

**ACTC ADVANCED CHEMISTRY TUITION CENTRE, NAGERCOIL, KK DIST 9940847892****PLAN!****PREPARE!!****PRESENTATION!!!**

# +2 CHEMISTRY QUESTION BANK 2023

**(WITH TEXT BOOK PAGE NUMBER)****Question Bank from**

- **GOVT PUBLIC EXAM PREVIOUS YEAR QUESTION PAPER - MARCH 2020(M20), JULY 2020 (J20), SEP2020(S20), AUG2021(A21), MAY 2022(M22), JULY 2022(J22)**
- **GOVT MODEL QUESTION PAPER 2019-20 (GM).**
- **TAMIL NADU STATE TEACHER PARENTS ASSOCIATION 2019-20 PTA 1-6 Question paper (PTA).**
- **Text book inside**

**+2 CHEMISTRY GOVT PUBLIC PREVIOUS QUESTION PAPER ANALYSIS**\*C- Compulsory **CP- COMPULSORY PROBLEM** **PRO- PROBLEM**

LN	MARCH 2020					JULY 2020 (exam absent M20)					SEPTEMBER 2020 (Arrear)					AUGUST 2021 (Arrear)				
	1M	2M	3M	5M		1M	2M	3M	5M		1M	2M	3M	5M		2M	3M	5M		
1	1			1(5)	7	1	1	1	6		2	1		1(8)	6	1			1(5)	6
2	1		1C	1(2)	6	1	1C		1(3)	6	1		1	1(2)	6	1	1	1		6
3	1	1		1(3) 1(3)	9	1		1	1(3) 1(2)	9	1	1		1(3) 1(2)			1	1	1(3) 1(2)	9
4	1	1	1		6	-			1(5)	5	1	1	1	6	-	1			1(3) 1(2)	7
5	-	1	1	1(2) 1(3)	10	2	1	1	1(2)	9	1			1(5)	9	1		1	1(2) 1(3)	9
6	1	1		1(2)	5	1		1	1(2)	6	1	1		1(3)	6	1	1	1C		9
7	1			1(5)	6	-		1C	1(3) PR O	6				1(2)	5	1			1(2) 1(3)	6
8	1	1	1	1(2)	8	1	2	1	1(2)	10	1	1	1	1(3)	9	1	1		1(5) PRO	8
9	2		1	1(3) PRO	8	1			1(5)	6			1CP	1(5)	8	1	2	1		8
10	-	1	1	1(3)	8	-	1		1(5)	7	1		1	1(2) 1(2)	8	2	1		1(5)	9
11	2	1C		1(2) 1(3)	9	2		1	1(5)	10	1	1		1(3) 1(2)	8	1		1	1(2) 1(3)	9
12	1	1	1	1(2)	8	1		1	1(3) 1(2)	11	1		1	1(5) Me	9	2		1	1(5)	10
13	1	1		1(2) 1(3)	8		1	1		6	1	1		1(3) 1(2)	8	1	1C		1(5) PRO	8
14	1		1	1(3)	7	2	1		1(3)	7	1	1	1		6	1	1	1		6
15	1		1	1(1)	6	1			1(3) 1(2)	6	1		1	1(3)	7	REDUCED SYLLABUS NO QUESTION CORONA				
	15	9	9		110	15	9	9	50M	110	15	9	9	50M	110	15	9	9	50M	110

Inorganic	UNIT 1,2,3,4,5	Physical	Unit 6, 7,8,9,10	Organic	Unit 11, 12, 13, 14,15
Part I	Choose 5x1=5	Part I	Choose 5x1=5	Part I	Choose 5x1=5
Part II	3x2=6		3x2=6		3x2=6
Part III	3x3=9		3x3=9		3x3=9
Part IV	3x5=15		3x5=15		4x5=20
	35		35		40

**M20 – MARCH 2020, J20-JULY 2020, S20- SEPTEMBER 2020, A21-AUGUST 2021, MAY 2022 (M22), JULY 2022 (J22)**

**2M- 2 marks; 2MC- 2 marks Compulsory; 3M- 3 marks; 3MC- 3 marks Compulsory; 5M- 5 marks; GM2M- Govt model question paper 2 marks; & PTA Questions**

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**PLAN! PREPARE!!! PRESENTATION!!!**

LN	MAY 2022					JULY 2022					MARCH 2023					JUNE 2023				
	1M	2M	3M	5M		1M	2M	3M	5M		1M	2M	3M	5M		1M	2M	3M	5M	
1	1bb	1		1(3) 1(2)	8	1		1	1(5)	9										
2	1bb			1(2) 1(3)	6	1		1	1(5)	9										
3	1bb		1	1(3) 1(2)	9	1	1		1(5)	8										
4	1bb	1	1		6	1	1		1(5)	8										
5	1bb		1C	1(5)	9	1	1	1		6										
6	1bb	2		1(5)	10	1	1		1(2) 1(3)	8										
7	1bb	1	1	1(5)	11	1		1	1(5)	9										
8	1bb	1		1(2) 1(3)	8	1	1	1	-	6										
9	1bb		1	1(5)	9	1	1C PRO		1(5)	8										
10	2bb	1	1	1(5)	12	1		1	1(5)	9										
11	1bb	1	1		6	1	1	1		6										
12	1bb		1	1(5)	9	2	1		1(3) 1(2)	9										
13	1bb	1C		1(2) 1(3)	8	1		1C	1(3) 1(2)	9										
14	1bb		1		4	1	1	1		3										
15	REDUCED SYLLABUS NO QUESTION CORONA					REDUCED SYLLABUS NO QUESTION CORONA														
	15	9	9	50M	110	15	9	9	50M	110										

**MAY 2022 PUBLIC QUESTION PAPER B-TYPE 13/5/2022 Friday**

Q No	Answer	Book exercise	Book inside	Page No	Unit	BookMCQ No
1	c) Glycerine	Book exercise		270	14	25
2	a) Methanol	Book exercise		141	11	6
3	a) Basic	Book exercise		53	2	1
4	d) FeO	Book exercise		200	6	22
5	c)Hydrolysis	Book exercise		101	10	14
6	c)Sn/HCl	Book exercise		229	13	1
7	d) charge carried	Book exercise		63	9	6
8	a) Al <sub>2</sub> O <sub>3</sub> ·H <sub>2</sub> O	Book exercise		20	1	1
9	b) 30 min	Book exercise		229	7	22
10	a)HPO <sub>4</sub> <sup>2-</sup>	Book exercise		30	8	19
11	c)Fe(CO) <sub>5</sub>	Book exercise		171	5	16
12	c)+3	Book exercise		127	4	15
13	c)1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>3</sup>	Book exercise			3	3
14	a)liquid in gas	Book exercise		100	10	4
15	c)nucleophilic addition	Book exercise		189	12	2

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1. Difference between ore and mineral. (2) BB (S20) M22 2M
2. Write the ores of copper, silver, Aluminium, Iron, zinc (3)
3. What are all the steps involves in metallurgical process?(2)
4. Write note on gravity separation.(3) M22 5Mi
5. What is Gangue?(3)
6. Explain froth floatation.(6) (A21) 5M
7. Describe the underlying principle of forth floatation process. (4) PTA 2M
8. Give the depressing agents used in the froth floatation process and why we use depressing agents in that process? (OR)Role of depressing agent in froth floatation process?(4) PTA2M
9. which type of ores can be concentrated by froth flotation method give two example (J20)
10. What are leaching process?(4)
11. Explain Cyanide leaching.(4)
12. Explain how gold ore is leached by cyanide process. (4) GM5Mi
13. What is Cementation?(4)
14. What is Ammonia leaching.(4)
15. What is Acid leaching?(4) J22 3M
16. Explain magnetic separation.(6)
17. How will you manage sulphur dioxide produced during roasting process?(7)
18. Define the following terms (i) Roasting (6) (ii) Calcination (7) PTA3M
19. What is smelting? (8)
20. What is the role of limestone in the extraction of iron from its oxide  $Fe_2O_3$ .(8) (J20) (BB)  
ANS: Used As Flux. (Govt. key)
21. Explain the following terms with suitable example. a)Gangue b) Slag.(BB) PTA3M(S20)
22. Explain extraction of copper from copper pyrites.(8) PTA5M i
23. What is meant by blistered copper.(9)
24. Give one example for (i) Acidic flux (ii) Basic flux (8)
25. CO is reducing agent. Justify with an example.(8)& (LN 2 43) PTA 2M
26. Reduction by hydrogen, carbon, metal(9)
27. How  $Cr_2O_3$  is reduced to Cr by Al powder? (10) PTA 5M ii
28. Explain Auto reduction.(10)
29. Define, observation, application & Limitations of Ellingham diagram.(13)
30. Extraction of Aluminium-Hall -Herold process. (14)
31. Define refining process(15)
32. Write note on distillation. (15)
33. Write note on Liquation. (15)
34. Explain electro refining process. (15)
35. Explain electrolytic refining of silver. (16) PTA5M ii (Explain the principle of electrolytic refining with an example. J22 5M)
36. Explain zone refining process.(16) M20 5M, PTA 3M
37. Explain Mond process. (Describe the method for refining of Nickel.)(16) PTA5M ii, M22 5Mii
38. Explain Van-Arkel method. (17)
39. Application of Al, Zn, Fe, Cu & Au. & Revise Book Back (Evaluation) Question Answer

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**ACTC ADVANCED CHEMISTRY TUITION CENTRE, NAGERCOIL, KK DIST 9940847892****PLAN!****PREPARE!!****PRESENTATION!!!****LESSON 2 p-block elements-I**

1. Explain general characteristics(properties) of p-block elements. (27)
2. Write note on metallic nature of p-block elements. (28)
3. Give one example for Icosogens, Tetragen, pnictogen, chalcogen. (28)
4. What are the anomalous properties of the first elements of the p-block elements. (29) **S20 A21**
5. There is only a marginal difference in decrease in ionization enthalpy from Aluminium to Thallium – Explain why? (29)**M20 Compulsory 3 mark**
6. What is inert pair effect.(30) **M22 5Mi**
7. Occurrence of Group 13 (Boron group) elements. (31)
8. Chemical properties of Boron. ( $X_2$ ,  $N_2$ ,  $O_2$ ,  $H_2SO_4$ ,  $HNO_3$ ,  $NaOH$ ) (33)
9. A hydride of 2<sup>nd</sup> period alkali metal (A) on reaction with compound of Boron (B) in the presence of ether to give a reducing agent (C). Identify A, B and C. **J20 2M Compulsory**
10. Uses of Boron. (22)
11. Preparation, properties & uses of Borax. (33,34)
12. Preparation of Boric acid. (34)
13. How will you identify borate radical? Write the reactions involved (35) **PTA3M, GM2M**
14. How will you convert Boric acid to boron nitride?(35) **PTA2M**
15. Structure of Boric acid.(35)
16. What are the Uses of Boric acid. (35) **J22 3M, M22 5Mi**
17. How does borane react with  $H_2O$  &  $NaOH$ ? ( 36 )
18. How will you prepare borazine? (37)
19. Explain structure of diborane. (37)
20. Mention the uses of diborane (38)
21. How will you prepare  $BF_3$ ? ( 38 )
22. Write note on McAfee process ( 39 )
23. Mention the uses of Aluminium chloride ( 39 )
24. How to prepare potash alum & uses. (40) **J20, PTA 5M ii**
25. What is burnt alum.(40)
26. Occurrence of Group 14 (Carbon group) elements.(40)
27. What is catenation? Write the conditions for catenation property. (41) **J22 5M** Substantiate this statement. (41) **S20 2M** Write any two conditions for catenation. **M20 2M**
28. What are allotropes of carbon, difference between graphite & diamond. (41)
29. Write note on fullerenes.(42)
30. Write note on nanotubes.(42)
31. Write note on graphene.(43)
32. What is producer gas, water gas? (43,45)
33. Write note on Fischer Tropsch synthesis.
34. Write note on structure & uses of  $CO$  &  $CO_2$  (44, 45)
35. How does  $SiCl_4$  react with alcohol &  $NH_3$ ? (46)
36. Explain the preparation, uses of silicone. (47 & 48)
37. What are the types, properties of silicones. (47)
38. Explain types of silicates. (48)
39. Explain Zeolites. (50)

**ALL CHEMICAL EQUATION & Revise Book Back (Evaluation) Question Answer****E.MUTHUSAMY MSc(Che), MSc(Psy), MEd., MPhil., MA(Eng), MA(T), MA(PA), MA(Soc), BLISc., DMLT.****B. SARANYA MUTHUSAMY BE., BEd., You Tube: ACTC Chemistry Whatsapp: 9940847892**

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**LESSON 3 p-block elements -II**

1. Occurrence of 15<sup>th</sup> group elements. (57)
2. Physical properties of 15<sup>th</sup> group elements. (57)
3. How will you prepare nitrogen from sodium azide and atmospheric air? (57)
4. N<sub>2</sub> is a chemically inert. Why? (58)
5. What are the uses of Nitrogen? (58)
6. Write note on Haber's process (58) **S20**
7. How is ammonia prepared in the laboratory? (58)
8. What is the reaction of Ammonia with Iron and copper salts? (60) **PTA5M ii**
9. Structure of Ammonia. (60)
10. Preparation of Nitric acid Oswald process. (61)
11. How is fuming nitric acid decompose on exposure? (61)
12. How does nitric acid act as nitrating agent? (62)
13. How does copper react with nitric acid? (63)
14. Give the uses of nitric acid (63)
15. Draw the structures of oxides of nitrogen (65)
16. Draw the structures of oxoacids of nitrogen (65) **M22 ONE MARK (hyponitrous acid)**
17. Explain Allotropy of phosphorus. (67)
18. How does phosphorous act as a reducing agent? (68) or  
Complete the reaction: P<sub>4</sub> + NaOH + H<sub>2</sub>O → (68) **PTA 2MC**
19. How does phosphorous react with HNO<sub>3</sub>? (68)
20. Uses of phosphorus. (68)
21. How is phosphine prepared? (68)
22. What is Holmes signal and uses? (70) **S20 5Mii**
23. Draw the structure of PH<sub>3</sub> (70) PCl<sub>3</sub> (71) P<sub>2</sub>O<sub>3</sub> (P<sub>4</sub>O<sub>6</sub>) (71) P<sub>4</sub>O<sub>10</sub> (72)
24. Write the structure and basicity of following oxy acids. (72) **3M**  
i) hypo phosphoric acid ii) ortho phosphoric acid iii) pyro phosphoric acid
25. Occurrence of 16<sup>th</sup> group elements. (73)
26. Physical properties of 16<sup>th</sup> group elements. (73)
27. How is ozone estimated? (75) (Ozone (O<sub>3</sub>) act as a powerful oxidizing agent why?)
28. Laboratory preparation of O<sub>2</sub>. (74)
29. Laboratory Preparation of ozone (74)
30. Structure of ozone. (74)
31. How is ozone estimated? (75) (Ozone (O<sub>3</sub>) act as a powerful oxidizing agent why?) (75)  
**PTA2M**
32. Uses of oxygen (75) **M22 5Mi**
33. Allotropic form of sulphur. (73)
34. Lab preparation of SO<sub>2</sub>. (76)
35. SO<sub>2</sub> is an acidic or basic oxide. Why? (76)
36. Explain the bleaching action of Sulphur dioxide. (77) **A21**
37. Write note on contact process. (77)
38. Structure of SO<sub>2</sub>, Uses of SO<sub>2</sub>. (77)
39. Manufacture of Sulphuric acid by contact process. (77)
40. How does H<sub>2</sub>SO<sub>4</sub> act as an oxidizing agent? (78, 79)

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41. Sulphuric Acid is a dibasic acid. Prove it. (78) **J20 5Mii**
42. How is sulphate radical/sulphuric acid detected? (80)
43. Draw the structure of oxoacids of sulphur. (80) Sulphurous acid, sulphuric acid, Marshall's acid (**M20**), Caro's acid, Dithionic acid.  
(Write the molecular formula and draw the structure of sulphurous acid and Marshall's acid.) (**M20 5Mi**)
44. Occurrence of 17<sup>th</sup> group elements. (81)
45. Physical properties of 17<sup>th</sup> group elements. (81)
46. Why fluorine is more reactive than other halogens? (BB98) **PTA1 2M & PTA3 5M**
47. Manufacture of chlorine. (electrolytic process) (83)
48. How will you prepare chlorine in the laboratory? (82BB83) **2M**
49. Explain Deacon's process for manufacture of chlorine. (83) **J20 5Mii**
50. Physical properties of chlorine. (83)
51. Give the balanced equation for the reaction between chlorine with Cold NaOH and hot NaOH. (84) **S20**
52. Write about bleaching action of chlorine (85)
53. How will you prepare bleaching powder? (85) **M20 2M, M22 5Mii**
54. Uses of chlorine. (86)
55. How is aqua regia obtained? (86)
56. HF is not stored in glass bottles. Why? (88) **M20 5Mii**
57. What are interhalogen compounds. Give two examples. (89) **GM5Mi A21, M22 3 mark**
58. What are the properties of interhalogen compounds. (89) **PTA5M, J22 5M**
59. Give reasons: ICl is more reactive than I<sub>2</sub>. **PTA5Mi**
60. Structure of interhalogen compounds AX<sub>2</sub>, AX<sub>3</sub>, AX<sub>5</sub>, AX<sub>7</sub>. (90)
61. What type of hybridisation is found in the following? (a) BrF (b) BrF<sub>5</sub> (c) BrF<sub>3</sub> (d) IF<sub>7</sub> (91) (**J20**) **3M, PTA 5M ii**
62. Occurrence of 18<sup>th</sup> group elements. (91)
63. Physical properties of 18<sup>th</sup> group elements. (91)
64. How does XeF<sub>6</sub> react with NaOH? (92)
65. How does XeF<sub>6</sub> react with SiO<sub>2</sub>? (92)
66. How does sodium per xenate act as strong oxidizing property? (92)
67. Uses of Helium (93) **GM 3M, S20, A21, PTA 3M**
68. Uses of Neon. (93)
69. Uses of Argon. (93) **PTA 3M J22 2M**
70. Uses of Krypton. (93)
71. Uses of Xenon. (93)
72. Uses of Radon. (93)
73. List any five compounds of xenon and mention the type of hybridization and structure of the compounds. (93) **PTA5M**

**ALL CHEMICAL EQUATION & Revise Book Back (Evaluation) Question Answer**

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- d-block elements are called transition elements. Justify this statement. (101)
- What are transition elements? Write two characteristics of the transition elements?(102)  
**PTA5Mi**
- Write general Electronic configuration of d-block, f-block elements.
- Classify the following elements into d-block and f-block elements: (M20)  
i) Tungsten      ii) Ruthenium      iii) promethium      iv) Einsteinium
- How many series are in d-block elements? What are they?(102)
- Why there is a slight variation in the atomic radii from Cr to Cu?(104)
- Applying Aufbau principle, write down the electronic configuration of  $\text{Cr}^{3+}$  and Cu.(102)
- What are the metallic behavior of d-block elements.(102)
- Transition metals show high melting points. Why? (103) **PTA 2MARK**
- d-block elements have variable oxidation state. Why?(106) **PTA 2MARK**
- Write a note about oxidation state of 3d series.(106)
- $\text{Mn}^{2+}$  is more stable than  $\text{Mn}^{4+}$ . Why? (106)
- Which is more stable  $\text{Fe}^{3+}$  or  $\text{Fe}^{2+}$ ? Why? **BB M22 2M**
- Ru and Os have highest oxidation state in which compounds? Explain with example.(106)
- Which metal in the 3d series exhibits +1 oxidation state most frequently and why?(106-107)  
**S20 3MARK**
- Define – Standard electrode potential.(107)
- Explain why  $\text{Cr}^{3+}$  is strongly reducing while  $\text{Mn}^{3+}$  is strongly oxidizing? (108)**PTA5M i**
- Write note on diamagnetic. Give example.(109)
- Write note on paramagnetic. Give example (109)
- $\text{Sc}^{3+}$ ,  $\text{V}^{5+}$  are diamagnetic. Give reason (110)
- Calculate the magnetic moment of  $\text{Ti}^{3+}$  &  $\text{V}^{4+}$ . (110)
- Calculate the number of unpaired electrons in  $\text{Ti}^{3+}$ ,  $\text{Mn}^{2+}$  and calculate the spin only magnetic moment.(110) **A21, PTA 3MARK**
- Most of the transition metals act as catalyst. Justify this statement.(110) **PTA 3M**
- Explain the catalytic hydrogenation of alkene to alkane with equation. (110)
- What is Zeigler – Natta catalyst? In which reaction it is used? Give equation.(110) **J22 2M**
- Hume-Rothery rule. (110)
- d-block elements readily form Alloy. Give reason.(111)
- d-block elements formation of interstitial compounds (What are interstitial Compound)(111)  
(What are interstitial compounds? How they differ from the properties of its pure metals?)(111) **S20, A21 2MARK, PTA 3MARK**
- What are the properties of interstitial compounds? (111) **M22 3MARK**
- d-block elements readily form complexes. Give reason. (111)
- Describe preparation of potassium dichromate. (112)
- How does potassium dichromate decompose on heating? (113)
- Draw the structure of chromate, dichromate ions (113)
- How does potassium dichromate act as oxidizing agent? (113)
- Explain chromyl chloride test (114) **M20 3M**
- Mention the uses of potassium dichromate ( 114 )
- How will you prepare potassium permanganate? ( 115 )

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38. What is Bayer's reagent? (117)  
 39. Give the uses of potassium permanganate (118)  
 40. Justify the position of Lanthanoids and Actinoids in the periodic table. (BB) (118) **PTA 5M i**  
 41. Electronic configuration of Lanthanoids. (120) Write down the electronic configuration of Gd (Z=64), Th (Z=90).  
 42. Oxidation state of Lanthanoids. (120)  
 43. Define, cause and consequences of Lanthanoid contraction. (121) **J20 5 MARK, PTA 3M**  
 44. Which is more basic among  $\text{Lu}(\text{OH})_3$  and  $\text{La}(\text{OH})_3$ ? Why? (BB) (121) **2M**  
 45. Electronic configuration of actinoids. (122)  
 46. Oxidation state of actinoids. (123)  
 47. Differences between Lanthanoids and Actinoids. (123) **PTA 5 MARK, J22 5M**  
 & **Revise Book Back (Evaluation) Question Answer**

**LESSON 5 Coordination Chemistry**

1. Difference between double salt and coordination compounds. (131) (A21, PTA 3M)  
 2. Explain Werner theory & limitation. (132) (S20 5M, M22 5M)  
 3. Write note on coordination entity. (133)  
 4. Define central metal ion (133)  
 5. Write note on ligand (134)  
 6. Write the IUPAC ligand name for the following: a)  $\text{Cr}_2\text{O}_4^{2-}$  b)  $\text{H}_2\text{O}$  c)  $\text{Cl}^-$  (137) **J22 3M**  
 7. Define coordination number. (134) **M22 2M**  
 8. Define oxidation number. (134)  
 9. Write note on coordination sphere. (134)  
 10. Write note on coordination polyhedron. (134)  
 11. In the complex  $[\text{Pt}(\text{NO}_2)(\text{H}_2\text{O})(\text{NH}_3)_2\text{Br}]$  identify the following i) central metal ion, ii) Ligand iii) coordination entity iv) oxidation number of the central metal ion. V) coordination number.  
 12. Write the IUPAC name of the following: a)  $[\text{Ag}(\text{NH}_3)_2]^+$  B)  $[\text{Co}(\text{NH}_3)_5\text{Cl}]^{2+}$  (M20 2M)  
 Write the following for the complex  $[\text{Ag}(\text{NH}_3)_2]^+$ . **M22 3M Compulsory**  
 a) Ligand b) Central metal ion c) IUPAC Name  
 13. Write the IUPAC Name for the compound  $\text{Na}_2[\text{Ni}(\text{EDTA})]$  (140) **PTA 5M i**  
 14. For the example  $[\text{Fe}(\text{en})_2\text{Cl}_2]\text{Cl}_2$ , Identify **PTA 5M i**  
 1) Oxidation number of Fe, 2) Hybridization and shape  
 3) Magnetic behavior, 4) Number of geometric isomers  
 5) Whether there may be optical isomer also? 6) IUPAC name  
 15. Explain types of complexes. (135)  
 16. Classification based on the net charge on the complex. (135)  
 17. Classification based on kind of ligands. (135)  
 18. Nomenclature of coordination compounds IUPAC Names (138-141)  
 19. Explain structural isomer. (coordination, Linkage, ionization, solvate isomers) (142)  
 20. Write any two hydrate isomers of the complex with the molecular formula  $\text{CrCl}_3 \cdot 6\text{H}_2\text{O}$ . (142) **M20 2M**  
 21. Explain Geometrical isomers (cis, trans isomer) (144, 145)  
 22. Why tetrahedral complexes do not exhibit geometrical isomerism? (BB) (143)  
 23. Define mer, fac isomer. (145)  
 24. Explain optical isomerism of coordination compounds with an example. (146)

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25. Explain Valence Bond theory (VB Theory) & limitations. (S20, J22 2M, PTA)(149 &152)
26. Apply VB theory in  $[\text{Ni}(\text{CO})_4]$ ,  $[\text{Ni}(\text{CN})_4]^{2-}$  (S20),  $[\text{Co}(\text{CN})_6]^{3-}$ ,  $[\text{Co}(\text{F})_6]^{3-}$  (M20) (149-151)
27. Explain crystal field theory. (153)
28. In an Octahedral crystal field, draw the figure to show splitting of d orbitals (154) J20 3M
29. In a tetrahedral crystal field, draw the figure to show splitting of d orbitals (155)
30. Write note on spectrochemical series (156)
31. Calculate the CFSE value of  $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$  in high spin and low spin complexes. (157)
32. Calculate the CFSE value of  $[\text{Fe}(\text{CN})_6]^{3-}$  in high spin and low spin complexes. (158)
33.  $[\text{Sc}(\text{H}_2\text{O})_6]^{3+}$  is colourless – Explain. M20 3M (Explain about d-d transition) (159,160)
34. How is metal carbonyls classified based on the number of metal atom? (161)
35. How is metal carbonyls classified based on the structure? (161, 162)
36. How can the stability of coordination complexes be interpreted? (163)
37. Define Labile, inert complexes. (163)
38. What is meant by stability constants? (163) significance of stability constants? (164)
39. Explain the importance and application of coordination compounds (166)
40. Mention the metal complexes and its metal ions are used in biological system (167) S20 3M
41. Give one test to differentiate  $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{SO}_4$  and  $[\text{Co}(\text{NH}_3)_4\text{SO}_4]\text{Cl}$ . (BB) PTA 2M

*ALL IUPAC Name, Apply VB Theory & Revise Book Back (Evaluation) Question Answer*

**PHYSICAL CHEMISTRY LESSON 6 SOLID STATE**

1. Write General characteristics of solids. (177)
2. Difference between crystalline solids and Amorphous solids.(178)PTA 3M, M22 5M
3. Define Isotropy and anisotropy.(178) (S20)
4. Explain ionic solids. (Introduction, NaCl diagram, Characteristics)(179)
5. What are the characteristics of Ionic solids. (179) PTA 5M i
6. What are covalent solids? (179) M22 2M
7. Explain molecular solids. (179)
8. What are Metallic solids? (180)
9. Classify the following into covalent, molecular, ionic and metallic solids.( A21) 3MARK C  
(i) Diamond (ii) Brass (iii) NaCl (iv) Naphthalene (v) Glucose (vi)  $\text{SiO}_2$
10. Define unit cell. (180) J22 2M
11. What is meant by term "Coordination Number"? What is the Coordination Number of atoms in a bcc structure? (180) A21 3MARK, M22 2M
12. Define primitive & non-primitive unit cells? (181)
13. Define the terms crystal lattice and unit cell. (180) PTA3M
14. Write note on SC. (183)
15. Write note on BCC. (183)
16. Sketch Face Centered cubic unit cell(FCC) and Calculate the number of atoms present in it. (184) PTA3M
17. What is Bragg's equation? (184) PTA 2M
18. How will you calculate the density of the unit cell? (185)
19. Barium has a body centered cubic unit cell with a length of 508pm along an edge. What is the density of barium in  $\text{g cm}^{-3}$ ? (185)
20. An element has a face centered cubic unit cell with a length of 352.4 pm along an edge. The density of the element is  $8.9 \text{ g cm}^{-3}$ . How many atoms are present in 100 g of an element. (186)

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21. What is packing efficiency? (187) **PTA 5M ii, J22 5Mi**
22. How will you calculate the packing efficiency for simple cubic?(187)
23. Calculate the percentage efficiency of packing in body centered cubic system (188)**PTA3M**
24. What is void and its type? (190)
25. Note on radius ratio (192)
26. If the radius of the compound is between 0.155 – 0.225, find out the co-ordination number and structure of the compound. (192) **J20 5Mi**
27. If the no. of close packed sphere is 6, calculate the number of Octahedral voids and Tetrahedral voids generated. **(M20) 2MARK**  
**ANS:** Octahedral voids (n)=6 Tetrahedral voids(2n)=2x6=12
28. Calculate the percentage efficiency of packing in face centered cubic system (192)
29. Outline the classification of point defects. (193) **PTA5M ii**
30. Explain Schottky defect **(GM5M, S20)**& Frenkel defect. **(M20, PTA3M, J22 5Mii)** (Stoichiometric defects) (193)
31. Explain 'f' centers with a neat diagram.(194) **PTA5M ii**
32. Explain Metal excess defect, metal deficiency defect. (Non-Stoichiometric defects) (194)
33. Write note on Impurity defect. (195)
34. Atoms X and Y form BCC crystalline structure. Atom X is present at the corners of the cube and Y is at the centre of the cube. What is the formula of the compound? **(BB)PTA 5M i**
35. An atom crystallizes in FCC crystal lattice and has a density of  $10\text{gcm}^{-3}$  with unit cell edge length of 100pm. Calculate the number of atoms present in 1g of crystal. **(BBQ22201) 5M**
36. A face centred cubic solid of an element (atomic mass  $60\text{gmol}^{-1}$ ) has a cube edge of  $4\text{\AA}$ . Calculate its density.(186) **GM3M**
37. What is piezoelectricity? (195)

**ALL INSIDE PROBLEM & Revise Book Back (Evaluation) Question Answer**

**LESSON 7 CHEMICAL KINETICS**

1. Define chemical kinetics. (205)
2. Define Rate of a chemical reaction. (205)
3. Write note on stoichiometry and rate of a reaction.(206)
4. Define average rate and instantaneous rate. **(BB)207**
5. Define rate law and rate constant. **(BB)208**
6. Give the difference between rate of a reaction and rate constant.(209) **PTA,A21**
7. What is an elementary reaction? (210)
8. Explain the rate determining step with an example. (210) **PTA 3M**
9. Define order and molecularity of a reaction. (210) **J22 3M**
10. Give the differences between order and molecularity of a reaction. (210)
11. The rate of the reaction  $x + 2y \rightarrow \text{product}$  is  $4 \times 10^{-3}\text{mol L}^{-1}\text{s}^{-1}$ , if  $[x] = [y] = 0.2\text{M}$  and rate constant at 400k is  $2 \times 10^{-2}\text{s}^{-1}$ , what is the overall order of the reaction?(211) **(J20) 2 MARK**
12. Derive integrated rate law for a first order reaction  $A \rightarrow \text{product}$ .(212) **(PTA,M20) 5MARK**
13. Derive and describe the graphical representation of first order reaction.(212)
14. Give examples for the first order reaction. (213) **M22 2M**
15. Explain pseudo first order reaction with an example. (214)**GM3M**
16. Derive integrated rate law for a zero order reaction  $A \rightarrow \text{product}$  & Example (214) **A21 5M, J22 5M**

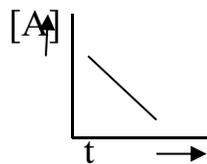
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17. Define half life period of reaction. Show that for a first order reaction half life period is independent of initial concentration. (215) PTA 5M i
18. Calculate the half period for a zero order reaction. (215) PTA 2M
19. Write Arrhenius equation and explain the terms involved. (220) M22 3M
20. Write the rate law for the following reactions. (BB)
- (a) A reaction that is 3/2 order in x and zero order in y.
- (b) A reaction that is second order in NO and first order in Br<sub>2</sub>.
21. The rate constant for a first order reaction is  $1.54 \times 10^{-3} \text{ s}^{-1}$ . Calculate its half life time. (BB 231) (PTA, J20)
22. Identify the order for the following reactions (BB)
- (i) Rusting of Iron
- (ii) Radioactive disintegration of  ${}_{92}\text{U}^{238}$
- (iii)  $2\text{A} + 3\text{B} \rightarrow \text{products}$ ;  $\text{rate} = k [\text{A}]^{\frac{1}{2}} [\text{B}]^2$
23. Write any three differences between order and molecularity. (210) 5M i
24. Explain about collision theory (217, 218)
25. Derive Arrhenius equation to calculate activation energy from the rate constant  $k_1$  and  $k_2$  at temperature  $T_1$  and  $T_2$  respectively. (220, 221) PTA 3MC
26. What are the factors affecting the reaction rate? (222)
27. Explain the effect of catalyst on reaction rate with an example. (222) S20 5Mi
28. Powdered CaCO<sub>3</sub> reacts much faster with dilute HCl than with the same mass of CaCO<sub>3</sub> as marble. Give reason. J20 3M COMPULSORY
29. The rate of formation of dimer in a second order reaction is  $7.5 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$  at 0.05 mol L<sup>-1</sup> monomer concentration. Calculate the rate constant. (BB) 5M i
30. For the general reaction  $\text{A} \rightarrow \text{B}$ . Plot of concentration of A Vs time is given in the graph below. Answer the following Questions on the basis of this graph. (214) PTA 3M



- i) What is the order of the reaction?
- ii) What is the slope of the curve?
- iii) What is the Unit of rate constant?

31. A first order reaction is 40% complete in 50 minutes. Calculate the value of the rate constant. In what time will the reaction be 80% complete? (BBQ<sub>30</sub>231) GM5Mi

**ALL INSIDE PROBLEM & Revise Book Back (Evaluation) Question Answer**

**PREPARED BY:**

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# **ACTC CHEMISTRY TUITION CENTRE**



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**ACTC ADVANCED CHEMISTRY TUITION CENTRE, NAGERCOIL, KK DIST 9940847892****PLAN!****PREPARE!!****PRESENTATION!!!****SSON 8 IONIC EQUILLIBRIUM**

- Mention the Arrhenius concept of acid and base (2)
- What are the limitations of Arrhenius concepts? (3) **M22 2M**
- Classify acid or base using Arrhenius concept.  $\text{HNO}_3$ ,  $\text{CH}_3\text{COOH}$ ,  $\text{Ba}(\text{OH})_2$ ,  $\text{H}_3\text{PO}_4$ (3)
- Define Lowery - Bronsted concept of acids and bases (3)
- 0.1 M Solution of HF is weak acid. But 5M solute ion of HF is stronger acid. Why? **PTA3M**
- What are conjugate acid – base pairs? Give example. (3) **PTA5M ii**
- Write a balanced equation for the dissociation of the following in water and identify the conjugate acid-base pairs. i)  $\text{NH}_4^+$  ii)  $\text{H}_2\text{SO}_4$  iii)  $\text{CH}_3\text{COOH}$  iv)  $\text{HCl}$  v)  $\text{HF}$  (4)
- Classify the following into Lewis acids and Lewis bases. **S20 5Mi**  
A)  $\text{BF}_3$  (B)  $\text{CO}_2$  (C)  $\text{MgO}$  (D)  $\text{CH}_3^-$
- What are Lewis acids and bases? Give two example (4) **M20 2M, J22 2M**
- Difference between Lewis acids and Lewis base. (5)
- Identify the Lewis acid and the Lewis base in the following reactions (5)  
 $\text{Cr}^{3+} + 6\text{H}_2\text{O} \rightarrow [\text{Cr}(\text{H}_2\text{O})_6]^{3+}$
- Identify the Lewis acid and Lewis base in the following reactions.(5)  
i)  $\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$  ii)  $\text{CH}_3\text{-O-CH}_3 + \text{AlCl}_3 \rightarrow (\text{CH}_3)_2\text{O-AlCl}_3$
- $\text{H}_3\text{BO}_3$  accepts hydroxide ion from water as shown below (6)  
 $\text{H}_3\text{BO}_3(\text{aq}) + \text{H}_2\text{O} \rightleftharpoons \text{B}(\text{OH})_4^- + \text{H}^+$   
Predict the nature of  $\text{H}_3\text{BO}_3$  using Lewis concept.
- Identify the conjugate Acid Base pair for the following reactions in aqueous solution.(30) **BB J20 5Mii**  
(A)  $\text{HS}^-(\text{aq}) + \text{HF} \rightleftharpoons \text{F}^-(\text{aq}) + \text{H}_2\text{S}(\text{aq})$   
(B)  $\text{HPO}_4^{2-} + \text{SO}_3^{2-} \rightleftharpoons \text{PO}_4^{3-} + \text{HSO}_3^-$
- How will you measure the strength of an acid? (6)
- Define ionic product of water. Give its value at room temperature (7) **S20, PTA 3MARK**
- Calculate the concentration of  $\text{OH}^-$  in a fruit juice which contains  $2 \times 10^{-3}\text{M}$ ,  $\text{H}_3\text{O}^+$  ion. Identify the nature of the solution(8)
- Derive the relationship between pH and pOH. (9,10) **PTA 5M i**
- Define pH (9) **M22 5M**
- Write the pH value of the following substances: **M20 5Mi**  
A) Vinegar B) Black coffee C) Baking Soda D) Soapy Water
- Calculate the pH of  $10^{-7}\text{M}$   $\text{HCl}$ .(11) **PTA3M**
- State Ostwald dilution law. Derive an expression Ostwald's dilution law. (12) **J20 3MARK, PTA 3MARK**
- A solution of 0.10M of a weak electrolyte is found to be dissociated to the extent of 1.20% at  $25^\circ\text{C}$ . Find the dissociation constant of the acid.(14)
- Calculate the pH of 0.1M  $\text{CH}_3\text{COOH}$  solution. Dissociation constant of acetic acid is  $1.8 \times 10^{-5}$ . (15)
- Define common Ion effect with an example (15) **J20 2M, PTA 2M, M22 5M ii**
- What are buffer solutions? Give an example (16) **J22 3M**
- What are the two types of buffer solution? Give example for each type.(16) **PTA5M ii**
- Explain the buffer action of a solution (16) **J20 2MARK**
- Explain buffer action of acidic buffer. (16) **PTA 3M**
- Derive Henderson-Hasselbalch equation(Derive Henderson equation) (18) **M20,GM 3M**

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- PLAN!** **PREPARE!!** **PRESENTATION!!!**
31. Find the pH of a buffer solution containing 0.20 mole per litre sodium acetate and 0.18 mole per litre acetic acid.  $K_a$  for acetic acid  $1.8 \times 10^{-5}$ . (19) **A21 3MARK**
32. Calculate the pH of 0.1M  $\text{CH}_3\text{COONa}$  solution ( $pK_a$  for  $\text{CH}_3\text{COOH}$  is 4.74)(20) **S20 2MARK**
33. Derive expression for hydrolysis constant and pH of salt of weak acid and strong base. (21) **5M**
34. Derive expression for hydrolysis constant and pH of salt of strong acid and weak base. (22)
35. Define solubility product (25)
36. Give a condition for a compound to be precipitated (25)
37. How will you calculate solubility product from molar solubility? (26) **PTA 5M ii**
38. Write the expression for the solubility product of  $\text{Ca}_3(\text{PO}_4)_2$ ,  $\text{BaSO}_4$ . (26)
39. Define solubility product of a compound. (25) **2M**
40. The  $K_a$  value of HCN is  $10^{-9}$ . What is the pH of 0.4M HCN solution? (31) **PTA 5M**
41. 50ml of 0.05M  $\text{HNO}_3$  is added to 50ml of 0.025M KOH. Calculate the pH of the resultant solution. (BBQ<sub>15</sub>31) **GM 2MC**
42.  $K_{SP}$  of  $\text{Ag}_2\text{CrO}_4$  is  $1.1 \times 10^{-12}$ . What is the solubility of  $\text{Ag}_2\text{CrO}_4$  in 0.1M  $\text{K}_2\text{CrO}_4$ ? (BBQ<sub>25</sub>31) **GM 5M ii**

**ALL INSIDE PROBLEM & Revise Book Back (Evaluation) Question Answer**

**LESSON 9 ELECTROCHEMISTRY**

- Define electrochemistry. (34)
- State Ohm's law. (34)
- A conductivity cell has two platinum electrodes separated by a distance 1.5 cm and the cross sectional area of each electrode is  $4.5 \text{ cm}^2$ . Using this cell, the resistance of 0.5N electrolytic solution was measured  $15 \Omega$ . Find the specific conductance of the solution. (36) **M20 2M**
- Define molar conductivity. (36)
- Define molar conductance and specific conductance How they are related? (36) **PTA 5M i**
- Define Equivalent conductance. (37) **A21 2M**
- What are the factors affecting electrolytic conductance. (37) **A21 2M, M22 3M**
- Explain measurement of conductivity of ionic solutions. (38)
- Why is AC current used instead of DC in measuring the electrolytic conductance? (38) (BBQ<sub>11</sub>66) **PTA 5M ii**
- Explain variation of molar conductivity with concentration. (39)
- Write Debye-Hückel and Onsager equation for a uni-univalent electrolyte. (41) **GM 5M ii**
- State Kohlrausch's law and applications. (41) **J20 5MARK, GM 5M ii**
- Describe the construction of Daniel cell and write its cell reaction. (45) **G3M**
- 
- What is the role of salt bridge in Galvanic cell? (46) **PTA 5M ii**
- What are the conventions used in Galvanic cell notation. (46) **PTA 5M i**
- Write note on standard hydrogen electrode (SHE). (48) **PTA 3 3M & PTA 5 5M i**
- Explain Thermodynamics of cell reactions. (50)
- Derive Nernst equation. (51) **S20 5M, M22 5M, J22 5M**
- Explain Electrolytic cell and electrolysis. (53)
- State Faraday's law of electrolysis First law, Second law. (54) **A21 3MARK GM 3M**

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22. A solution of silver nitrate is electrolyzed for 30 minutes with a current of 2 amperes. Calculate the mass of silver deposited at the cathode. (55) **S20 3MARK COM, J22 2M COM**
23. Write note on Leclanche cell. (56)
24. Write note on Mercury button cell. (57)
25. Write note on secondary batteries. (58)
26. Write note on fuel cell. (59)
27. Explain electrochemical mechanism of corrosion (60)
28. Write note on Electrochemical series. (62)
29. What are electrochemical series? How is it useful to predict corrosion? (62) **PTA 3M**
30. How are metals protected from corrosion by cathodic protection method? (61) **M20 3M**
31. The reaction  $Zn(s) + Co^{2+} \leftrightarrow Co(s) + Zn^{2+}$  occurs in a cell. Compute the standard emf of the cell. Given that  $E^\circ_{Zn/Zn^{2+}} = +0.76V$  and  $E^\circ_{Co/Co^{2+}} = +0.28V$ . **PTA 2MC**
32. Reduction potential of two metals  $M_1$  and  $M_2$  are  $E^\circ_{M_1^{2+}/M_1} = -2.3 V$  and  $E^\circ_{M_2^{2+}/M_2} = 0.2 V$ . Predict which one is better for coating the surface of iron. Given  $E^\circ_{Fe^{2+}/Fe} = -0.44 V$  (BBQ1766) **PTA 5M ii**
33. Is it possible to store copper sulphate in an iron vessel for a long time? Given  $E^\circ_{Cu^{2+}/Cu} = 0.34 V$  and  $E^\circ_{Fe^{2+}/Fe} = -0.44V$  (BBQ1566) **PTA 2M i**
34. Calculate  $\Lambda^\circ$   $CH_3COOH$  using appropriate molar conductance of the electrolytes listed below at infinite dilution at  $25^\circ C$  (BBQ563) **PTA 5M i**

Electrolyte	NaCl	KCl	$CH_3COONa$
$\Lambda^\circ$ (S $cm^2 mol^{-1}$ )	126.5	426.2	91.0

35. The equivalent conductance of M/36 solution of a weak monobasic acid is 6 mho  $cm^2 equiv^{-1}$  and at infinite dilution is 400 mho  $cm^2 equiv^{-1}$ . Calculate the dissociation constant of this acid. (BBQ1764) **PTA 2MC**

**ALL INSIDE PROBLEM & Revise Book Back (Evaluation) Question Answer**

### LESSON 10 SURFACE CHEMISTRY

- Define adsorption and absorption. (70)
- Characteristics of adsorption. (71) **PTA 2M**
- Distinction between chemical and physical adsorption. (71) **PTA 5Mi**
- Explain Factors affecting adsorption. (72)
- Write note on Adsorption and isobars. (73)
- Explain Freundlich adsorption isotherm and limitations. (73)
- Explain applications of adsorption. (75)
- Define catalyst and catalysis. (77)
- Define positive catalysis (77)
- Define homogenous catalysis & example (77) **M22 3M**
- Define heterogeneous catalysis & example (77) **J22 3M**
- What are the characteristics of catalysis? (78) **J20 5MARK, PTA 5Mi, M22 5M**
- Define promoters, catalytic poison with suitable example (79) **PTA 2M**
- Define auto catalysis and example (79)
- Identify the auto catalyst in the following reaction. (79) **J20 5Mi**  
 A)  $CH_3COOC_2H_5 + H_2O \rightarrow CH_3COOH + C_2H_5OH$     B)  $2AsH_3 \rightarrow 2As + 3H_2$
- Define negative catalysis and example (79)

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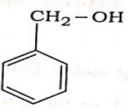
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17. Explain intermediate compound formation theory & limitations.(80) **GM 5Mi**
18. Explain adsorption theory of catalysis.(81) **A21 5MARK J22 5M**
19. What is the role of adsorption in the heterogeneous catalysis?(81) **PTA 3M**
20. What are active centres? (82) **PTA 2M**
21. Define, Mechanism, characteristics of Enzyme Catalysis. (83)
22. Explain Zeolite catalysis.(84)
23. What is Nano Catalysis? Give example.(86)**GM 5M i**
24. Define colloid & particle size.(86)
25. Define dispersed phase and dispersing medium. (86)
26. Define lyophilic colloids & example(87)
27. Define lyophobic colloids & example (87)
28. Explain the classification of colloids based on the physical state.(87)
29. Write the dispersed phase and dispersion medium of butter. (88) **M20 2MARK**
30. Write note on preparation of Colloids- mechanical dispersion.(88)
31. Write note on preparation of Colloids- electro dispersion.(89)
32. Write note on preparation of Colloids- Ultrasonic dispersion.(89)
33. Peptising agent is added to convert precipitate into colloidal solution. Illustrate with an example. 90 (BB103) **PTA 2M** (Write note on preparation of Colloids- peptisation.)
34. Write any three condensation methods of preparation of colloids. (Chemical method) (90) **PTA5M i**
35. What happens when hydrogen sulphide gas is passed through a solution of arsenic oxide? Name the chemical method. (90) **PTA 2M**
36. Write note on Dialysis.(91)
37. Write note on Electrodialysis.(91)
38. Write note on Ultrafiltration.(91) **GM5Mi**
39. Mention the shapes of the following colloidal particles. (93) **M20 3MARK**
  - i)  $As_2S_3$
  - ii) Blue gold sol
  - iii) Tungstic acid sol
40. Define Tyndall effect. (93) **20 2M**
41. Define Brownian movement (94)
42. What is the significance of Brownian movement? (94) **PTA 5M ii**
43. Write note on Helmholtz double layer.(94) **PTA 3M**
44. Explain Electrophoresis.(94) **PTA 5Mi, M22 2M**
45. Define electroosmosis. (95)
46. Define coagulation. Various method of coagulation. (96)
47. What is flocculation value? (96) **PTA 2M**
48. Define Gold number.(96)
49. Explain types, identification of Emulsions. (97)
50. What is inversion of phase? Give an example. (98) **S20 3M**
51. Explain various application of colloids (98)

**& Revise Book Back (Evaluation) Question Answer**

**ORGANIC CHEMISTRY LESSON 11 Hydroxy compounds and ethers**

- Classification of alcohol.(105)
- Write the IUPAC names of the following compounds. **PTA2M**
  - $C_6H_5 - O - CH_2 - CH - CH_3$  (134M)      ii)  $CH_2 = CH - CH_2 - CH_2OH$  (107M)
  - $$\begin{array}{c} | \\ CH_3 \end{array}$$
  - iii) Neopentyl alcohol (106)      iv) Glycerol (107)
- Give the IUPAC names: **M22 2M**

<p>(a) <math>\begin{array}{c} CH_3 \\   \\ CH_3 - C - OH \\   \\ CH_3 \end{array}</math></p>	<p>(b) </p>
--	--
- Write all the possible isomers of an alcohol having molecular formula  $C_5H_{12}O$ .(107)
- Structure of alcohol.(107)
- Preparation of primary, secondary, Tertiary alcohol(108)
- Propene to 2-propanol.(107)
- Formaldehyde to primary alcohol (Formaldehyde to ethanol).  
(Formaldehyde +  $C_6H_5MgBr \rightarrow ?$ ) (108)
- Acetaldehyde to secondary alcohol (Acetaldehyde to isopropyl alcohol)  
(Acetaldehyde +  $CH_3CH_2MgBr \rightarrow ?$ )(108)
- Acetone to Tertiary alcohol (Acetone to tert-butyl alcohol)  
(Acetone +  $CH_3CH_2CH_2CH_2MgBr \rightarrow ?$ .) (108)
- Ethyl methanoate to isopropyl alcohol. (108)
- How will you prepare the following by using Grignard reagent? **M20 5M ii**
  - Propan-1-ol
  - propan-2-ol
- Write note on Hydroboration.(109)
- Crotonaldehyde to crotyl alcohol. (109)
- Benzyl alcohol to benzaldehyde(109)
- What is Baeyer's reagent? How it is useful to convert ethane to ethane-1, 2-diol?**S20 5Mi**  
(What happens when ethylene reacts with cold dilute alkaline  $KMnO_4$ ?) (110)
- Write note on saponification. (110)
- Lucas test for primary, secondary, tertiary alcohol. (110) **J20 5MARK, S20 5M i**
- Victor Meyer test for primary, secondary, tertiary alcohol. (111)
- State Saytzeff's rule.(2, 3-dimethylpentan-3-ol)114
- Swern oxidation.(117) **PTA 5M i**
- Biological oxidation(118)
- Esterification.(118)
- Oxidation primary, secondary alcohol, Tertiary alcohol.(117)
- Is it possible to oxidize t-butyl alcohol using acidified dichromate to form a carbonyl compound? (BB142) **PTA2M**
- $Cu/573K$  of primary, secondary alcohol, Tertiary alcohol.(118)

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27. Glycol to ethene, Glycol to dinitroglycol. (119)
28. Glycol to oxirane, glycol to acetaldehyde. **A21 3M i** (119)
29. Write the reaction of ethylene glycol with  $\text{Con H}_2\text{SO}_4$ ? (120)
30. Oxidation of Ethylene glycol. (120)
31. Write the chemical equation for oxidation of ethylene glycol with periodic acid. (120) **PTA 2M**
32. Glycerol to TNG. (121)
33. What happens when glycerol react  $\text{KHSO}_4$ ? (121) (Glycerol to Acrolein) **A21 3Mi GM 3M PTA 5Mi**
34. Oxidation of glycerol (121)
35. What is meant by glycerose. (121)
36. Uses of glycerol. (122) **J22 2M**
37. Preparation of phenol  
From chlorobenzene – Dow process, From benzene sulphonic acid, From aniline, From benzene or cumene
38. Chemical properties of phenol  
 $\text{Zn}$ ,  $\text{NH}_3$  / Anhydrous  $\text{ZnCl}_2$ ,  $\text{CH}_3\text{COCl}$ ,  $\text{C}_6\text{H}_5\text{COCl}$ ,  $\text{NaOH}/\text{CH}_3$ , Oxidation - Acidified  $\text{K}_2\text{Cr}_2\text{O}_7$ , Reduction –  $\text{Ni}/160^\circ\text{C}$ , Nitrosation –  $\text{HNO}_2/278\text{K}$ , 20%  $\text{HNO}_3/298\text{K}$ , Conc  $\text{HNO}_3/\text{Conc H}_2\text{SO}_4$ , Sulphonation – Conc.  $\text{H}_2\text{SO}_4$ ,  $\text{Br}_2/\text{H}_2\text{O}$ ,  $\text{Br}_2/\text{CCl}_4/278\text{K}$ , Kolbe (or) Kolbe schmit reaction,
39. Riemer-tiemann reaction, **PTA 5M i** Phthalein reaction.
40. Schotten – Baumann reaction. (127) **PTA 5M i**
41. Test to differentiate alcohol and phenols, Uses of phenol
42. Give the coupling reaction of phenol. (131) **M20 5Mi**
43. How the following conversions are effected? **GM 5M**  
i) phenol to salicylaldehyde (130)  
ii) phenol to phenolphthalein (131)  
iii) Glycol to 1,4 dioxane (120)
44. Acidity of phenol. (124)
45. Why is C – O – C bond angle in ether slightly greater than the tetrahedral bond angle? (133) **M20 2MARK COMPULSORY**
46. Preparation of ether (134) Write any one method of preparation for diethyl ether? **M22 3M**
47. Chemical properties of ether. (136)
48. Explain auto oxidation of ethers. (137) **S20 5M ii**
49. Mention the mechanism in the following reactions: (137) **J20 3MARK**  
(a) One mole of HI reacts with methoxy ethane **ANS:  $\text{SN}^2$**   
(b) One mole of HI reacts with 2-methoxy 2-methyl propane **ANS:  $\text{SN}^1$**
50. Electrophilic substitution reactions. (137)
51. Give the uses of diethyl ether. (138) **A21 5Mii**
52. A ether (A)  $\text{C}_6\text{H}_{12}\text{O}$  when heated with excess of hot concentrated HI, produced two alkyl halides, which on hydrolysis forms compound (B) and (C). Oxidation Of (B) gives an acid (D) where as oxidation of (C) gives ketone (E). Identify A, B, C, D and E and write the chemical equation. **PTA 5M**
53. Anisole  $\xrightarrow{t\text{-butylchloride AlCl}_3}$  A  $\xrightarrow{\text{Cl}_2/\text{FeCl}_3}$  B  $\xrightarrow{\text{HBr}}$  C. Complete the above reaction and Find A, B, C. (BB142) **PTA 3M**

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54. dehydration of glycerol (121) **5M ii**
55. How will you prepare 2-methyl hexan-2-ol from Grignard reagent? (108) **3M**
56. Write the mechanism of acid catalysed dehydration of ethanol to give ethene. (115) **5M i**
57. What are the tests to differentiate ethanol and phenols? (131) **3M**
58. An organic compound (A) –  $C_3H_8O_3$  used as a sweetening agent, which on oxidation with Fenton's reagent gives a mixture of compounds B and C. Identify A, B and C. Write Possible reactions. **3MC(121,122)**
59. Give four uses of diethyl ether. (138) **2M**
60. What will be the product (X and A) for the following reaction? **5M ii** (BBQ<sub>22</sub>143)
- $$\text{Acetyl chloride} \xrightarrow{CH_3MgBr/H_3O^+} X \xrightarrow{\text{acidic}/K_2Cr_2O_7} A$$
61. The major product formed when 1-ethoxy prop-1-ene is heated with one equivalent of HI (BBQ<sub>2</sub>142) **PTA 3Mi**
62. What happens when 1-phenyl ethanol is treated with acidified  $KMnO_4$ ? (BBQ<sub>9</sub>142) **PTA 3Mii**
63. An organic compound  $C_2H_6O$  (A) heated with  $Con H_2SO_4$  at 443K to give an unsaturated hydrocarbon  $C_2H_4$  (B), which on treatment with Bayer's reagent to give compound  $C_2H_6O_2$  (C) which is used as antifreeze in automobile radiator. Compound (C) distilled with  $con H_2SO_4$  to give cyclic compound  $C_4H_8O_2$  (D). Compound (D) is heated with  $Con H_2SO_4$  at 413K to give compound  $C_4H_{10}O$  (E). Identify Compounds (A) to (E) and write equations. **PTA 5M**
64. How the following conversions are effected? **5M**
- phenol to salicylaldehyde (130)
  - phenol to phenolphthalein (131)
  - Glycol to 1,4 dioxane (120)

**LESSON 12 Carbonyl compounds and carboxylic acid**

- How will you prepare ethanal by ozonolysis? (149)
- How are the following conversions effected? **PTA 2M**
  - Hex-3-yne  $\rightarrow$  hexan-3-one (150 model)
  - benzaldehyde  $\rightarrow$  2-hydroxy phenyl acetic acid. (BBQ<sub>8</sub>193)
- Rosenmund reduction (151) **J22 2M**
- Name the catalyst used in Rosenmund reduction and state its importance. (151) **M20 2M**
- How is the following conversion affected? Hex-4-enitrile  $\rightarrow$  hex-4-enal (151) **PTA 5M ii**
- Stephen's reaction (151)
- Etard reaction (151) **PTA 3M**
- Gattermann Koch reaction (151)
- Friedel crafts acylation (151)
- How is benzaldehyde manufactured commercially? (152)
- How are the following conversions affected? **5M ii**
  - (X) Benzene  $\rightarrow$  acetophenone (153) (Y) Benzaldehyde  $\rightarrow$  hydrobenzamide (159)
- How will you prepare aldimine? (158)
- What is Urotropine? How it is prepared? and uses (158) **J20 2M**
- Popoff's rule (159)
- Clemmensen reduction (160)
- Wolf kishner reduction (161)

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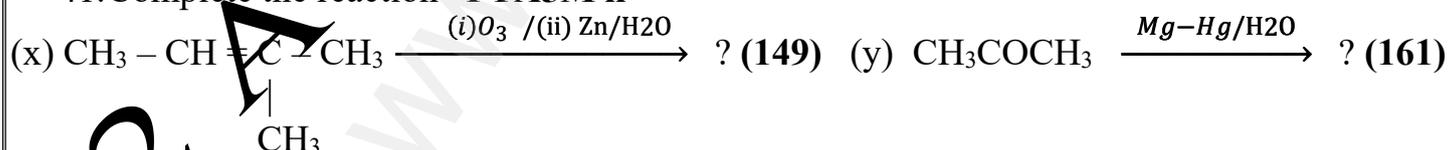
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17. Haloform reaction (161) **M22 3M**  
 18. Crossed aldol condensation (162)  
 19. What happens on heating of aldol? (161)  
 20. Explain Aldol condensation with mechanism (161) **S20 5M GM 3M**  
 21. Explain Cannizzaro reaction with mechanism (163) **PTA 3M**  
 22. Claisen Schmidt condensation (163)  
 23. Crossed cannizzaro reaction (164)  
 24. Perkin's reaction (165)  
 25. Knoevenagel reaction (165) **PTA 3Mi**  
 26. How will you convert benzaldehyde into the following compounds? (165) **A21 5M**  
 (i) Benzoin **J22 5Mii** (ii) Cinnamic acid (iii) Malachite green **PTA 2M**  
 27. Note on Schiff's base (165)  
 28. Test for aldehyde (166)  
 29. Explain Benedict's solution test. (167) **S20 3M**  
 30. What is Formalin? What is its use? (167) **M20 5Mi**  
 31. How will you prepare benzoic acid from toluene? (171) **PTA 3M**  
 32. What happens when ethanoic acid reacts with ethanol in the presence of con  $H_2SO_4$ . Give its complete mechanism. (173) **5M i**  
 33. How does sodium salt react with soda lime? (175)  
 34. HVZ reaction (176)  
 35. Formic acid reduces Tollens reagent whereas acetic acid does not reduce. (177) **M20 3M**  
 Explain the reducing action of formic acid with example. (177) **M22 5M**  
 36. uses of formic acid (188)  
 37. Test for carboxylic acid (177) **J20 3M, A21 3M, J22 5Mi**  
 38. Why formic acid act as strong reducing agent? Give one equation to show its reducing property. (177) **2M**  
 39. Identify A and B (by bond line structure) **J20 5Mi**



40. Effect of substituents on the acidity of carboxylic acid. (178)

41. Complete the reaction **PTA 5M ii**



42. Arrange the following in the increasing order of relative reactivity of acid derivatives and mention the reason alone.  $CH_3COOC_2H_5$ ,  $CH_3COCl$ ,  $CH_3CONH_2$ ,  $CH_3COOCOCH_3$  (180)  
**J20 5Mii**

43. A carbonyl compound A having molecular formula  $C_5H_{10}O$  forms crystalline precipitate with sodium bisulphate and gives positive iodoform test. A does not reduce Fehling's solution. Identify 'A'. (BBQ<sub>10</sub>193) **PTA 5M ii**

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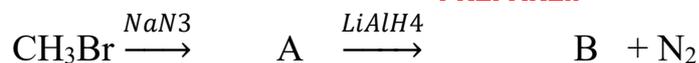
44. An organic compound  $C_3H_4$  (A) on hydration with  $Hg^{2+} / H_2SO_4$  gives compound (B) which gives positive iodoform test. Compound (B) heated with  $NH_2 - NH_2 / C_2H_5ONa$  to give hydrocarbon (C). (B) also treated with HCHO in the presence of dil NaOH gives compound (D). Identify A, B, C and D. Write the chemical reactions involved. **5M**
45. An organic Compound (A)  $-C_2H_4O$  reduces Tollen's and fehling's solution. A-react with methanol and HCl to give compound (B)  $- C_4H_{10}O_2$ . A-on reaction with Methanal in the presence of dilute NaOH to give compound (C)  $- C_3H_6O_2$ . Identify Compounds A, B and C with necessary reactions. (155,163) **PTA5M ii**
46. An organic Compound  $C_2H_5Br$  (A) on treatment with Mg in dry ether gives (B) which on treatment with  $CO_2$  followed by acidification gives (C). Identify (A), (B) & (C) and write possible equations. **PTA3MC**
47. Compound A of molecular formula  $C_7H_6O$  reduces Tollen's reagent when A reacts with 50% NaOH gives compound B of molecular formula  $C_7H_8O$  and C of molecular formula  $C_7H_5O_2Na$ . compound C on treatment with dil HCl gives compound D of molecular formula  $C_7H_6O_2$ . When D is heated with sodalime gives compound E. identify A,B,C,D & E. Write the corresponding equations. (163) **GM 5M**

**LESSON 13 Organic Nitrogen Compounds**

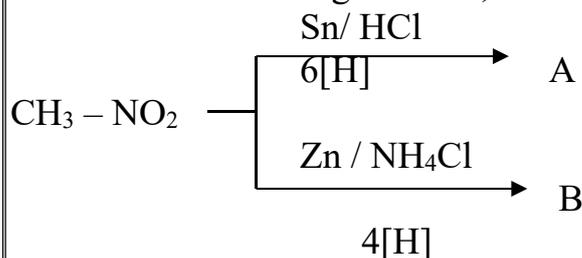
1. Explain Isomerism of nitro compound? (199)
2. There are two isomers with the formula  $CH_3NO_2$ . How will you distinguish between them? (199) **PTA3M**
3. Acidic nature of nitro alkanes. (200)
4. How will you prepare oil of mirbane? (201)
5. Ethyl nitrite to ethanol. (202)
6. How is Chloropicrin prepared (203) **M20 2 MARK**
7. Hofmann's bromide reaction (209)
8. Gabriel phthalimide synthesis (209) **J20 3M, M22 5Mii**
9. Hoffmann's ammonolysis (209)
10. Sabatier – Mailhe method (210)
11. Schotten – Baumann reaction (214)
12. Diazotisation (215)
13. Libermann's nitroso test. (215)
14. Carbylamine reaction (216) **M22 5Mi**
15. Mustard oil reaction (216) **S20 5Mii, J22 5Mii**
16. Hofmann-Mustard oil reaction. (216)  
(How will you prepare phenyl mustard oil?)
17. How does aniline react with  $Br_2/H_2O$  (Bromination of aniline)? (217) **J22 5Mi**
18. Why aniline does not undergo Friedel Crafts reaction. (218) **J20 2M**
19. Identify A and B. (208) **M20 5Mii**

$$A \xrightarrow{Na(Hg)/C_2H_5OH} CH_3 - CH_2 - NH_2$$

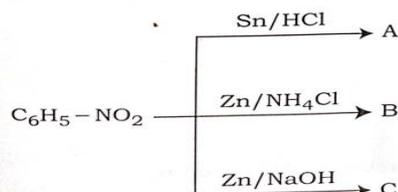
$$B \xrightarrow[4[H]]{Na(Hg)/C_2H_5OH} CH_3 - NH - CH_3$$
20. Identify A and B. (208) **M22 2M Compulsory**

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21. From the following reaction, identify A and B. (202) **A21 2M Compulsory**



22. From the following reaction, identify A, B and C. (203) **J22 3M Compulsory**



23. How will you convert nitrobenzene into (203)

Aniline, Phenyl hydroxyl amine, Nitrosobenzene, Azobenzene, azoxybenzene, hydrozobenzene, m - nitro aniline (204), 1,3,5 - trinitrobenzene (204), 3-nitro benzene sulphonic acid, 3-chloro nitro benzene.

24. Name the reducing agent used in the reduction of nitrobenzene to the following compounds.

(203) **S20 5Mi**

(A) Aniline

(B) Phenyl hydroxylamine

(C) Nitrosobenzene

25. How will you distinguish between primary secondary and tertiary aliphatic amines?

26. Arrange the following: (212BB234) **PTA 2M**

(i) In decreasing order of the  $pK_b$  values:  $\text{C}_2\text{H}_5\text{NH}_2$ ,  $\text{C}_6\text{H}_5\text{N}(\text{CH}_3)_2$ ,  $(\text{C}_2\text{H}_5)_2\text{NH}$ ,  $\text{CH}_3\text{NH}_2$

(ii) Increasing order of basic strength:  $\text{C}_2\text{H}_5\text{NH}_2$ ,  $\text{C}_6\text{H}_5\text{N}(\text{CH}_3)_2$ ,  $(\text{C}_2\text{H}_5)_2\text{NH}$ ,  $\text{CH}_3\text{NH}_2$

27. How the following conversion are effected? **PTA 5M**

(i) Nitro benzene  $\rightarrow$  N-phenyl hydroxyl amine (203)

(ii) Propanamide  $\rightarrow$  Propan - 1 - amine (209)

(iii) Aniline  $\rightarrow$  p - nitroaniline (218)

28. An organic compound (A) on reduction gives compound (B). (B) on treatment with  $\text{CHCl}_3$  and alcoholic  $\text{KOH}$  gives (C). (C) on catalytic reduction gives N - methyl aniline. Identify A,B,C and write its equation. (216 Carbylamine) **PTA 3M**

29. Account the following **PTA 3M (BBQ8234)**

i) Aniline does not undergo Friedal - crafts reaction

ii) Ethylamine is soluble in water whereas aniline is not

iii) Amines are more basic than amides.

30. An organic compound (A) -  $\text{C}_7\text{H}_7\text{NO}$  on treatment with  $\text{Br}_2$  and  $\text{KOH}$  gives an amine (B), which gives carbylamines test. (B) upon diazotization to give (C). (C) on coupling with P. cresol to give compound (D). Identify A,B,C and D with necessary reaction. (209) **PTA 5M**

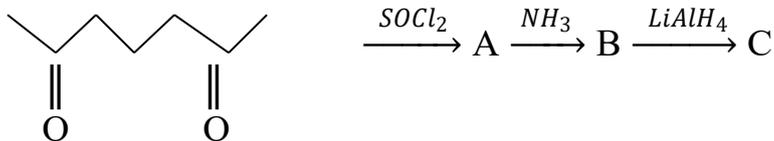
31. An organic compound (A) -  $\text{CNCl}$  react with methyl magnesium Bromide to give compound B -  $(\text{C}_2\text{H}_3\text{N})$ . B-upon catalytic reduction to give compound C -  $(\text{C}_2\text{H}_7\text{N})$ . C gives carbylamine test. Identify compound A,B and C and write the reactions. (224, 225) **3MC**

32. Identify A,B and C (BBQ13235) **PTA 5M i**



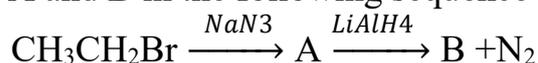
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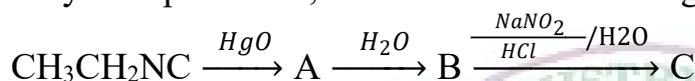
33. An aromatic nitro compound (A) on reduction with Sn/HCl gives compound (B)  $\text{C}_6\text{H}_7\text{N}$ , which on treatment with Benzoyl chloride in the presence of pyridine to give compound (C). Compound (B) on treatment with  $\text{CH}_3\text{Br}$  to give compound (D) which further reacts with  $\text{NaNO}_2/\text{HCl}$  to give compound (E) with yellow oil liquid. Identify (A) to (E) and write the reactions. **PTA 5M**

34. Identify A and B in the following sequence of reactions. (210) **M22 2M COMPULSORY**



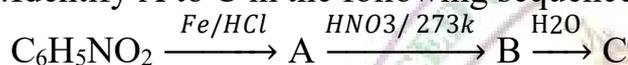
35. Write a note on Sabatier – mailhe method? (210) **PTA 2M**

36. Identify Compounds A, B and C in the following sequence of reaction. (BBQ<sub>51</sub>233) **3MC**



37. Write the uses of nitroalkanes. (228) **PTA 5M ii**

38. Identify A to C in the following sequence? (BBQ<sub>51</sub>233) **GM 3M**



39. write short notes on **GM 5M** a) Mustard oil reaction (210) b) Carbylamines reaction (216) c) Gabriel phthalamide synthesis. (209)

**LESSON 14 Biomolecules**

- Outline the classification of carbohydrates giving example for each (239)
- What are the different types of monosaccharides (240)
- Elucidate the structure of glucose (241)
- Draw the cyclic structure of glucose (243) (Write the structure of  $\alpha - \text{D}(+) \text{Glucopyranose}$ .) (243) **2M**
- Define anomer. Give example. (243)
- Define mutarotation. (244)
- Define epimers and epimerization. (244) What are epimers? Give an example. **M22 3M**
- Explain the structure of Fructose. (245) **GM5M**
- Draw the structure of  $\text{D}(+) \text{Fructose}$ . (246) **J22 2M**
- What happens when fructose is partially reduced with sodium amalgam and water? (245) **PTA 5M ii**
- Draw the cyclic structure of fructose. (246)
- Write about the structure of sucrose (247)
- Explain the structure of lactose (247)
- Explain the structure of maltose (248)
- What is glycosidic linkage? (247) **M20 5Mi**
- Mention the importance of Carbohydrates (250)
- Define isoelectric point (252)
- What is Called Zwitter ion? Give an example (252) **J22 3M**
- Write a short note on peptide bond (252)
- Write a short note on peptide bond. (252) **5M i**
- Write the Zwitter ion structure of alanine. (252) **2M**
- Write a note on denaturation of proteins. (256) **S20 2M**

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23. Write any three biological importance of lipids. (256) **S20 3M**  
 24. Name the vitamins whose deficiency causes. (a) Rickets (b) Scurvy (259) **J20 2M**  
 25. Explain composition and structure of nucleic acids. (260)  
 26. Explain types of RNA molecules (264) **J20 5Mi**  
 27. Give any four differences between DNA and RNA. (264) **A21 3M GM2M**  
 28. Write note on DNA finger printing. (265)  
 29. Explain Hormones (266).

**LESSON 15 CHEMISTRY IN EVERYDAY LIFE**

1. Give a brief account on Antioxidants. (283) **J20 5Mi**  
 2. How do you classify the following into various class of drugs? (277-282) **J20 5Mii**  
 (A) Milk of Magnesia (B) Aspirin (C) Penicillin (D) Procaine  
 3. How do antiseptics differ from disinfectants? (282) **S20 5Mi**  
 4. What are bio degradable polymers? Give two examples. (293) **PTA 3M**  
 5. How the tranquilizers work in body? 277 (BB 296) **PTA 2M**  
 6. Write a note on TFM value. ((284) **PTA 5M ii**  
 7. How nylon -6 is prepared? (289) **PTA 2M**  
 8. Write the made of action and uses of antacids. Give an example (279) **PTA 5M ii**  
 9. What are food preservatives? (283) **PTA 2MC**  
 10. Define food additives. (282) State any three advantages of food additives. (283) **M20 3M**  
 11. Define the term therapeutic index. How is it related to the safety of the drug? (273) **PTA 5M ii**  
 12. How is terylene prepared? (290) **PTA 2M**  
 13. Explain the mechanism of cleaning action of soaps and detergents. (284) **PTA 5M i**  
 14. What are Antiseptics? Give an example. (283) **PTA 2M**  
 15. How polymers are classified on the basis of structure and molecular forces, give examples of each one. (286) **PTA 5M**  
 16. Write short notes on Antioxidants. (283) **G2M**  
 17. Explain the preparation of Nylon-6,6 and Buna-S. (289, 292) **G3M**  
 18. What is Vulcanization? (292) **M20 5Mii**  
 19. How is neoprene prepared? **S20 3M**

**2M- 2 marks; 2MC- 2 marks Compulsory; 3M- 3 marks; 5M- 5 marks; GM2M- Govt model question paper 2 marks;**

**ALL THE BEST SCORE MORE MARKS**

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## NAME REACTIONS ORGANIC CHEMISTRY

### LESSON 11

1. Markonikoff's rule. (108)
2. Grignard reagent(108)
3. Hydroboration(109)
4. Bayer's reagent. (110)
5. Saponification(110)
6. Lucas test (111)
7. Victor meyer's test(111)
8. Saytzeff's rule.(116)
9. Swern oxidation.(117)
10. Esterification(118)
11. Biological oxidation(118)
12. TNG(121)
13. Acrolein (121)
14. Dows process(127)
15. Schotten- baumann reaction(128)
16. Williamson ether synthesis(128)
17. Kolbe's (or) Kolbe's schmit reaction(131)
18. Riemer tiemann reaction(131)
19. Phthalein reaction(132)
20. Coupling reaction(132)
21. Friedel craft's reaction(139)

### LESSON 12

22. Ozonolysis(149)
23. Rosenmund reduction(151)
24. Stephen's reaction(151)
25. Gattermann koch reaction(152)
26. Friedel crafts acylation (152)
27. Urotropine(158)
28. Popoff's rule (159)
29. Clemmensen reduction(160)
30. Wolfkisher reduction(161)
31. Aldol condensation(162)
32. Crossed aldol condensation(163)
33. Claisen-schmidt condensation(163)
34. Cannizaro reaction(164)
35. Crossed cannizaro reaction(164)

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36. Benzoin condensation(164)
37. Penkings reaction(165)
38. Knoevenagal reaction(165)
39. Schiff's base(165)
40. Malachite green dye(165)
41. Tollens reagent test(166)
42. Fehlings solution test(166)
43. Benedicts solution test(167)
44. Schiff's reagent test(167)
45. Esterification (173)
46. Kolbs electrolytic(175)
47. decarboxylation(175)
48. Trans esterification(185)
49. Claisen condensation(186)
50. Hoffmanns degradation(188)

**LESSON-13**

51. Chloropicrin (203)
52. Mendius reaction (208)
53. Gabriel phthalimide synthesis(209)
54. Hoffmanns ammonolysis(209)
55. Sabatier –Mailhe method(210)
56. Schotten –baumann reaction(214,215)
57. Carbylamine reaction(216)
58. Mustard oil reaction(216)
59. Gattermann reaction(220)
60. sandmeyer reaction(220)
61. Baltz –schiemann reaction(221)
62. Gomberg reaction(221)
63. Coupling reaction(223)
64. Condensation reaction(225)
65. Carbylamins reaction(226)
66. Levine and hawser acetylation(225)
67. Libermanns nitroso test(215)

**ORGANIC PROBLEM**

1. An organic compound (A) of molecular formula  $C_3H_8O$  gives turbidity within 5-10 minutes on reaction with anhydrous  $ZnCl_2/HCl$ . Compound (A) on treatment with PCC (Pyridinium chloro chromate) gives a carbonyl compound(B) which on further chlorination gives compound (C) of molecular formula  $C_3H_3OCl_3$ . Identify (A),(B) and (C) and explain the reactions. (117)
2. An organic compound (A)  $C_2H_6O$  liberates hydrogen on treatment with metallic sodium. (A) on mild oxidation gives (B)  $C_2H_4O$  which answers iodoform test. (B) when treated with  $CH_3OH/HCl$  to give (C) ( $C_4H_{11}O_2$ ). Identify (A), (B) and (C) and explain the reactions. (155)
3. An organic compound A of molecular formula  $C_6H_6O$  gives a violet colourisation with neutral  $FeCl_3$ . Compound A on treatment with  $NaOH$  gives compound B. Compound B on

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- treatment with  $\text{CO}_2$  at 400 K under pressure gives C. This product on acidification gives compound D ( $\text{C}_7\text{H}_6\text{O}_3$ ) which is used in medicine. Identify A, B, C and D and explain the reactions. (130 Kolbe reaction)
- An organic compound A of molecular formula  $\text{C}_3\text{H}_6\text{O}$  on reduction with  $\text{LiAlH}_4$  gives B. Compound B gives blue colour in Victor Mayer's test and also forms a chloride C with  $\text{SOCl}_2$ . (A) reacts with  $\text{NH}_3$  to form D ( $\text{C}_6\text{H}_{13}\text{NO}$ ). Identify A, B, C and D and explain the reactions. (158)
  - An organic compound (A)  $\text{C}_3\text{H}_8\text{O}$  answers Lucas test-within 5-10 minutes and on oxidation forms B ( $\text{C}_3\text{H}_6\text{O}$ ). This on further oxidation forms C ( $\text{C}_2\text{H}_4\text{O}_2$ ) which gives effervescence with  $\text{Na}_2\text{CO}_3$ . B also undergoes iodoform reaction. Identify A, B, and C. Explain the conversion of A to B and C.
  - An organic compound (A) of molecular formula  $\text{C}_6\text{H}_6\text{O}$  gives violet colour with neutral  $\text{FeCl}_3$ . (A) react with  $\text{CHCl}_3/\text{aq. NaOH}$  to gives B. (A) also reacts with  $\text{C}_6\text{H}_5\text{NO}$  to give the compound (C) which is a red orange dye. Identify (A), (B) and (C). Explain with suitable reactions. (130)
  - Compound (A) of molecular formula  $\text{C}_3\text{H}_8\text{O}$  liberates hydrogen with sodium metal. (A) with  $\text{P/I}_2$  gives (B). Compound (B) on treatment with silver nitrate gives (C) which gives blue colour with nitrous acid. Identify (A), (B), (C) and explain the reactions. (112)
  - Compound (A) with molecular formula  $\text{C}_6\text{H}_6\text{O}$  gives violet colour with neutral  $\text{FeCl}_3$ , reacts with  $\text{CHCl}_3$  and  $\text{NaOH}$  gives (B) with molecular formula  $\text{C}_7\text{H}_6\text{O}_2$ . Compound (A) reacts with Ammonia at 473 K in the presence of  $\text{ZnCl}_2$  and gives compound (C) with molecular formula  $\text{C}_7\text{H}_7\text{N}$ . Compound (D) undergoes carbylamine test. Identify (A), (B), and (C). Explain the reactions. 127
  - An organic compound  $\text{C}_2\text{H}_6\text{O}$  (A) reacts with  $\text{H}_2\text{SO}_4$  at 443 K and gives (B) of molecular formula  $\text{C}_2\text{H}_4$ . (B) Reacts with cold alkaline  $\text{KMnO}_4$  (Baeyer's reagent) to give (C) of molecular formula  $\text{C}_2\text{H}_6\text{O}_2$ . Identify (A), (B), (C). Explain the reactions. (115 & 110)
  - An organic compound A ( $\text{C}_2\text{H}_6\text{O}$ ) liberates hydrogen with sodium metal. A when heated with alumina at 620 K gives an alkene B which when passed through Bayer's reagent gives C ( $\text{C}_2\text{H}_6\text{O}_2$ ). C reacts with  $\text{Fe}_3$  and gives back B. Identify A, B and C. Write the reactions. (110, 115, 119)
  - A ether (A)  $\text{C}_5\text{H}_{12}\text{O}$  when heated with excess of hot concentrated HI, produced two alkyl halides, which on hydrolysis forms compound (B) and (C). Oxidation Of (B) gives an acid (D) where as oxidation of (C) gives ketone (E). Identify A, B, C, D and E and write the chemical equation.
  - An organic compound (A) on reduction gives compound (B). (B) on treatment with  $\text{CHCl}_3$  and alcoholic  $\text{KOH}$  gives (C). (C) on catalytic reduction gives N – methyl aniline. Identify A, B, C.
  - An organic compound  $\text{C}_3\text{H}_4$  (A) on hydration with  $\text{Hg}^{2+} / \text{H}_2\text{SO}_4$  gives compound (B) which gives positive iodoform test. Compound (B) heated with  $\text{NH}_2 - \text{NH}_2 / \text{C}_2\text{H}_5\text{ONa}$  to give hydrocarbon (C). (B) also treated with  $\text{HCHO}$  in the presence of dil  $\text{NaOH}$  gives compound (D). Identify A, B, C and D. Write the chemical reactions involved.
  - An organic compound (A) –  $\text{C}_3\text{H}_8\text{O}_3$  used a sweetening agent, which on oxidation with Fenton's reagent gives a mixture of compounds B and C. Identify A, B and C. Write Possible reactions.
  - An organic compound (A) –  $\text{C}_7\text{H}_7\text{NO}$  on treatment with  $\text{Br}_2$  and  $\text{KOH}$  gives an amine (B), which gives carbylamine test. (B) upon diazotization to give (C). (C) on coupling with P. cresol to give compound (D). Identify A, B, C and D with necessary reaction.

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16. An organic Compound (A)- $C_2H_4O$  reduces Tollen's and fehling's solution. A-react with methanol and HCl to give compound (B) –  $C_4H_{10}O_2$ . A-on reaction with Methanal in the presence of dilute NaOH to give compound (C) –  $C_3H_6O_2$ . Identify Compounds A, B and C with necessary reactions.
17. An organic Compound  $C_3H_5Br$  (A) on treatment with Mg in dry ether gives (B) which on treatment with  $CO_2$  followed by acidification gives (C). Identify (A), (B) & (C) and write possible equations.
18. An aromatic nitro compound (A) on reduction with Sn/HCl gives compound (B)  $C_6H_7N$ , which on treatment with Benzoyl chloride in the presence of pyridine to give compound (C). Compound (B) on treatment with  $CH_3Br$  to give compound (D) which further reacts with  $NaNO_2/HCl$  to give compound (E) with yellow oil liquid. Identify (A) to (E) and write the reactions.
19. An organic compound  $C_2H_6O$  (A) heated with Con  $H_2SO_4$  at 443K to give an unsaturated hydrocarbon  $C_2H_4$  (B), which on treatment with Bayer's reagent to give compound  $C_2H_6O_2$  (C) which is used as antifreeze in automobile radiator. Compound (C) distilled with con  $H_2SO_4$  to give cyclic compound  $C_4H_8O_2$  (D). Compound (A) is heated with Con  $H_2SO_4$  at 413K to give compound  $C_4H_{10}O$  (E). Identify Compounds (A) to (E) and write equations.
20. Compound A of molecular formula  $C_7H_6O$  reduces Tolle's reagent when A reacts with 50% NaOH gives compound B of molecular formula  $C_7H_8O$  and C of molecular formula  $C_7H_5O_2Na$ . compound C on treatment with dil HCl gives compound D of molecular formula  $C_7H_6O_2$ . When D is heated with sodalime gives compound E. identify A,B,C,D & E. (163) GM5M

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**PREPARED BY:**

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