CH_2O . A solution containing six gram of X exerts the same osmotic pressure as that of 0.025M glucose solution at the same temperature. The molecular formula of X is
 a) $\text{C}_2\text{H}_4\text{O}_2$ b) $\text{C}_8\text{H}_{16}\text{O}_8$ c) $\text{C}_4\text{H}_8\text{O}_4$ d) CH_2O
16. The K_H for the solution of oxygen dissolved in water is $4 \times 10^4 \text{ atm}$ at a given temperature. If the partial pressure of oxygen in air is 0.4 atm, the mole fraction of oxygen in solution is
 a) 4.6×10^3 b) 1.6×10^4 c) 1×10^{-5} d) 1×10^5
17. Normality of 1.25M sulphuric acid is a) 1.25 N b) 3.75 N c) 2.5 N d) 2.25 N
18. Two liquids X and Y on mixing gives a warm solution. The solution is
 a) ideal b) non-ideal and shows positive deviation from Raoult's law
 c) ideal and shows negative deviation from Raoult's Law
 d) non-ideal and shows negative deviation from Raoult's Law
19. The relative lowering of vapour pressure of a sugar solution in water is 3.5×10^{-3} . The mole fraction of water in that solution is a) 0.0035 b) 0.35 c) $0.0035/18$ d) 0.9965
20. The mass of a non-volatile solute (molar mass 80 g mol^{-1}) which should be dissolved in 92g of toluene to reduce its vapour pressure to 90% a) 10g b) 20g c) 9.2 g d) 8.89g
21. For a solution, the plot of osmotic pressure (p) verses the concentration (c in mol L^{-1}) gives a straight line with slope $310R$ where 'R' is the gas constant. The temperature at which osmotic pressure measured is
 a) $310 \times 0.082 \text{ K}$ b) 310°C c) 37°C d) $310 \text{ K}/0.082$
22. 200ml of an aqueous solution of a protein contains 1.26g of protein. At 300K, the osmotic pressure of this solution is found to be $2.52 \times 10^{-3} \text{ bar}$. The molar mass of protein will be ($R = 0.083 \text{ L bar mol}^{-1} \text{ K}^{-1}$)
 a) $62.22 \text{ Kg mol}^{-1}$ b) 12444 g mol^{-1} c) 300 g mol^{-1} d) none of these
23. The Van't Hoff factor (i) for a dilute aqueous solution of the strong electrolyte barium hydroxide is
 a) 0 b) 1 c) 2 d) 3
24. What is the molality of a 10% W/W aqueous sodium hydroxide solution ? a) 2.778 b) 2.5 c) 10 d) 0.4
25. The correct equation for the degree of an associating solute, 'n' molecules of which undergoes association in solution, is

$$\text{a) } \alpha = \frac{n(i-1)}{n-1} \quad \text{b) } \alpha^2 = \frac{n(1-i)}{(n-1)}$$

$$\text{c) } \alpha = \frac{n(i-1)}{1-n} \quad \text{d) } \alpha = \frac{n(1-i)}{n(1-i)}$$

26. Which of the following aqueous solutions has the highest boiling point ?

- a) 0.1M KNO_3 b) 0.1 MNa_3PO_4 c) 0.1 M BaCl_2 d) 0.1 M K_2SO_4

27. The freezing point depression constant for water is $1.86^\circ \text{K Kg mol}^{-1}$. If 5g Na_2SO_4 is dissolved in 45g water, the depression in freezing point is 3.64°C . The Vant Hoff factor for Na_2SO_4 is

- a) 2.50 b) 2.63 c) 3.64 d) 5.50

28. Equimolal aqueous solutions of NaCl and KCl are prepared. If the freezing point of NaCl is -2°C , the freezing point of KCl solution is expected to be

- a) -2°C b) -4°C c) -1°C d) 0°C

29. Phenol dimerises in benzene having van't Hoff factor 0.54. What is the degree of association ?

- a) 0.46 b) 92 c) 46 d) 0.92

30. Assertion : An ideal solution obeys Raoult's Law

Reason : In an ideal solution, solvent-solvent as well as solute-solute interactions are similar to solute-solvent interactions.

- a) both assertion and reason are true and reason is the correct explanation of assertion
b) both assertion and reason are true but reason is not the correct explanation of assertion
c) assertion is true but reason is false
d) both assertion and reason are false

Additional:

- In which mode of expression of concentration of a solution remains independent of temperature?
a) Molarity b) Normality c) Formality d) Molality
- Which of the following is correct for a solution showing positive deviations from Raoult's law?
a) $\Delta V = +ve$, $\Delta H = +ve$ b) $\Delta V = -ve$, $\Delta H = -ve$ c) $\Delta V = +ve$, $\Delta H = -ve$ d) $\Delta V = -ve$, $\Delta H = +ve$
- If liquids A and B form an ideal solution
a) The entropy of mixing is zero b) The Gibbs free energy is zero
c) The Gibbs free energy as well as the entropy of mixing are each zero
d) The enthalpy of mixing is zero
- Water and ethanol form non - ideal solution with positive deviation from Raoult's law. This solution, will have vapour pressure
a) equal to vapour pressure of pure water b) less than vapour pressure of pure water
c) more than vapour pressure of pure water
d) less than vapour pressure of pure ethanol
- Which liquid pair shows a positive deviation from Raoult's law ?
a) Acetone - chloroform b) Benzene - methanol c) Water - nitric acid d) Water - hydrochloric acid
- In the phenomenon of osmosis, the membrane allow passage of _____.
a) Solute only b) Solvent only c) Both solute and solvent d) None of these
- Molecular weight of glucose is 180, A solution of glucose which contains 18 g per liter is
a) 2 molal b) 1 molal c) 0.1 molal d) 18 molal
- Normality of 2 M sulphuric acid is
a) 2 N b) 4 N c) N/2 d) N/4
- Which of the following is a colligative property?
a) Osmotic pressure b) Boiling point c) Vapour pressure d) Freezing point
- The freezing point of 0.2 molal K_2SO_4 is -1.1°C . Calculate van't Hoff factor and percentage degree of dissociation of K_2SO_4 . K_f for water is 1.86°
a) 97.5 b) 90.75 c) 105.5 d) 85.75

10. Chemical Bonding

EVALUATION:

I. Choose the best answer.

- In which of the following Compounds does the central atom obey the octet rule?
a) XeF_4 b) AlCl_3 c) SF_6 d) SCl_2
- In the molecule $\text{O}_\text{A} = \text{C} = \text{O}_\text{B}$ the formal charge on O_A , C and O_B are respectively.
a) -1, 0, +1 b) +1, 0, -1 c) -2, 0, +2 d) 0, 0, 0
- Which of the following is electron deficient? a) PH_3 b) $(\text{CH}_3)_2$ c) BH_3 d) NH_3
- Which of the following molecule contain no π bond? a) SO_2 b) NO_2 c) CO_2 d) H_2O
- The ratio of number of sigma (σ) and pi (π) bonds in 2- butynal is
a) 8/3 b) 5/3 c) 8/2 d) 9/2
- Which one of the following is the likely bond angles of sulphur tetrafluoride molecule?
a) $120^\circ, 80^\circ$ b) $109^\circ 28'$ c) 90° d) $89^\circ, 117^\circ$
- Assertion:** Oxygen molecule is paramagnetic.
Reason : It has two unpaired electron in its bonding molecular orbital
a) both assertion and reason are true and reason is the correct explanation of assertion
b) both assertion and reason are true but reason is not the correct explanation of assertion
c) assertion is true but reason is false d) Both assertion and reason are false
- According to Valence bond theory, a bond between two atoms is formed when
a) fully filled atomic orbitals overlap b) half filled atomic orbitals overlap
c) non- bonding atomic orbitals overlap d) empty atomic orbitals overlap
- In ClF_3 , NF_3 and BF_3 molecules the chlorine, nitrogen and boron atoms are
a) sp^3 hybridised b) sp^3 , sp^3 and sp^2 respectively
c) sp^2 hybridised d) sp^3d , sp^3 and sp^2 hybridised respectively
- When one s and three p orbitals hybridise,
a) four equivalent orbitals at 90° to each other will be formed
b) four equivalent orbitals at $109^\circ 28'$ to each other will be formed.
c) four equivalent orbitals, that are lying the same plane will be formed
d) none of these
- Which of these represents the correct order of their increasing bond order.
a) $\text{C}_2 < \text{C}_2^{2-} < \text{O}_2^{2-} < \text{O}_2$ b) $\text{C}_2^{2-} < \text{C}_2^+ < \text{O}_2 < \text{O}_2^{2-}$
c) $\text{O}_2^{2-} < \text{O}_2 < \text{C}_2^{2-} < \text{C}_2^+$ d) $\text{O}_2^{2-} < \text{C}_2^+ < \text{O}_2 < \text{C}_2^{2-}$
- Hybridisation of central atom in PCl_5 involves the mixing of orbitals.
a) s, px, py, dx^2 , $\text{dx}^2\text{-y}^2$ b) s, px, py, pxy, $\text{dx}^2\text{-y}^2$
c) s, px, py, pz, $\text{dx}^2\text{-y}^2$ d) s, px, py, dxy, $\text{dx}^2\text{-y}^2$
- The correct order of O-O bond length in hydrogen peroxide, ozone and oxygen is
a) $\text{H}_2\text{O}_2 > \text{O}_3 > \text{O}_2$ b) $\text{O}_2 > \text{O}_3 > \text{H}_2\text{O}_2$ c) $\text{O}_2 > \text{H}_2\text{O}_2 > \text{O}_3$ d) $\text{O}_3 > \text{O}_2 > \text{H}_2\text{O}_2$
- Which one of the following is diamagnetic?
a) O_2 b) O_2^{2-} c) O_2^+ d) None of these
- Bond order of a species is 2.5 and the number of electrons in its bonding molecular orbital is found to be 8
The no. of electrons in its antibonding molecular orbital is
a) three b) four c) Zero d) can not be calculated from the given information.
- Shape and hybridisation of IF_5 are
a) Trigonal bipyramidal, Sp^3d^2 b) Trigonal bipyramidal, Sp^3d
c) Square pyramidal, Sp^3d^2 d) Octahedral, Sp^3d^2

17. Pick out the incorrect statement from the following
- sp^3 hybrid orbitals are equivalent and are at an angle of $109^\circ 28'$ with each other
 - dsp^2 hybrid orbitals are equivalent and bond angle between any two of them is 90°
 - All five sp^3d hybrid orbitals are not equivalent out of these five sp^3d hybrid orbitals, three are at an angle of 120° , remainir two are perpendicular to the plane containing the other three
 - none of these
18. The molecules having same hybridisation, shape and number of lone pairs of electons are
- SeF_4 , XeO_2F_2
 - SF_4 , XeF_2
 - $XeOF_4$, TeF_4
 - $SeCl_4$, XeF_4
19. In which of the following molecules / ions BF_3 , NO_2^- , H_2O the central atom is sp^2 hybridised?
- NH_2^- and H_2O
 - NO_2^- and H_2O
 - BF_3 and NO_2^-
 - BF_3 and NH_2^-
20. Some of the following properties of two species, NO_3^- and H_3O^+ are described below. which one of them is correct?
- dissimilar in hybridisation for the central atom with different structure.
 - isostructural with same hybridisation for the Central atom.
 - different hybridiration for the central atom with same structure
 - none of these
21. The types of hybridiration on the five carbon atom from right to left in the, 2,3 pentadiene.
- sp^3 , sp^2 , sp , sp^2 , sp^3
 - sp^3 , sp , sp , sp , sp^3
 - sp^2 , sp , sp^2 , sp^2 , sp^3
 - sp^3 , sp^3 , sp^2 , sp^3 , sp^3
22. XeF_2 is isostructural with
- $SbCl_2$
 - $BaCl_2$
 - TeF_2
 - ICl_2^-
23. The percentage of s-character of the hybrid orbitals in methane, ethane, ethene and ethyne are respectively
- 25, 25, 33.3, 50
 - 50, 50, 33.3, 25
 - 50, 25, 33.3, 50
 - 50, 25, 25, 50
24. Of the following molecules, which have shape similar to carbon dioxide?
- $SnCl_2$
 - NO_2
 - C_2H_2
 - All of these.
25. According to VSEPR theory, the repulsion between different parts of electrons obey the order.
- $l.p - l.p > b.p - b.p > l.p - b.p$
 - $b.p - b.p > b.p - l.p > l.p - b.p$
 - $l.p - l.p > b.p - l.p > b.p - b.p$
 - $b.p - b.p > l.p - l.p > b.p - l.p$
26. Shape of ClF_3 is
- Planar triangular
 - Pyramidal
 - 'T' Shaped
 - none of these
27. Non- Zero dipole moment is shown by
- CO_2
 - p-dichlorobenzene
 - carbontetrachloride
 - water.
28. Which of the following conditions is not correct for resonating structures?
- the contributing structure must have the same number of unpaired electrons
 - the contributing structures should have similar energies
 - the resonance hybrid should have higher energy than any of the contributing structure.
 - none of these
29. Among the following, the compound that contains, ionic, covalent and Co-ordinate linkage is
- NH_4Cl
 - NH_3
 - $NaCl$
 - none of these
30. CaO and $NaCl$ have the same crystal structure and approximately the same radii. If U is the lattice energy of $NaCl$, the approximate lattice energy of CaO is
- U
 - $2U$
 - $U/2$
 - $4U$

Additional:

- Lattice energy of an ionic compound depends upon:
 - Charge on the ions only
 - Size of the ions only
 - Packing of the ions only
 - Charge and size of the ion
- Which of the following has the highest ionic character?
 - $MgCl_2$
 - $CaCl_2$
 - $BaCl_2$
 - $BeCl_2$
- Among the following the maximum covalent character is shown by the compound
 - $FeCl_2$
 - $SnCl_2$
 - $AlCl_3$
 - $MgCl_2$

4. The electronegativity of H and Cl are 2.1 and 3.0 respectively. The correct statement (s) about the nature of HCl is/are: a) 17 % ionic b) 83 % ionic c) 50 % ionic d) 100 % ionic
5. Pick out the molecule which has zero dipole moment a) NH_3 b) H_2O c) BCl_3 d) SO_2
6. Of the following molecules, the one, which has permanent dipole moment is: a) SiF_4 b) BF_3 c) PF_3 d) PF_5
7. The hybridization of carbon atoms in $\text{C} - \text{C}$ single bond of $\text{H} - \text{C} \equiv \text{C} = \text{CH} = \text{CH}_2$ is
a) $\text{sp}^3 - \text{sp}^3$ b) $\text{sp}^2 - \text{sp}$ c) $\text{sp} - \text{sp}^2$ d) $\text{sp}^3 - \text{sp}$
8. The strength of bonds by overlapping of atomic orbitals is in the order
a) $s - s > s - p > p - p$ b) $s - s < p - p < s - p$ c) $s - p < s - s < p - p$ d) $p - p < s - s < s - p$
9. The structure of IF_7 is a) square pyramidal b) trigonal bipyramidal
c) octahedral d) pentagonal bipyramidal
10. The structure of XeOF_4 is a) tetrahedral b) square pyramidal c) square planar d) octahedral
11. Which one of the following compounds has sp^2 hybridization? a) CO_2 b) SO_2 c) NO_2^+ d) CO
12. The shape of XeO_2F_2 molecule is a) Trigonal bipyramidal b) square planar c) tetrahedral d) see - saw
13. According to MO theory, a) O_2^+ is paramagnetic and bond order is greater than O_2
b) O_2^+ is paramagnetic and bond order is less than O_2
c) O_2^+ is diamagnetic and bond order is less than O_2
d) O_2^+ is diamagnetic and bond order is more than O_2
14. Bond order of O_2 , O_2^+ , O_2^- and O_2^{2-} is in order
a) $\text{O}_2^- < \text{O}_2^{2-} < \text{O}_2 < \text{O}_2^+$ b) $\text{O}_2^{2-} < \text{O}_2^- < \text{O}_2 < \text{O}_2^+$ c) $\text{O}_2^+ < \text{O}_2 < \text{O}_2^- < \text{O}_2^{2-}$ d) $\text{O}_2 < \text{O}_2^+ < \text{O}_2^- < \text{O}_2^{2-}$
15. Which of the following is paramagnetic? a) O_2^- b) CN^- c) CO d) NO^+
16. Which of the following has zero dipole moment? a) CH_2Cl_2 b) CH_4 c) NH_3 d) PH_3
17. The isoelectronic pair is a) Cl_2 , ICl_2^- b) ICl_2^- , ClO_2 c) IF_2^+ , I_3^- d) ClO_2^- , ClF_2^+
18. The bond order is maximum in a) O_2 b) O_2^- c) O_2^+ d) O_2^{2-}
19. The number of ionic, covalent, and coordinate bond NH_4Cl are respectively
a) 1, 3 and 1 b) 1, 3 and 2 c) 1, 2 and 3 d) 1, 1 and 3
20. The hybridization of orbitals of N atom in NO_3^- , NO_3^+ and NH_4^+ are respectively.
a) sp , sp^2 , sp^3 b) sp^2 , sp , sp^3 c) sp , sp^3 , sp^2 d) sp^2 , sp^3 , sp

11. Fundamentals of Organic Chemistry

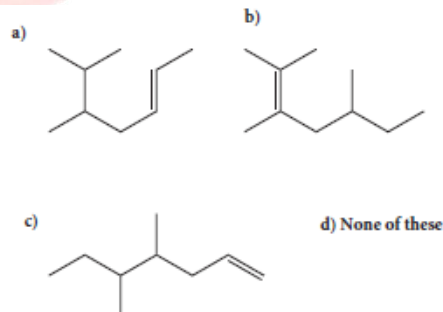
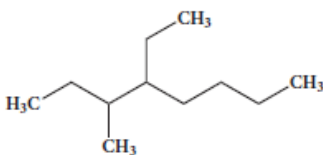
EVALUATION:

I. Choose the best answer.

1. Select the molecule which has only one π bond.
a) $\text{CH}_3 - \text{CH} = \text{CH} - \text{CH}_3$ b) $\text{CH}_3 - \text{CH} = \text{CH} - \text{CHO}$ c) $\text{CH}_3 - \text{CH} = \text{CH} - \text{COOH}$ d) All of these
2. In the hydrocarbon $\overset{7}{\text{CH}_3} - \overset{6}{\text{CH}_2} - \overset{5}{\text{CH}} = \overset{4}{\text{CH}} - \overset{3}{\text{CH}_2} - \overset{2}{\text{C}} = \overset{1}{\text{CH}}$ state of hybridisation of carbon 1,2,3,4 and 7 are in the following sequence. a) sp , sp , sp^3 , sp^2 , sp^3 b) sp^2 , sp , sp^3 , sp^2 , sp^3 c) sp , sp , sp^2 , sp , sp^3 d) none of these
3. The general formula for alkadiene is a) C_nH_{2n} b) $\text{C}_n\text{H}_{2n-1}$ c) $\text{C}_n\text{H}_{2n-2}$ d) C_nH_{n-2}
4. Structure of the compound whose IUPAC name is 5,6 - dimethylhept - 2 - ene is

5. The IUPAC name of the Compound is

- a) 2,3 - Dimethylheptane
b) 3- Methyl -4- ethyloctane
c) 5-ethyl -6-methyloctane
d) 4-Ethyl -3 - methyloctane



6. Which one of the following names does not fit a real name?

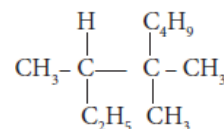
- a) 3 - Methyl -3-hexanone b) 4-Methyl -3- hexanone
c) 3- Methyl -3- hexanol d) 2- Methyl cyclo hexanone.

7. The IUPAC name of the compound $\text{CH}_3\text{-CH=CH-C}\equiv\text{CH}$ is

- a) Pent - 4 - yn-2-ene b) Pent -3-en-1-yne c) pent - 2- en - 4 - yne d) Pent - 1 - yn -3 -ene

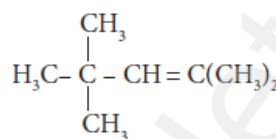
8. IUPAC name

- a) 3,4,4 - Trimethylheptane b) 2 - Ethyl -3, 3- dimethyl heptanes
c) 3, 4,4 - Trimethyloctane d) 2 - Butyl -2 -methyl - 3 - ethyl-butane.



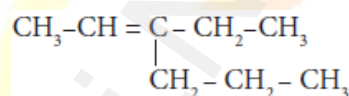
9. The IUPAC name

- a) 2,4,4 - Trimethylpent -2-ene b) 2,4,4 - Trimethylpent -3-ene
c) 2,2,4 - Trimethylpent -3-ene d) 2,2,4 - Trimethylpent -2-ene



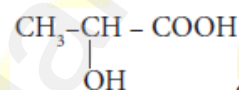
10. The IUPAC name of the compound

- a) 3 - Ethyl -2- hexene b) 3 - Propyl -3- hexene
c) 4 - Ethyl - 4 - hexene d) 3 - Propyl -2-hexene

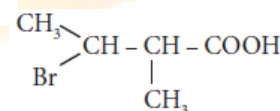


11. The IUPAC name of the compound

- a) 2 - Hydroxypropionic acid b) 2 - Hydroxy Propanoic acid



- c) Propan - 2- ol -1 - oic acid d) 1 - Carboxyethanol.

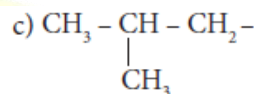
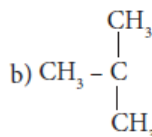


12. The IUPAC name

- a) 2 - Bromo -3 - methyl butanoic acid b) 2-methyl- 3- bromobutanoic acid
c) 3 - Bromo - 2 - methylbutanoic acid d) 3 - Bromo - 2, 3 - dimethyl propanoic acid.

13. The structure of isobutyl group in an organic compound is

- a) $\text{CH}_3\text{-CH}_2\text{-CH}_2\text{-CH}_2\text{-}$



14. The number of stereoisomers of 1, 2 - dihydroxy cyclopentane

- a) 1 b) 2 c) 3 d) 4

15. Which of the following is optically active?

- a) 3 - Chloropentane b) 2 Chloro propane c) Meso - tartaric acid d) Glucose
16. The isomer of ethanol is a) acetaldehyde b) dimethylether c) acetone d) methyl carbinol

17. How many cyclic and acyclic isomers are possible for the molecular formula $\text{C}_3\text{H}_6\text{O}$?

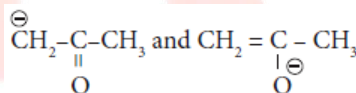
- a) 4 b) 5 c) 9 d) 10

18. Which one of the following shows functional isomerism?

- a) ethylene b) Propane c) ethanol d) CH_2Cl_2

19. Find the isomers

- a) resonating structure b) tautomers
c) Optical isomers d) Conformers.



20. Nitrogen detection in an organic compound is carried out by Lassaigne's test. The blue colour formed is due to the formation of.

- a) $\text{Fe}_3[\text{Fe}(\text{CN})_6]_2$ b) $\text{Fe}_4[\text{Fe}(\text{CN})_6]_3$ c) $\text{Fe}_4[\text{Fe}(\text{CN})_6]_2$ d) $\text{Fe}_3[\text{Fe}(\text{CN})_6]_3$

21. Lassaigne's test for the detection of nitrogen fails in

- a) $\text{H}_2\text{N-CO-NH.NH}_2.\text{HCl}$ b) $\text{NH}_2\text{-NH}_2.\text{HCl}$
c) $\text{C}_6\text{H}_5\text{-NH-NH}_2.\text{HCl}$ d) $\text{C}_6\text{H}_5\text{CONH}_2$

22. Connect pair of compounds which give blue colouration / precipitate and white precipitate respectively, when their Lassaigne's test is separately done.
- a) $\text{NH}_2\text{NH}_2\text{HCl}$ and $\text{ClCH}_2\text{-CHO}$ b) NH_2CSNH_2 and $\text{CH}_3\text{-CH}_2\text{Cl}$
 c) $\text{NH}_2\text{CH}_2\text{COOH}$ and NH_2CONH_2 d) $\text{C}_6\text{H}_5\text{NH}_2$ and $\text{ClCH}_2\text{-CHO}$.
23. Sodium nitropruside reacts with sulphide ion to give a purple colour due to the formation of
- a) $[\text{Fe}(\text{CN})_5\text{NO}]^{3-}$ b) $[\text{Fe}(\text{NO})_5\text{CN}]^+$ c) $[\text{Fe}(\text{CN})_5\text{NOS}]^{4-}$ d) $[\text{Fe}(\text{CN})_5\text{NOS}]^{3-}$
24. An organic Compound weighing 0.15g gave on carius estimation, 0.12g of silver bromide. The percentage of bromine in the Compound will be close to
- a) 46% b) 34% c) 3.4% d) 4.6%
25. A sample of 0.5g of an organic compound was treated according to Kjeldahl's method. The ammonia evolved was absorbed in 50mL of 0.5M H_2SO_4 . The remaining acid after neutralisation by ammonia consumed 80mL of 0.5 MNaOH, The percentage of nitrogen in the organic compound is.
- a) 14% b) 28% c) 42% d) 56%
26. In an organic compound, phosphorus is estimated as
- a) $\text{Mg}_2\text{P}_2\text{O}_7$ b) $\text{Mg}_3(\text{PO}_4)_2$ c) H_3PO_4 d) P_2O_5
27. Ortho and para-nitro phenol can be separated by
- a) azeotropic distillation b) destructive distillation
 c) steam distillation d) cannot be separated
28. The purity of an organic compound is determined by
- a) Chromatography b) Crystallisation c) melting or boiling point d) both (a) and (c)
29. A liquid which decomposes at its boiling point can be purified by
- a) distillation at atmospheric pressure b) distillation under reduced pressure
 c) fractional distillation d) steam distillation.
30. Assertion: 3- carbethoxy -2- butenoic acid.
- $$\begin{array}{c} \text{CH}_3 - \text{C} = \text{CH} - \text{COOH} \\ | \\ \text{COOC}_2\text{H}_5 \end{array}$$
- Reason: The principal functional group gets lowest number followed by double bond (or) triple bond.
- (a) both the assertion and reason are true and the reason is the correct explanation of assertion.
 (b) both assertion and reason are true and the reason is not the correct explanation of assertion.
 (c) assertion is true but reason is false (d) both the assertion and reason are false.
- Additional :**
1. The first organic compound synthesized in the laboratory from an inorganic compound is
- a) NH_4NCO b) $\text{NH}_2\text{-CO-NH}_2$ c) CH_3COOH d) CH_4
2. Diethyl ether and n - propyl methyl ether are
- a) Metamers b) Chain isomers c) Geometrical isomers d) Position isomers
3. The total number of structural isomers for the compound of the formula $\text{C}_4\text{H}_{10}\text{O}$ is
- a) 7 b) 6 c) 4 d) 3
4. According to Huckel's rule a compound, is said to be aromatic if it contains
- a) $4n$ bonds b) $4n$ atoms c) $(4n + 2)$ atoms d) $(4n + 2)$ π electrons
5. IUPAC name of $\text{CH} \equiv \text{C} - \text{CH} = \text{CH}_2$ is
- a) but - 3 - ene - 1 - yne b) but - 1 - ene - 3 - yne
 c) but - 1 - yne - 3 - ene d) but - 3 - yne - 1 - ene
6. The IUPAC name of the Compound $\text{CH}_3\text{-CH(OH)-COOH}$ is
- a) Lactic acid b) 2 - Hydroxy propanoic acid c) 3 - Hydroxy propanoic acid d) Carboxy propanol
7. Tautomerism is shown by
- a) $\text{R-C} \equiv \text{N}$ b) R-NO_2 c) R-OH d) R-COOH
8. The method of separation of enantiomers from racemic mixture is known as
- a) inversion b) recemisation c) resolution d) asymmetric synthesis
9. Which of the following exhibit cis - trans isomerism
- a) propene b) 1 - butane c) 2 - butane d) benzene

10. d – tartaric acid and l – tartaric acid can be separated by
 a) Salt formation b) Fractional distillation c) Fractional crystallization d) Chromatography
11. Paper chromatography is a) Adsorption chromatography b) partition chromatography
 c) Ion exchange chromatography d) all of these
12. Simple distillation can be used to separate liquids which differ in their boiling points at least by
 a) 5°C b) 10°C c) 40 – 50°C d) 100°C
13. A very common adsorbent used in column chromatography is
 a) Powdered charcoal b) Alumina c) Chalk d) Sodium carbonate
14. Lassaigne's test is used in qualitative analysis to detect
 a) Nitrogen b) Sulphur c) Chlorine d) All of these
15. The presence of halogen in an organic compound is detected by
 a) Iodoform test b) Silver nitrate test c) Beilstein's test d) Million's test
16. In Kjeldahl's method, the nitrogen presence is estimated as
 a) N₂ b) NH₃ c) NO₂ d) N₂O₃
17. Ortho and para nitro phenols can be separated by
 a) crystallization b) distillation c) sublimation d) solvent extraction

12. Basic Concepts of Organic Reactions

EVALUATION:

I. Choose the best answer.

1. For the following reactions
 Which of the following statement is correct?
 (a) (A) is elimination, (B) and (C) are substitution
 (b) (A) is substitution, (B) and (C) are elimination
 (c) (A) and (B) are elimination and (C) is addition reaction
 (d) (A) is elimination, B is substitution and (C) is addition reaction.
2. What is the hybridisation state of benzyl carbonium ion? (a) sp² (b) sp²d (c) sp³ (d) sp²d
3. Decreasing order of nucleophilicity is
 (a) OH⁻ > NH₂⁻ > -OCH₃ > RNH₂ (b) NH₂⁻ > OH⁻ > -OCH₃ > RNH₂
 (c) NH₂⁻ > CH₃O⁻ > OH⁻ > RNH₂ (d) CH₃O⁻ > NH₂⁻ > OH⁻ > RNH₂
4. Which of the following species is not electrophilic in nature?
 (a) Cl⁺ (b) BH₃ (c) H₃O⁺ (d) NO₂⁺
5. Homolytic fission of covalent bond leads to the formation of
 (a) electrophile (b) nucleophile (c) Carbo cation (d) free radical
6. Hyper Conjugation is also known as
 (a) no bond resonance (b) Baker - nathan effect (c) both (a) and (b) (d) none of these
7. Which of the group has highest +I effect?
 (a) CH₃- (b) CH₃-CH₂- (c) (CH₃)₂-CH- (d) (CH₃)₃-C-
8. Which of the following species does not exert a resonance effect?
 (a) C₆H₅OH (b) C₆H₅Cl (c) C₆H₅NH₂ (d) C₆H₅NH₃⁺
9. -I effect is shown by (a) -Cl (b) -Br (c) both (a) and (b) (d) -CH₃
10. Which of the following carbocation will be most stable?
 (a) Ph₃C⁺ (b) CH₃-C⁺H₂-
 (c) (CH₃)₂-C⁺H (d) CH₂=CH-C⁺H₂

11. Assertion: Tertiary Carbocations are generally formed more easily than primary Carbocations ions.

Reason: Hyper conjugation as well as inductive effect due to additional alkyl group stabilize tertiary carbonium ions.

(a) both assertion and reason are true and reason is the correct explanation of assertion.

(b) both assertion and reason are true but reason is not the correct explanation of assertion.

(c) Assertion is true but reason is false

(d) Both assertion and reason are false

12. Heterolytic fission of C–C bond results in the formation of

(a) free radical (b) Carbanion (c) Carbocation (d) Carbanion and Carbocation

13. Which of the following represent a set of nucleophiles?

(a) BF_3 , H_2O , NH_2^- (b) AlCl_3 , BF_3 , NH_3 (c) CN^- , RCH_2^- , ROH (d) H^+ , RNH_3^+ , $:\text{CCl}_2$

14. Which of the following species does not act as a nucleophile?

(a) ROH (b) ROR (c) PCl_3 (d) BF_3

15. The geometrical shape of carbocation is (a) Linear (b) tetrahedral (c) Planar (d) Pyramidal

Additional:

1. The shape of carbonium ion is (a) Planar (b) Linear (c) Pyramidal (d) Tetrahedral

2. Which of the following species is paramagnetic

(a) A carbonium ion (b) A free radical (c) A carbanion (d) All of these

3. In carbonium ion the carbon bearing the positive charge is

(a) sp hybridized (b) sp^2 hybridized (c) sp^3 hybridized (d) un hybridized

4. Electrophiles are (a) Lewis bases (b) Lewis acids (c) Amphoteric (d) All of these

5. Electrophiles are (a) Electron loving species (b) Electron hating species

(c) Nucleus loving reagents (d) Nucleus hating reagents

6. The electromeric effect in organic compounds is a

(a) Temporary effect (b) Permanent effect (c) Temporary or permanent effect (d) All of the above

7. + I effect is shown by (a) $-\text{NO}_2$ (b) $-\text{Cl}$ (c) $-\text{Br}$ (d) $-\text{CH}_3$

8. The reaction intermediate produced by homolytic cleavage of bond is called

(a) carbocations (b) carbanions (c) free radicals (d) carbenes

9. The most stable carbonium ion is (a) Methyl carbonium ion (b) Primary carbonium ion

(c) Secondary carbonium ion (d) Tertiary carbonium ion

10. Alkenes readily undergo (a) Substitution reactions (b) Addition reactions

(c) Elimination reactions (d) Rearrangement reactions

11. Nitration of benzene is (a) nucleophilic substitution (b) nucleophilic addition

(c) electrophilic substitution (d) free radical substitution

12. Which is most stable carbocation? (a) n – propyl cation (b) iso – propyl cation

(c) Ethyl cation (d) Triphenylmethyl cation

13. Hydrocarbons

EVALUATION:

I. Choose the best answer.

1. The correct statement regarding the comparison of staggered and eclipsed conformations of ethane, is

a) the eclipsed conformation of ethane is more stable than staggered conformation even though the eclipsed conformation has torsional strain.

- b) the staggered conformation of ethane is more stable than eclipsed conformation, because staggered conformation has no torsional strain.
- c) the staggered conformation of ethane is less stable than eclipsed conformation, because staggered conformation has torsional strain.
- d) the staggered conformation of ethane is less stable than eclipsed conformation, because staggered conformation has no torsional strain.

2. $C_2H_5Br + 2Na \xrightarrow{\text{dry ether}} C_4H_{10} + 2NaBr$ The above reaction is an example of which of the following

- a) Reimer Tiemann reaction b) Wurtz reaction c) Aldol condensation d) Hoffmann reaction

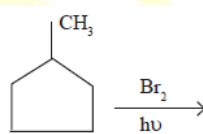
3. An alkyl bromide (A) reacts with sodium in ether to form 4, 5- diethyloctane, the compound (A) is

- a) $CH_3(CH_2)_3Br$
 b) $CH_3(CH_2)_5Br$
 c) $CH_3(CH_2)_3CH(Br)CH_3$
 d) $CH_3 - (CH_2)_2 - CH(Br) - CH_2 - CH_3$

4. The C - H bond and C - C bond in ethane are formed by which of the following

- types of overlap a) $sp^3 - s$ and $sp^3 - sp^3$ b) $sp^2 - s$ and $sp^2 - sp^2$
 c) $sp - sp$ and $sp - sp$ d) $p - s$ and $p - p$

5. In the following reaction,



The major product obtained is

- a) b)
 c) d)

6. Which of the following is optically active

- a) 2 - methyl pentane b) citric acid c) Glycerol d) none of these

7. The compounds formed at anode in the electrolysis of an aqueous solution of potassium acetate are

- a) CH_4 and H_2 b) CH_4 and CO_2 c) C_2H_6 and CO_2 d) C_2H_4 and Cl_2

8. The general formula for cyclo alkanes

- a) C_nH_n b) C_nH_{2n} c) C_nH_{2n-2} d) C_nH_{2n+2}

9. The compound that will react most readily with gaseous bromine has the formula

- a) C_3H_6 b) C_2H_2 c) C_4H_{10} d) C_2H_4

10. Which of the following compounds shall not produce propene by reaction with HBr followed by elimination (or) only direct elimination reaction

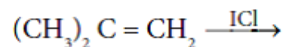
- a) b) $CH_3 - CH_2 - CH_2 - OH$
 c) $H_2C = C = O$ d) $CH_3 - CH_2 - CH_2Br$

11. Which among the following alkenes on reductive ozonolysis produces only propanone ?

- a) 2 - Methyl propene b) 2 - Methyl but - 2 - ene
 c) 2, 3 - Dimethyl but - 1 - ene d) 2, 3 - Dimethyl but - 2 - ene

12. The major product formed when 2 - bromo - 2 - methyl butane is refluxed with ethanolic KOH is

- a) 2 - methylbut - 2 - ene b) 2 - methyl butan - 1 - ol
 c) 2 - methyl but - 1 - ene d) 2 - methyl butan - 2 - ol

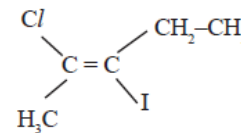


13. Major product of the below mentioned reaction is,

- a) 2-chloro -1- iodo - 2 - methyl propane b) 1-chloro-2-iodo-2-methylpropane
c) 1,2 - dichloro - 2 - methyl propane d) 1, 2 - diiodo - 2 - methyl propane

14. The IUPAC name of the following compound is

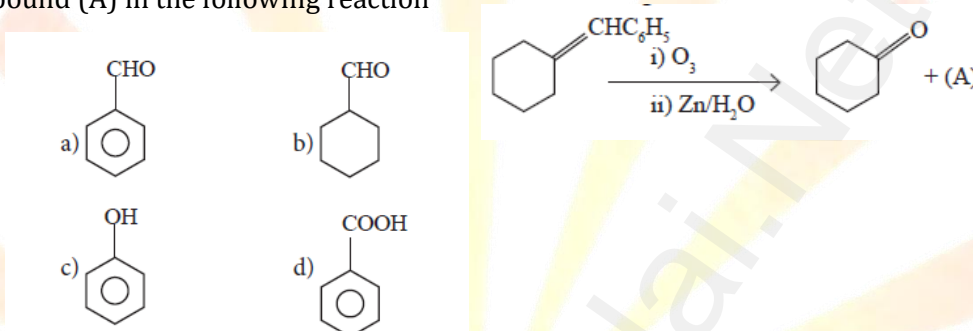
- a) trans-2-chloro-3-iodo - 2 - pentene b) cis-3 - iodo - 4 - chloro - 3 - pentene
c) trans-3-iodo-4-chloro - 3 - pentene d) cis-2 - chloro - 3 - iodo - 2 - pentene



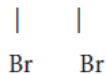
15. Cis - 2 - butene and trans - 2 - butene are

- a) conformational isomers b) structural isomers c) configurational isomers d) optical isomers

16. Identify the compound (A) in the following reaction



17. $\text{CH}_2 - \text{CH}_2 \xrightarrow{\text{(A)}} \text{CH} \equiv \text{CH}$, where A is,



- a) Zn b) Conc. H_2SO_4 c) alc. KOH d) dil H_2SO_4

18. Consider the nitration of benzene using mixed con H_2SO_4 and HNO_3 if a large quantity of KHSO_4 is added to the mixture, the rate of nitration will be

- a) unchanged b) doubled c) faster d) slower

19. In which of the following molecules, all atoms are co-planar

- a) c)
b) d) both (a) and (b)

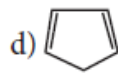
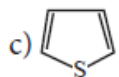
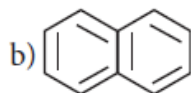
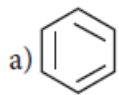
20. Propyne on passing through red hot iron tube gives

- a) b)
c) d) none of these

21. $\xrightarrow{\text{HCl}}$ (A) is

- a) c) both (a) and (b)
b) d)

22. Which one of the following is non aromatic ?



23. Which of the following compounds will not undergo Friedal - crafts reaction easily ?

- a) Nitro benzene b) Toluene c) Cumene d) Xylene

24. Some meta-directing substituents in aromatic substitution are given. Which one is most deactivating ?

- a) - COOH b) - NO₂ c) - C ≡ N d) - SO₃H

25. Which of the following can be used as the halide component for friedal - crafts reaction ?

- a) Chloro benzene b) Bromo benzene c) chloro ethene d) isopropyl chloride

26. An alkane is obtained by decarboxylation of sodium propionate. Same alkane can be prepared by

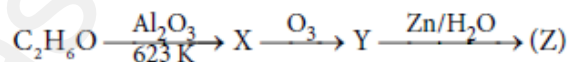
- a) Catalytic hydrogenation of propene b) action of sodium metal on iodomethane
c) reduction of 1 - chloro propane d) reduction of bromomethane

27. Which of the following is aliphatic saturated hydrocarbon

- a) C₈H₁₈ b) C₉H₁₈ c) C₈H₁₄ d) All of these

28. Identify the compound 'Z' in the following reaction

- a) Formaldehyde b) Acetaldehyde
c) Formic acid d) none of these



29. Peroxide effect (Kharasch effect) can be studied in case of

- a) Oct - 4 - ene b) hex - 3 - ene c) pent - 1 - ene d) but - 2 - ene

30. 2 - butyne on chlorination gives

- a) 1 - chloro butane b) 1, 2 - dichloro butane
c) 1, 1, 2, 2 - tetrachlorobutane d) 2, 2, 3, 3 - tetra chloro butane

Additional:

1. The gas supplied in cylinders for cooking is

- a) marsh gas b) LPG
c) mixture CH₄ and C₂H₆ d) mixture of ethane and propane

2. Adam's catalyst is:

- a) platinum metal b) palladium c) nickel metal d) PtO₂

3. The most stable conformation of ethane is

- a) Eclipsed b) Skew c) Staggered d) All are equally stable

4. In Wurtz reaction, n - hexane is obtained from

- a) n - propyl chloride b) n - butyl chloride c) Ethyl chloride d) isopropyl chloride

5. The alkene that exhibits geometrical isomerism is

- a) propene b) 2 - methyl propene c) 2 - butane d) 2 - methyl - 2 - butane

6. Baeyer's reagent is

- a) Aqueous bromine solution b) Neutral permanganate solution
c) Acidified permanganate solution d) Alkaline potassium permanganate solution

7. Polytetrafluoroethylene is commercially known as

- a) Teflon b) Freon c) Lewisite d) Westron

8. The peroxide effect involves

- a) Ionic mechanism b) Free - radical mechanism
c) Heterolytic fission of double bond d) Homolytic fission of double bond

9. The number of possible alkynes with molecular formula C₅H₈ is

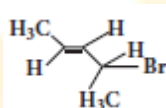
- a) 3 b) 4 c) 5 d) 6

10. Westron is the solvent obtained by the reaction of chlorine with
 a) Ethylene b) Ethyne c) Ethane d) Methane
11. Coal tar is obtained as a by product during
 a) Destructive distillation of wood b) Destructive distillation of coal
 c) Destructive distillation of bones d) steam distillation of light oil
12. Chemical name of the insecticide gammaxene is
 a) DDT b) Benzene hexa chloride c) Chloral d) Hexa chloro ethane
13. Benzene is purified by a) distillation b) fractional distillation c) Evaporation d) sublimation
14. Which of the following is not meta directing group? a) $-\text{SO}_3\text{H}$ b) $-\text{NO}_2$ c) $-\text{CN}$ d) $-\text{NH}_2$

14. Haloalkanes and Haloarenes

EVALUATION:

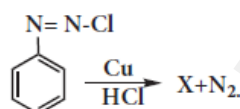
1. The IUPAC name of



- a) 2-Bromo pent-3-ene b) 4-Bromo pent-2-ene c) 2-Bromo pent-4-ene d) 4-Bromo pent-1-ene
2. Of the following compounds, which has the highest boiling point?
 a) n-Butyl chloride b) Isobutyl chloride c) t-Butyl chloride d) n-propyl chloride
3. Arrange the following compounds in increasing order of their density
 A) CCl_4 B) CHCl_3 C) CH_2Cl_2 D) CH_3Cl
 a) $\text{D} < \text{C} < \text{B} < \text{A}$ b) $\text{C} > \text{B} > \text{A} > \text{D}$ c) $\text{A} < \text{B} < \text{C} < \text{D}$ d) $\text{C} > \text{A} > \text{B} > \text{D}$
4. With respect to the position of $-\text{Cl}$ in the compound $\text{CH}_3 - \text{CH} = \text{CH} - \text{CH}_2 - \text{Cl}$, it is classified as
 a) Vinyl b) Allyl c) Secondary d) Aralkyl
5. What should be the correct IUPAC name of diethyl chloromethane?
 a) 3 - Chloro pentane b) 1-Chloropentane
 c) 1-Chloro-1, 1, diethyl methane d) 1 -Chloro-1-ethyl propane
6. C - X bond is strongest in a) Chloromethane b) Iodomethane c) Bromomethane d) Fluoromethane

7.

In the reaction

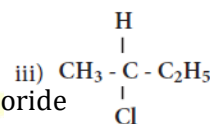
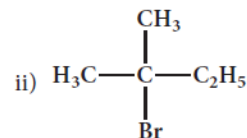
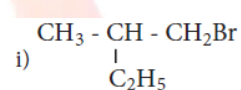


X is _____

- a) b) c) d)

8. Which of the following compounds will give racemic mixture on nucleophilic substitution by OH^- ion?

- a) (i) b) (ii) and (iii) c) (iii) d) (i) and (ii)
9. The treatment of ethyl formate with excess of RMgX gives
 a) $\text{R} - \text{C} - \text{R} - \text{O}$ b) $\text{R} - \text{CH} - \text{R} - \text{OH}$ c) $\text{R} - \text{CHO}$ d) $\text{R} - \text{O} - \text{R}$



10. Benzene reacts with Cl_2 in the presence of FeCl_3 and in absence of sunlight to form
 a) Chlorobenzene b) Benzyl chloride c) Benzal chloride d) Benzene hexachloride
11. The name of $\text{C}_2\text{F}_4\text{Cl}_2$ is _____ a) Freon-112 b) Freon-113 c) Freon-114 d) Freon-115
12. Which of the following reagent is helpful to differentiate ethylene dichloride and ethylidene chloride?
 a) $\text{Zn} / \text{methanol}$ b) $\text{KOH} / \text{ethanol}$ c) aqueous KOH d) $\text{ZnCl}_2 / \text{Con HCl}$

13. Match the compounds given in Column I with suitable items given in Column II

Code

- a) A → 2 B → 4 C → 1 D → 3 b) A → 3 B → 2 C → 4 D → 1
c) A → 1 B → 2 C → 3 D → 4 d) A → 3 B → 1 C → 4 D → 2

| | Column I (Compound) | | Column II (Uses) |
|---|-----------------------|---|-------------------|
| A | Iodoform | 1 | Fire extinguisher |
| B | Carbon tetra chloride | 2 | Insecticide |
| C | CFC | 3 | Antiseptic |
| D | DDT | 4 | Refrigerants |

14. **Assertion:** In mono haloarenes, electrophilic substitution occurs at ortho and para positions.

Reason: Halogen atom is a ring deactivator

- (i) If both assertion and reason are true and reason is the correct explanation of assertion.
(ii) If both assertion and reason are true but reason is not the correct explanation of assertion.
(iii) If assertion is true but reason is false. (iv) If both assertion and reason are false.

15. Consider the reaction, $\text{CH}_3\text{CH}_2\text{CH}_2\text{Br} + \text{NaCN} \rightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{CN} + \text{NaBr}$

This reaction will be the fastest in a) ethanol b) methanol c) DMF (N, N' – dimethyl formamide) d) water

16. Freon-12 is manufactured from tetrachloro methane by

- a) Wurtz reaction b) Swarts reaction c) Haloform reaction d) Gattermann reaction

17. The most easily hydrolysed molecule under $\text{S}_{\text{N}}1$ condition is

- a) allyl chloride b) ethyl chloride c) isopropylchloride d) benzyl chloride

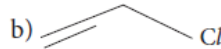
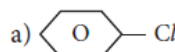
18. The carbo cation formed in $\text{S}_{\text{N}}1$ reaction of alkyl halide in the slow step is

- a) sp^3 hybridised b) sp^2 hybridised c) sp hybridised d) none of these

19. The major products obtained when chlorobenzene is nitrated with HNO_3 and conc. H_2SO_4

- a) 1-chloro-4-nitrobenzene b) 1-chloro-2-nitrobenzene
c) 1-chloro-3-nitrobenzene d) 1-chloro-1-nitrobenzene

20. Which one of the following is most reactive towards nucleophilic substitution reaction ?



21. Ethylidene chloride on treatment with aqueous KOH gives

- a) acetaldehyde b) ethyleneglycol c) formaldehyde d) glyoxal

22. The raw material for Rasching process

- a) chloro benzene b) phenol c) benzene d) anisole

23. Chloroform reacts with nitric acid to produce

- a) nitro toluene b) nitro glycerine c) chloropicrin d) chloropicric acid

24. acetone $\xrightarrow[\text{ii) } \text{H}_2\text{O} / \text{H}^+]{\text{i) } \text{CH}_3\text{MgI}}$ X, X is

- a) 2-propanol b) 2-methyl-2-propanol
c) 1-propanol d) acetanol

25. Silverpropionate when refluxed with Bromine in carbontetrachloride gives

- a) propionic acid b) chloro ethane c) bromo ethane d) chloro propane

Additional :

1. Grignard reagent is formed when alkyl halide reacts with which one of the following

- a) Mg in alcohol b) Mg in acid c) Mg in dry ether d) MgO

2. Which of the following is used as refrigerant?

- a) CH_3COCH_3 b) CCl_4 c) $\text{C}_2\text{H}_5\text{Cl}$ d) CF_4

3. In Dow's process the starting raw material is

- a) Phenol b) Chloro benzene
c) Aniline d) Diazobenzene

4. Chloro benzene is prepared commercially by
 a) Dow's process b) Decon's process c) Raschig process d) Etard's process
5. Chloro benzene on treatment with sodium in dry ether gives diphenyl. The name of the reaction is
 a) Fitting reaction b) Wurtz fittig reaction c) Wurtz reaction d) Sandmeyer reaction
6. The raw materials for the commercial manufacture of DDT are
 a) chloro benzene and chloroform b) chloro benzene and chloro methane
 c) chloro benzene and chloral d) chloro benzene and iodoform
7. Iodoform is used as a) anaesthetic b) antiseptic c) analgesic d) anti fibrin
8. Freon - 12 is a) CF_3Cl b) CHCl_2F c) CF_2Cl_2 d) DDT
9. The name of DDT a) p, p' - dichloro diphenyl trichloro ethane b) p, p' - dichloro diphenyl trichloro ethene
 c) p, p' - dichloro diphenyl ttnchloro benzene d) p, p' - tetra chloro ethane
10. With Zn - Cu couple and $\text{C}_2\text{H}_5\text{OH}$, ethyl Iodide reacts to give
 a) ethers b) diethyl ether c) Iodoform d) Ethane
11. The reaction of alkyl halide with benzene in presence of anhydrous AlCl_3 gives alkyl benzene the reaction is known as
 a) Friedel - craft's reaction b) Carbylamine reaction
 c) Gattermann reaction d) Wurtz reaction
12. S_N^2 reaction leads to a) inversion of configuration b) retention of configuration
 c) partial racemisation d) no racemisation
13. The order of reactivity of various alkyl halides toward $\text{S}_\text{N}1$ reaction is
 a) $3^\circ > 2^\circ > 1^\circ$ b) $1^\circ > 2^\circ > 3^\circ$ c) $3^\circ = 2^\circ = 1^\circ$ d) $1^\circ > 3^\circ > 2^\circ$
14. Chloro benzene can be prepared by reacting benzene diazonium chloride with
 a) HCl b) $\text{Cu}_2\text{Cl}_2 / \text{HCl}$ c) $\text{Cl}_2 / \text{AlCl}_3$ d) HNO_2

15. Environmental Chemistry

EVALUATION:

I. Choose the best answer.

1. The gaseous envelope around the earth is known as atmosphere. The region lying between an altitudes of 11-50 km is _____
 a) Troposphere b) Mesosphere c) Thermosphere d) stratosphere
2. Which of the following is natural and human disturbance in ecology?
 a) Forest fire b) Floods c) Acid rain d) Green house effect
3. Bhopal Gas Tragedy is a case of a) thermal pollution b) air pollution c) nuclear pollution d) land pollution
4. Haemoglobin of the blood forms carboxy haemoglobin with
 a) Carbon dioxide b) Carbon tetra chloride c) Carbon monoxide d) Carbonic acid
5. Which sequence for green house gases is based on GWP?
 a) $\text{CFC} > \text{N}_2\text{O} > \text{CO}_2 > \text{CH}_4$ b) $\text{CFC} > \text{CO}_2 > \text{N}_2\text{O} > \text{CH}_4$
 c) $\text{CFC} > \text{N}_2\text{O} > \text{CH}_4 > \text{CO}_2$ d) $\text{CFC} > \text{CH}_4 > \text{N}_2\text{O} > \text{CO}_2$
6. Photo chemical smog formed in congested metropolitan cities mainly consists of
 a) Ozone, SO_2 and hydrocarbons b) Ozone, PAN and NO_2
 c) PAN, smoke and SO_2 d) Hydrocarbons, SO_2 and CO_2
7. The pH of normal rain water is a) 6.5 b) 7.5 c) 5.6 d) 4.6
8. Ozone depletion will cause a) forest fires b) eutrophication c) bio magnification d) global warming
9. Identify the wrong statement in the following
 a) The clean water would have a BOD value of more than 5 ppm
 b) Greenhouse effect is also called as Global warming
 c) Minute solid particles in air is known as particulate pollutants
 d) Biosphere is the protective blanket of gases surrounding the earth

10. Living in the atmosphere of CO is dangerous because it
 a) Combines with O₂ present inside to form CO₂ b) Reduces organic matter of tissues
 c) Combines with haemoglobin and makes it incapable to absorb oxygen d) Diluted the blood
11. Release of oxides of nitrogen and hydrocarbons into the atmosphere by motor vehicles is prevented by using
 a) grit chamber b) scrubbers c) trickling filters d) catalytic convertors
12. Biochemical oxygen Demand value less than 5 ppm indicates a water sample to be
 a) highly polluted b) poor in dissolved oxygen c) rich in dissolved oxygen d) low COD
13. Match the List I with List II and select the correct answer using the code given below the lists

| List I | | List II | |
|--------|--------------------------|---------|-----------------|
| A | Depletion of ozone layer | 1 | CO ₂ |
| B | Acid rain | 2 | NO |
| C | Photochemical smog | 3 | SO ₂ |
| D | Green house effect | 4 | CFC |

| | A | B | C | D |
|---|---|---|---|---|
| a | 3 | 4 | 1 | 2 |
| b | 2 | 1 | 4 | 3 |
| c | 4 | 3 | 2 | 1 |
| d | 2 | 4 | 1 | 3 |

14.

| List I | | List II | |
|--------|------------------------------|---------|-------------------|
| A | Stone leprosy | 1 | CO |
| B | Biological magnification | 2 | Green house gases |
| C | Global warming | 3 | Acid rain |
| D | Combination with haemoglobin | 4 | DDT |

| | A | B | C | D |
|---|---|---|---|---|
| a | 1 | 2 | 3 | 4 |
| b | 3 | 4 | 2 | 1 |
| c | 2 | 3 | 4 | 1 |
| d | 4 | 2 | 1 | 3 |

The questions gives below consists of an assertion the reason. Choose the correct option out of the choices given below each question

- i) Both (A) and R are correct and (R) is the correct explanation of (A)
 ii) Both (A) and R are correct and (R) is not the correct explanation of (A)
 iii) Both (A) and R are not correct
 iv) (A) is correct but (R) is not correct

15. Assertion (A): If BOD level of water in a reservoir is more than 5 ppm it is highly polluted
 Reason(R) : High biological oxygen demand means high activity of bacteria in water
 a) i b) ii c) iii d) iv

16. Assertion (A): Excessive use of chlorinated pesticide causes soil and water pollution.
 Reason (R) : Such pesticides are non-biodegradable.
 a) i b) ii c) iii d) iv

17. Assertion (A): Oxygen plays a key role in the troposphere
 Reason (R): Troposphere is not responsible for all biological activities
 a) i b) ii c) iii d) iv

Additional :

1. The gas responsible for ozone depletion: (a) NO and freons (b) SO₂ (c) CO₂ (d) CO
2. In Antarctica ozone depletion is due to the formation of following compound
 (a) acrolein (b) peroxyacetyl nitrate (c) SO₂ and NO₂ (d) chlorine nitrate
3. Classical smog occurs in places of (a) excess SO₂ (b) low temperature (c) high temperature (d) excess NH₃

4. Which gas is responsible for 'Bhopal Gas Tragedy' in 1984?
(a) CO (b) Methyl isocyanate (c) SO₂ and NO₂ (d) Ethyl isocyanate
5. Which of the following is the coldest region of atmosphere
(a) Thermosphere (b) Mesosphere (c) Troposphere (d) Stratosphere
6. Formation of London smog takes place in (a) Winter during day time (b) summer during day time
(c) summer during morning time (d) winter during morning time
7. The substance which is not regarded as a pollutant? (a) NO₂ (b) CO₂ (c) O₃ (d) Hydrocarbons
8. Minamata disease of Japan is due to pollution of (a) Aresenic (b) Lead (c) Cynide (d) Mercury
9. Which is known as "Third poison of environment" and also creates 'Blue baby syndrome'
(a) Nitrate present in water (b) Phosphate and detergents found in water
(c) Cyanide (d) Pesticides
10. Green chemistry means such reactions which:
(a) produce colour during reactions (b) reduce the use and production of hazardous chemicals
(c) are related to the depletion of ozone layer (d) study the reactions in plants
11. What is DDT among the following? (a) Greenhouse gas (b) A fertilizer
(c) Biodegradable pollutant (d) Non - biodegradable pollutant
12. Black - foot disease is caused due to groundwater contaminated with excess of
(a) Nitrate (b) Fluoride (c) Arsenic (d) Sulphur
13. Which of the following metal is a water pollutant and causes sterility in human. being?
(a) As (b) Mn (c) Mg (d) Hg

--ALL THE BEST----



Time + Effort = Success

Thank God

XI CHEMISTRY VOLUME I & II**Answer Key****1. Basic Concepts of****Chemistry and Chemical****Calculations**

1. a) 40 ml of CO₂ gas
2. (d) 200 u
3. (c) assertion is true but
4. (b) oxygen
5. a) 102 g
6. c) 6.022×10^{20}
7. c) 16 %
8. c) 0.075
9. d) 1 mole of HCl
10. (c) $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow$
11. b) $\text{P}_4 + 3 \text{NaOH} + 3 \text{H}_2\text{O}$
12. (b) 52.7
13. d) 6.022×10^{24}
14. a) NO
15. a) 6.022×10^{23}
16. (c) $\text{S}_2\text{O}_4^{2-} < \text{SO}_3^{2-} < \text{S}_2\text{O}_6^{2-} <$
17. c) molar mass of ferrous
18. d) the mass of one mole
19. c) the ratio between B is
20. a) 3.59 g
21. b) 44 g mol⁻¹
22. c) both (a) & (b)
23. a) propene
24. a) relative atomic mass is
25. b) 6C^{12}

Additional :

1. (a) 6.023×10^{23} atoms of C
2. (c) 0.01
3. (b) 6.023×10^{23}
4. (c) 5
5. (a) H₂O
6. (a) Vapour Density $\times 2$
7. (d) 1
8. (d) 44
9. (b) -1
10. (b) 111 g mol⁻¹
11. (d) 46 amu
12. (d) 9

13. (c) 6 moles
14. (b) 0.5 mole
15. (c) 11.2 L
16. (b) $\text{Fe}_2\text{O}_3 \cdot x \text{H}_2\text{O}$
17. (b) + 6
18. (c) - 3
19. (c) Zn
20. (d) $4\text{H}_3\text{PO}_3 \rightarrow 3\text{H}_3\text{PO}_4 +$
21. (b) 142
22. (b) + 3
23. (d) Cl_2O_7 .

2. Quantum Mechanical**Model of Atom**

1. (c) 30
2. (c) $4.42 \times 10^{-18} \text{ J}$
3. (b) $\lambda_1 = 2\lambda_2$
4. (d) Stark effect
5. (b)
6. (d) $n=6$ to $n=5$
7. (a)
8. (c) $d_{z^2}, d_{x^2-y^2}$
9. (b) Spin quantum
10. (b) $[\text{Xe}]4f^7 6s^2 ; [\text{Xe}]4f^7$
11. (c) $4l+2$
12. (d) $\sqrt{6} h/2\pi$
13. (c) 2
14. (c)
15. (a) 9
16. (a) $ns \rightarrow (n-2)f \rightarrow (n-1)d \rightarrow np$
17. (b) (ii), (iv) & (v)
18. (b) 17
19. (a) Zero
20. (c) $1/2m \sqrt{h/\pi}$
21. (c) $6.6 \times 10^{-31} \text{ cm}$
22. (d) 0.4
23. (d) -9E
24. (a) $H\psi = E\psi$
25. (d) $\Delta E \cdot \Delta X \geq h/4\pi$

Additional :

1. (d) -3.4 eV atom⁻¹
2. (a) -1.51 eV atom⁻¹
3. (d) 32
4. (b) 2
5. (a) Zeeman effect.
6. (b) 3, 1, -1, + $\frac{1}{2}$
7. (c) nodal surface
8. (a) $s > p > d > f$
9. (a) 1, 0, 0, + $\frac{1}{2}$
10. (b)
- $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1$
11. (c) 3d

3. Periodic Classification of Elements

1. (d) bismuth
2. (b) AB₂
3. (d) f-block elements
4. (a) $\text{I} < \text{Br} < \text{Cl} < \text{F}$
5. (d) fluorine
6. (c) Aluminium
7. (b) $\text{Na} < \text{Al} < \text{Mg} < \text{Si} < \text{P}$
8. (a)
9. (d) $\text{Ca} < \text{Al} < \text{C} < \text{O} < \text{F}$
10. (c) $\text{Cl} > \text{F} > \text{Br} > \text{I}$
11. (d) Hydrogen
12. (c) Argon
13. (a) $\text{Y} > \text{Z} > \text{X} > \text{A}$
14. (c)
15. (a) $1s^2, 2s^2, 2p^6, 3s^1$
16. (c) Oxygen
17. (c) +527 kcal mol⁻¹
18. (a) $s > p > d > f$
19. (d) None of these
20. (b) 575 kJ mol⁻¹
21. (a)
22. (a) Generally increases
23. (d) Be and Al

Additional :

1. (d) Be, B, C
2. (d) Ga, Ge

3. (b) atomic numbers
 4. (a) Group 17
 5. (d) Sb
 6. (b) $(n-1)d^{1-10} ns^{0-2}$
 7. (a) representative
 8. (c) $ns^2 np^6$
 9. (d) Oxidation number
 10. (c) Atomic radius
 11. (c) Cl^-
 12. (a) Zr, Hf
 13. (b) Noble gases
 14. (c) Fluorine
 15. (d) Lithium
 16. (b) Li, Na
 17. (d) $K > Mg > Al > B$
 18. (d) All the above

4. Hydrogen

1. (c)
 2. (c) $CO + H_2$
 3. (b)
 4. (d) group one elements
 5. (c) $1p+2n$
 6. (a) Palladium, Vanadium
 7. (a)
 8. (a) 1.2 g
 9. (d) EDTA
 10. (c) $CaCl_2$
 11. (a) sodium aluminium
 12. (a)
 13. (c) $CrO(O_2)_2$
 14. (c) $5/2$
 15. (d) 8.4
 16. (d) sp^3 and sp^3
 17. (c) monobasic acid
 18. (a) tetrahedrally by 4-H
 19. (b) intra-molecular hydrogen
 20. (c) both (a) & (b)
 21. (c) amphoteric oxide

Additional :

1. (d) Platinum
 2. (c) tritium
 3. (c) 12.3

4. (c) $CO + H_2$
 5. (c) Magnesium and Calcium
 6. (d) chlorides and sulphates
 7. (a) $NaOAl_2O_3 \cdot xSiO_2 \cdot yH_2O$
 8. (d) Heavy water
 9. (c) EDTA
 10. (b) 30
 11. (c) Salicylaldehyde
 12. (d) Interstitial hydrides
 13. (b) Vanderwaals bond <
 14. (b) o-nitrophenol
 15. (c) $CH_4 \cdot 20 H_2O$

5. Alkali and Alkaline Earth

Metals

1. (c) Density: $Li < K < Na <$
 2. (a) Li^+ has minimum
 3. (d) none of these
 4. (b) Li
 5. (c) kerosene
 6. (a) superoxide and pa
 7. (c) Potassium carbonate
 8. (b) Magnesium
 9. (b) $MI < MBr < MCl < MF$
 10. (a) Castners process
 11. (c) $Ca(CN)_2$
 12. (a) $MgCl_2$
 13. (a) p-2, q-1, r-4, s-5, t-6,
 14. (d) both assertion and
 15. (a)
 16. (b) $MgCO_3 > CaCO_3 >$
 17. (c) Its salts are rarely
 18. (c) milk of lime
 19. (b) $NaHCO_3$
 20. (b) $Ca(OH)_2$
 21. (a) Ca^{2+} ions are not
 22. (b) CaF_2
 23. (a) $CaSO_4 \cdot 2H_2O$
 24. (b) $CaNCN$
 25. (d) Li_2CO_3

Additional :

1. (b) Cs
 2. (a) +1
 3. (d) lilac
 4. (d) K^+
 5. (b) strontium
 6. (b) Calcium hydroxide
 7. (d) amphoteric
 8. (d) (i) and (iii)
 9. (c) $Ca(OCl)_2$
 10. (b) dead burnt plast
 11. (a) $Na < K < Rb < Cs < Li$
 12. (a) KO_2

6. Gaseous State

1. (d) at high pressure the
 2. (d) inversely proportion
 3. (c)
 4. (b) exert no attractive
 5. (a) $1/3$
 6. (b) Boyle temperature
 7. (c) diffusion
 8. (b) near the hydrogen
 9. (d) units of pressure and
 10. (c) $8.3 J mol^{-1} K^{-1}$
 11. (a) Boyle's Law
 12. (c) NH_3
 13. (d) I, II and III
 14. (c) $0.41 dm^3$
 15. (c) P
 16. (b) 4
 17. (c) $1/8$
 18. (b) $1/T$
 19. (a) P
 20. (b) NH_3
 21. (c) $mol^{-1} L$ and $L^2 atm$
 22. (d) both assertion and
 23. (c) $3.41 g L^{-1}$
 24. (c)
 25. (d) HI

Additional :

1. (b) increases by three times
 2. (c) $1/273$
 3. (c) square root of molar

4. (d) CO_2
 5. (c) -273.15°C
 6. (b) 303.98 K
 7. (b) 4 : 3

7. Thermodynamics

- (b) DH
- (d) decrease in free
- (b) $q = 0$
- (d) $= 0$
- (a) $w = -\Delta U$
- (d) mass/volume
- (a) -900 J
- (b) negative
- (b) -67.6 kcal
- (a) graphite is more
- (d) -462 kJ
- (d) frictional energy
- (d) $\Delta H < \Delta U$
- (c) $+3 \text{ kJ}$
- (a) -2.48 kJ
- (b) -500 R
- (d) $ba-22$
- (d) -635.66 kJ
- (c) 80 kJ mol^{-1}
- (a) $\Delta H < 0$ and $\Delta S > 0$
- (c) adiabatic expansion
- (d) $(-, -, +)$
- (b) 27°C
- (d) $\text{CaCO}_3(\text{S}) \rightarrow \text{CaO}(\text{S}) +$
- (a) 300K

Additional :

- (c) inversion temperature
- (c) 10^{-4} K
- (b) 1, 3 and 4
- (c) 2 and 5
- (a) $\text{J K}^{-1} \text{ mol}^{-1}$
- (b) 75%
- (c) equal to zero
- (d) $-w - P\Delta V$
- (d) $= 0$
- (d) $\Delta G^\circ = -2.303 RT \log K$
- (c) Volume

- (a) $C_p - C_v = R$
- (c) heat of combustion
- (d) Mole
- (b) 1 Calorie
- (a) $-3227 \text{ kJ mol}^{-1}$
- (b) -57.32 kJ

8. Physical and Chemical Equilibrium

- a) 20
- b)
- a) The forward reaction is exothermic
- c) equilibrium is shifted to the left
- a) increase in pressure
- a)
- a)
- a) 0.06
- b) largely towards reverse direction
- d) $(RT)^2$
- a) $P = 24 \text{ KP}$
- d) $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g})$
- b) $x + 0.5$
- a) 36 : 1
- d) increase by 64
- c) more PCl_5 will be
- a) 33%
- b) 5
- c)
- c)
- b)
- d) remain the same
- a) $\Delta H > 0$ for the
- c)
- b) 5 litre

Additional:

- b) $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$

- b) proceeds in both
- c) does not depend on the
- b) The rates of forward and
- d) 7.33
- c) $(RT)^{-1/2}$
- a) $K = 10^3$
- c) Remains unaltered
- d) Concentration of
- c) High temperature
- a) Adding PCl_5

9. Solutions

- (d) 0.04 M
- (d) (a) and (c)
- (b) 7 mL
- (d) 1×10^{-5}
- (d) 2.5×10^{-4}
- (d) $\Delta G_{\text{mix}} = 0$
- (c) CO_2
- (c) $P_1 - x_2(P_1 - P_2)$
- (b) $\pi V = nRT$
- (d) ethanol + water
- (d) $5x/y$
- (c) 101°C
- (b) mole fraction of solute
- (d) 0.1 M $\text{Ba}(\text{NO}_3)_2$ and 0.1
- (b) $\text{C}_8\text{H}_{16}\text{O}_8$
- (c) 1×10^{-5}
- (c) 2.5 N
- (d) non - ideal and shows
- (d) 0.9965
- (d) 8.89g
- (c) 37°C
- (a) $62.22 \text{ Kg mol}^{-1}$
- (b) 1
- (b) 2.5
- (c)
- (a) 0.1 M KNO_3
- (a) 2.50
- (a) -2°C

29. (d) 0.92
30. (a) both assertion
- Additional:**
1. d) Molality
2. a) $\Delta V = +ve$, $\Delta H = +ve$
3. d) The enthalpy of mixing is zero
4. c) more than vapour pressure of pure water
5. b) Benzene – methanol
6. b) Solvent only
7. c) 0.1 molal
8. b) 4 N
9. a) Osmotic pressure
10. a) 97.5
- 10. Chemical Bonding**
1. (d) SCl_2
2. (d) 0, 0, 0
3. (c) BH_3
4. (d) H_2O
5. (d) $9/2$
6. (d) 89° , 117°
7. (c) assertion is true but reason is false
8. (b) half filled atomic orbitals overlap
9. (d)
10. (b)
11. (d) $O_2^{2-} < C_2^+ < O_2 < C_2^{2-}$
12. (c) s, p_x , p_y , p_z , $d_{x^2-y^2}$
13. (b) $O_2 > O_3 > H_2O_2$
14. (b) O_2^{2-}
15. (a) three
16. (c) Square pyramidal,
17. (c) All five sp^3d hybrid
18. (a) SeF_4 , XeO_2F_2
19. (c) BF_3 and NO_2
20. (a) dissimilar in
21. (a) sp^3 , sp^2 , sp , sp^2 , sp^3
22. (d) ICl_2^-
23. (a) 25, 25, 33.3, 50
24. (c) C_2H_2
25. (c) $l.p - l.p > b.p - l.p > b.p -$
26. (c) "T" Shaped
27. (d) water
28. (c) the resonance hybrid should have higher energy
29. (a) NH_4Cl
30. (d) 4U
- Additional Questions :**
1. d) Charge and size of the ion
2. c) $BaCl_2$
3. c) $AlCl_3$
4. a) 17 % ionic
5. c) BCl_3
6. c) PF_3
7. b) $sp^2 - sp$
8. a) $s - s > s - p > p - p$
9. d) pentagonal bipyramidal
10. b) square pyramidal
11. b) SO_2
12. d) see – saw
13. a) O_2^+ is paramagnetic and bond order is greater than O_2
14. b) $O_2^{2-} < O_2^- < O_2 < O_2^+$
15. a) O_2^-
16. b) CH_4
17. d) ClO_2^- , ClF_2^+
18. c) O_2^+
19. b) 1, 3 and 2
20. b) sp^2 , sp , sp^3
21. a) Metamers
- 11. Fundamentals of Organic Chemistry**
1. (a) $CH_3 - CH = CH - CH_3$
2. (a) sp , sp , sp^3 , sp^2 , sp^3
3. (c) C_nH_{2n-2}
4. (a)
5. (d) 4-Ethyl -3 - methyloctane)
6. (a) 3 – Methyl –3-hexanone
7. (b) Pent -3-en-1-yne
8. (c) 3, 4,4 – Trimethyloctane
9. (a) 2,4,4 – Trimethylpent -2
10. (a) 3 – Ethyl -2- hexene
11. (b) 2 – Hydroxy
12. (c) 3 - Bromo - 2 -
13. (c)
14. (c) 3
15. (d) Glucose
16. (b) dimethylether
17. (c) 9
18. (c) ethanol
19. (b) tautomers
20. (b) $Fe_4[Fe(CN)_6]_3$
21. (c) $C_6H_5 - NH - NH_2$. HCl
22. (d) $C_6H_5NH_2$ and
23. (c) $[Fe(CN)_5NOS]^{4-}$
24. (b) 34%
25. (b) 28%
26. (a) $Mg_2P_2O_7$
27. (c) steam distillation
28. (d) both (a) and (c)
29. (b) distillation under
30. (a)
- Additional:**
1. b) $NH_2 - CO - NH_2$
2. a) 7
3. a) Metamers
4. d) $(4n + 2) \pi$ electrons
5. b) but – 1 – ene – 3 – yne
6. b) 2 – Hydroxy propanoic
7. b) $R - NO_2$
8. c) resolution
9. c) 2 – butene
10. a) Salt formation
11. b) partition
12. c) $40 - 50^\circ C$
13. b) Alumina
14. d) All of these
15. c) Beilstein's test
16. b) NH_3
17. b) distillation

12. Basic Concepts of Organic**Reactions**

1. (d)
2. (a) sp^2
3. (b) $NH_2^- > OH^- > -OCH_3 >$
4. (c) H_3O^+
5. (d) free radical
6. (c) both (a) and (b)
7. (d) $(CH_3)_3C^-$
8. (d) $C_6H_5NH_3^+$
9. (c) both (a) and (b)
10. (a)
11. (a)
12. (d) Carbanion and
13. (c) CN^- , RCH_2^- , ROH
14. (d) BF_3
15. (c) Planar

Additional :

1. a) Planar
2. b) A free radical
3. b) sp^2 hybridized
4. b) Lewis acids
5. a) Electron loving species
6. a) Temporary effect
7. d) $-CH_3$
8. c) free radicals
9. d) Tertiary carbonium ion
10. b) Addition reactions
11. c) electrophilic substitution
12. d) Triphenylmethyl cation

13. Hydrocarbons

1. (b)
2. (b) Wurtz reaction
3. (d)
4. (a) $sp^3 - s$ and $sp^3 - sp^3$
5. (c)
6. (d) none of these
7. (c) C_2H_6 and CO_2
8. (b) C_nH_{2n}
9. (a) C_3H_6

10. (c)
11. (d) 2, 3 - Dimethyl 2 - ene
12. (a) 2 - methylbut - 2 - ene
13. (a) 2-chloro -1- iodo - 2 -
14. (a) trans-2-chloro-3-iodo -
15. (c) configurational isomers
16. (a)
17. (c) alc. KOH
18. (d) slower
19. (d)
20. (a)
21. (d)
22. (d)
23. (a) Nitro benzene
24. (b) $-NO_2$
25. (d) isopropyl chloride
26. (b) action of sodium metal
27. (a) C_8H_{18}
28. (a) Formaldehyde
29. (c) pent - 1 - ene
30. (d) 2, 2, 3, 3 - tetra chloro

Additional:

1. d) mixture of ethane and
2. d) PtO_2
3. c) Staggered
4. a) n - propyl chloride
5. c) 2 - butene
6. d) Alkaline potassium
7. a) Teflon
8. b) Free - radical mechanism
9. a) 3
10. b) Ethyne
11. b) Destructive distillation
12. b) Benzene hexa chloride
13. b) fractional distillation
14. d) $-NH_2$

14. Haloalkanes and**Haloarenes**

1. (b) 4-Bromo pent-2-ene
2. (a) n-Butyl chloride

3. (a) $D < C < B < A$
4. (b) Allyl
5. (a) 3 - Chloro pentane
6. (d) Fluoromethane
7. (b)
8. (c) (iii)
9. (b) $R-CH-R-OH$
10. (a) Chlorobenzene
11. (c) Freon-114
12. (c) aqueous KOH
13. (d) $A \rightarrow 3B \rightarrow 1C \rightarrow 4D \rightarrow 2$
14. (b)
15. (c) DMF
16. (b) Swarts reaction
17. (d) benzyl chloride
18. (b) sp^2 hybridised
19. (a) 1-chloro-4-nitrobenzene
20. (d)
21. (a) acetaldehyde
22. (c) benzene
23. (c) chloropicrin
24. (b) 2-methyl-2-propanol
25. (c) bromoethane

Additional:

1. c) Mg in dry ether
2. c) C_2H_5Cl
3. b) Chloro benzene
4. c) Raschig process
5. a) Fitting reaction
6. c) chloro benzene and chlora
7. b) antiseptic
8. c) CF_2Cl_2
9. a) p, p' - dichloro diphenyl
10. d) Ethane
11. a) Friedel - craft's reaction
12. a) inversion of
13. a) $3^\circ > 2^\circ > 1^\circ$
14. b) Cu_2Cl_2 / HCl

15. Environmental**Chemistry**

- 1.(d) stratosphere
- 2.(a) Forest fire
3. (b) air pollution
4. (c) Carbon monoxide
5. (c) $\text{CFC} > \text{N}_2\text{O} > \text{CH}_4 > \text{CO}_2$
6. (b) Ozone, PAN and NO_2
7. (C) 5.6
8. (c) bio magnification
9. (a) The clean water would have a BOD value of more than 5

10. (c) Combines with
11. (d) catalytic convertors
12. (c) rich in dissolved oxygen
13. (c)
14. (b)
15. (d) iv
16. (a) i
17. (d) iv

Additional :

1. (a) NO and freons
2. (a) acrolein
3. (b) low temperature

4. (b) Methyl isocynate
5. (b) Mesosphere
6. (d) winter during morning time
7. (b) CO_2
8. (d) Mercury
9. (b) Phosphate and detergents
10. (b) reduce the use and
11. (d) Non – biodegradable
12. (c) Arsenic
13. (b) Mn

Thank God

Note:

- ✓ I hope this material will be useful for test practice from the evaluation and additional MCQ with the help of teachers.
- ✓ It will be better to give importance to the evaluation part questions then can study additional questions.
- ✓ Above average students should study text book well for creative MCQ questions
- ✓ If any mistakes or your suggestions, please send your valuable thoughts to that email to help the students
- ✓ It has been updated on October 2022

DEDICATED TO : ALL THE TEACHERS AND STUDENTS

G. SURESH M.Sc, M.A, B.Ed,

**P.G. ASSISTANT IN CHEMISTRY
&
CAREER COUNSELOR**

THE CRESCENT MATRIC HR.SEC.SCHOOL

UTHAMAPALAYAM

THENI [DT]

E mail: vivekasuresh@gmail.com

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**“THANK GOD AND THANK YOU ALL”
“ALL THE BEST”**