

UNIT-TEST

9. SOLUTIONS

MARKS : 25

TIME : 45 MIN.

PART-I

Answer any Three questions.

 $3 \times 2 = 6$

1. What is osmotic pressure?
2. NH_3 and HCl do not obey Henry's law. Why?
3. Define the term "isotonic solution".
4. Define the i) molality ii) Normality
5. 50g of tap water contains 20mg of dissolved solids
What is the TDS value in ppm?

PART-II

Answer any Three questions

 $3 \times 3 = 9$

1. What are the limitations of Henry's law?
2. What is Van't Hoff factor 'i'?
3. What are the conditions when a solution tends to behave like an ideal solutions.
4. State and explain Henry's law.
5. Define the term solubility and what are the factors that influences solubility.

PART-III

Answer any TWO questions.

- 1] Write the formula to calculate the molar mass of a solute from relative lowering of vapour pressure values.
- 2] How will you determine the molar mass of solute from elevation of boiling point?
- 3] Define Raoult's law . b)

state Raoult Law and obtain expression for Lowering of vapour pressure when a non-volatile solute is dissolved in solvent?

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