

VIVEKANANDA VIDYALAYA HIGHER SECONDARY SCHYOOOL, ANTHIYUR, ERODE.
12TH CHEMISTRY - 2 & 3 MARKS QUESTIONS VOL - 1 2019 - 2020
Especially For The Slow Learners And Average Students

These Qns
Enough to get
good marks.

1. METALLURGY

1. What are the differences between minerals and ores?
2. What is the role of Limestone in the extraction of Iron from its oxide Fe_2O_3 ?
3. Which type of ores can be concentrated by froth floatation method? Give two examples for such ores.
4. Give the uses of Zinc.
5. Give the basic requirement for vapour phase refining.
6. Define gravity separation OR Hydraulic washing.
7. Write the applications of copper.
8. what is ammonia leaching?
9. CO is more stable than higher temperature. why?
10. Define vapour phase method Zr.
11. Write about magnetic separation.
12. Write about aluminothermic process.
13. Explain cyanide leaching.
14. Give the limitations of ellingham diagram.

2. p - Block Elements - I

1. write a note on fisher tropesch synthesis.
2. Give the uses of Borax
3. CO is a reducing agent. justify.
4. Give the structure of CO and CO_2 ?
5. give the uses of Silicones.
6. Write a note on Zeolites.
7. Complete the reaction
 $HCOOH + H_2SO_4 \rightarrow ?$
 $B_2H_6 + CH_3OH \rightarrow ?$
8. How Will You Identify borate radical?
9. How is borax prepared from Colemanite?
10. Define Allotropism.
11. What is the action of heat on boric acid?
12. Write the preparation of borazole.
13. write about Tecto silicates with example.
14. What is catenation ? property of carbon

3. p - Block Elements - II

1. What is inert pair effect?
2. what happens when PCl_5 is heated ?
3. Chalgogens belongs to p- block. Give Reason.
4. Gives the uses of helium.
5. Gives the uses of sulphuric acid.
6. how does ammonia reacts with $CuSO_4$?
7. Complete the reaction
 $NaCl + MnO_2 + H_2SO_4 \rightarrow$
 $XeO_6^{4-} + Mn^{2+} + H^+ \rightarrow$
8. Discuss the structure of SO_2 .
9. What type of hybridisation occur in BrF_5 and BrF_3
10. Give two equations to illustrate the chemical behavior of phosphine.
11. draw the structure of the following compounds
a) Marshall's acid b) dithionic acid
12. give the uses of argon.
13. how will you prepare chlorine in the laboratory ?
14. Explain why fluorine always exhibit -1 oxidation state ?

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4. Transition and Inner Transition Elements

1. What are transition metals? Give four examples.
2. What are interstitial compounds?
3. Why Gd³⁺ is colorless?
4. What are actinides? Give three examples.
5. Which is more stable? Fe³⁺ or Fe²⁺ - explain.
6. Explain why Cr²⁺ is strongly reducing while Mn³⁺ is strongly oxidizing.
7. Why do zirconium and Hafnium exhibit similar properties?
8. Which is stronger reducing agent Cr²⁺ or Fe²⁺?
9. Write the chromyl chloride test.
10. What is the action of heat on K₂Cr₂O₇?
11. why do transition elements form complex?
12. Write polymerization reaction using Zeigler - Natta catalyst.
13. Transition metals show high melting point why?
14. Which is more stable ? Fe³⁺ or Fe²⁺ Explain.

5.Coordination Chemistry

- 1.Any five IUPAC names in the book back question. Coordination Compounds.
- 2.Any five Coordination Compounds name with formula.
- 3.Write the structural formula for the rosy red precipitate of a complex (dimethyl glyoxime)
4. $[CuCl_4]^{2-}$ exists while $[CuI_4]^{2-}$ does not exist why?
- 5.What is linkage isomerism? explain with example.
- 6.Difference between double salts and coordination compounds?
- 7.Short note on Homoleptic and Heteroleptic complex.
- 8.Note on i) Ionisation isomers ii) Stereo isomers.
- 9.Calculate CFSE of low spin complex having $t_{2g}^5 e_g^0$ electronic configuration.
10. write a note on bridged carbonyls.
11. Explain d-d transition with an example.
12. write any three postulates of werner's theory.
13. Explain hydrate isomers.

6. Solid state

1. Define unit cell.
2. Give any three characteristics of ionic crystals.
- 3.What are point defects?
4. Difference between hexagonal close packing and cubic close packing.
5. What is Schottky defect?
6. why ionic crystals are hard and brittle.
7. Distinguish between tetrahedral and octahedral voids.
8. write a note on frenkel defects.
9. calculate the number of atoms in fcc unit cell.
10. differentiate isotropy from anisotropy
11. Draw the structure of NaCl.
12. Sketch the Sc, bcc and fcc structure.
13. Write the Bragg's equation.
14. Define crystal lattice.
15. What are molecular solids example.
16. Short note on metal excess defect.

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7. Chemical kinetics

1. Define half life of a reaction.
2. define rate law and rate constant.
3. what is meant by zero order reaction?
4. Note on pseudo first order reaction.
5. write Arrhenius equation and explain the terms.
6. what are the factors affecting the reaction rate?
7. what is activation energy E_a ?
8. Calculate the half life period for a zero order reaction.
9. express the rate of the reaction in terms of changes in the concentration of NO, O₂ and NO₂.
10. Give three examples of first order reaction.

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**" HARD WORK BEATS
TALENT WHEN TALENT
DOSEN'T WORK HARD"**