TIRUVANNAMALAI DISTRICT-2023-2024
MHS-GIRITHARAN PETTAI
SCIENCE-IX
REMEDIAL TEACHING TEST

LN-1.MEASUREMENT   **TEST-1**

I. Choose the correct answer.   (4×1=4)

1. Choose the correct one.
   a. mm< cm < m < km   b. mm > cm > m > km   c. km < m < cm < mm   d. mm > m> cm> km

2.1 metric ton is equal to
   a. 100 quintals   b. 10 quintals   c. 1/10 quintals   d. 1/100 quintals

3. Metre is the unit of _____
   a) Mass   b) Length   c) Time   d) Temperature

4.Thickness of a cricket ball is measured by_____
   a) Screw gauge   b)Vernier Caliper   c) Light year   d) Parsec

5.**State whether true or false**   (3×2=6)
   a)The SI unit of electric current is kilogram.
   b)Kilometre is one of the SI units of measurement.

6.Match the following.
   a) Length   -   kelvin
   b)Mass   -   metre
   c)Time   -   kilogram
   d)Temperature   -   second

7.What is the full form of SI system?

LN-1.MEASUREMENT   **TEST-2**

I. Choose the correct answer.   (4×1=4)

1.Rulers, measuring tapes and metre scales areused to measure
   a. mass   b. weight   c. time   d. length

2.Which among the following is not a device to measure mass?
   a. Spring balance   b. Beam balance   c. Physical balance   d. Digital balance

3.Radius of a thin wire is measured by _____
   a. Screw gauge   b. Vernier caliper   c.Digital balance   d. Beam balance

4.1 kg of rice is weighed by _____
   a. Beam balance   b. Physical balance   c.Spring balance   d. Digital balance

5.**Match the following**.   (3×2=6)
   a) Screw gauge   -   Vegetables
   b) Vernier caliper   -   Coins
   c) Beam balance   -   Gold ornaments
   d) Digital balance   -   Cricket ball

6.Define least count of any device
7.Differentiate mass and weight.

LN-2.MOTION   **TEST-3**

I. Choose the correct answer.   (4×1=4)

1.The centrifugal force is
   a) a real force.   b) the force of reaction of centripetal force.
   c) a virtual force.   d) directed towards the centre of the circular path.

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2.The area under velocity – time graph represents the
   a) velocity of the moving object.   b) displacement covered by the moving object.
   c) speed of the moving object.   d) acceleration of the moving object.
3.Speed is a _____ quantity
   a) scalar   b) vector   c) velocity   d) Acceleration

4.Negative acceleration is called _____
   a) deceleration   b) change of velocity   c) centripetal acceleration   d) displacement
5.Define velocity.   **(3×2=6)**
6.Compare speed and velocity.
7.Explain different types of motion

LN-3.FLUIDS   TEST-4

I. Choose the correct answer.   **(4×1=4)**
1.The size of an air bubble rising up in water
   a) decreases   b) increases   c) remains same   d) may increase or decrease
2.The instrument used to measure atmospheric pressure is_____
   a) Barometer   b) Ammeter   c) Voltmeter   d) Hydrometer
3.A drinking straw works on the existence of_____
    a) Liquid pressure   b) Force   c) Atmospheric pressure   d) Pascal's law
4.Clouds float in atmosphere because of their low
   a) density   b) pressure   c) velocity   d) mass
5.**State whether true or false**   **(3×2=6)**
   a)Archimedes' principle can also be applied to gases.
   b) Hydraulic press is used in the extraction of oil from oil seeds.
6.**Match the following.**
   a) Density   -   hρg
   b) 1 gwt   -   Milk
   c) Pascal's law   -   MassVolume
   d) Pressure exerted by a fluid -   Pressure
   e)Lactometer   -   980 dyne
7.The density of water is 1 g cm^{-3}. What is its density in S.I. units?

LN-10.MATTER AROUND US   TEST-5

I. Choose the correct answer.   **(4×1=4)**
1.Among the following _____ is a mixture
   a) Common Salt   b) Juice   c) Carbon dioxide   d) Pure Silver
2.The separation of denser particles from lighter particles done by rotation at high speed is called
_____
   a) Filtration   b) sedimentation   c) decantation   d) centrifugation
3.Alcohol can be separated from water by _____
    a) Fractional distillation   b) Filtration   c) sedimentation   d) separating funnel
4.Chromatography is based on the principle of _____
    a) Filtration   b) Different solubilities   c) solution   d) sublimation
5.Define Sublimation   **(3×2=6)**
6.Name the components in each of the following mixtures.
i. Ice cream   ii. Lemonade   iii. Air   iv. Soil
7.**State whether true or false**.
   a) Oil and water are immiscible with each other.
   b) Buttermilk is an example of heterogeneous Mixture.

LN-10.MATTER AROUND US   TEST-6

I. Choose the correct answer.   **(4×1=4)**

1.When we mix a drop of ink in water we get a _____
  a) Heterogeneous Mixture   b) Compound   c) Homogeneous Mixture   d) Suspension

2._____ is essential to perform separation by solvent extraction method.
  a) Separating funnel   b) filter paper   b) centrifuge machine   d) sieve

3.In petroleum refining, the method of separation used is _____
  a) Fractional distillation   b) Filtration   c) sedimentation   d) separating funnel

4 . A _____ mixture has no distinguishable boundary between its components.
  a) Compound   b) Hetrogeneous   c) Homogeneous mixture   d) Element

5.**Match the following**.   **(3×2=6)**
  a)Element   -   Settles down on standing
  b) Compound   -   Impure substance
  c)Colloid   -   Made up of molecules
  d) Suspension   -   Pure substance
  e)Mixture   -   Made up of atoms

6.Write the differences between elements and compounds
7.Explain Tyndall effect and Brownian movement with suitable diagram.

LN-11.ATOMIC STRUCTURE   TEST-7

I. Choose the correct answer.   **(4×1=4)**

1. The correct electronic configuration of potassium is _____
  a) 2,8,9   b) 2,8,1   c) 2,8,8,1   d) 2,8,8,3

2. The term nucleons refer to
  a) protons and electrons   b) only neutrons   c) electrons and neutrons   d) protons and neutrons

3. The valency of Argon is _____
  a) 1   b) 2   c) 0   d)   3

4. _____ isotope is used in the nuclear reactors.
  a)   Uranium-235   b) Cobalt – 60   c) Carbon -14   d) Rubidium - 82

5**. Match the following**.   **(3×2=6)**
  a) Dalton   -   1 . Hydrogen atom
                                model

  b) Chadwick   -   2. Discovery of
                                nucleus

  c) Rutherford   -   3. First atomic
                                theory

  d) Neils Bohr   -   4. Plum pudding
                                model

                   -      Discovery of neutrons

6. **State true or false.**
  a) Isotopes of an element have different atomic numbers. .
  b) Electrons have negligible mass and charge.
7. Write the electronic configuration of K and Cl

LN-11.ATOMIC STRUCTURE   TEST-8

I. Choose the correct answer.   (4×1=4)

1. Change in the number of neutrons in an atom changes it to
   a) an ion   . b) an isotope.   c) an isobar.   d) another element.
2. Calcium and Argon are examples of a pair of _____
   a) isobars   b) isotopes   c) isotones   d) noble gas
3. The number of neutrons present in is $^{7}_{3}Li$ _____

   a) 3   b) 7   c) 4   d) 10
4. The number of protons, neutrons and electrons present respectively in $^{80}_{35}Br$ are

   a) 80, 80, 35   b) 35, 55,80   c) 35, 35, 80   d) 35, 45, 35
5. For an atom 'X', K, L and M shells are completely filled. How many electrons will be   **(3×2=6)**
   present in it?
6.   Draw the structure of oxygen and sulphur atoms.
7. **State true or false**
   a) Electrons have negligible mass and charge.
   b) The maximum number of electron in L Shell is 10.

LN-12.PERIODIC CLASSIFICATION   TEST-9
OF ELEMENTS
I. Choose the correct answer.   (4×1=4)

1.Modern periodic law states that the physical and chemical properties of elements are the
   periodic functions of their
   a) atomic numbers   b) atomic masses   c) similarities   d) anomalies
2. Elements in the modern periodic table are arranged in groups and periods.
   a) 7, 18   b) 18, 7   c) 17, 8   d) 8, 17

3. Noble gases belong to group of the periodic table_____.
   a) 14   b) 16   c) 17   d) 18
4. Example for liquid metal is _____.
   a) Au   b) Cu   c) Hg   d) Ag

5. **Match the following**.   (3×2=6)
   a) Triads   -   Newlands

   b) Alkali metal   -   Calcium

   c) Law of octaves   -   Henry
                                       Moseley

   d) Alkaline   earth   -   Sodium
       metal

   e) Modern   -   Dobereiner
       Periodic Law

6. What are groups and periods in the modern periodic table?
7**. State whether true or false**
   a) Metals can gain electrons.
   b) Group 17 elements are named as Halogens.