SLIP TEST NO : 02

CHEMISTRY

CLASS :12
TIME : 1.30 hrs

Total Marks : 35

PART – I

Answer any 5 questions   **5X2=10**
1. Define Unitcell?
2. Give any two examples for zero order reactions?
3. What are co-valent solids? Give one example.
4. What is first order reaction.
5. Define packing efficiency.
6. What are primitive efficiency and non-primitive unit cells.
7. Define co-ordination number?
8. What are basic reactions?

PART – II

Answer any 5 questions   **5X3=15**
9. Write Bragg's equation and explain the terms?
10. Write the properties of Ionic crystals?
11. Define Half-life period?
12. Write the Arrhenius equation and explain the terms?
13. What are F - Centres?
14. Differentiate between order and molecularity.
15. What are molecular crystals and mention the types of molecular crystals?
16. Write note on pseudo first order reaction?

PART – III

Answer the following questions   **2X5=10**
17. a) Calculate the packing efficiency of body centered cubic unit cell?
                    (Or)
    b) i) Differentiate between crystalline and amorphous solids (any three) (3)
       ii) The rate constant for a first order reaction is $1.54X10^{-3}S^{-1}$
       Calculate its half life time.(2)

18. a) Derive the Integrated rate law equation for a first order reaction A$\rightarrow$ Products.
                    (Or)
    b) i) Write note on schottky defect? (3)
       ii) Calculate the number of atoms per unit cell in fcc unit cell. (2)