

Exercise 4 (Extra-Edge)

1. Which one of the following is present as an active ingredient in bleaching power for bleaching action?

[CBSE AIPMT 2011]

- (a) CaCl_2 (b) CaOCl_2
(c) **Ca(OCl)_2** (d) CaO_2Cl_2

2. Which of the following compounds has the lowest melting point?

[CBSE AIPMT 2011]

- (a) CaF_2 (b) CaCl_2
(c) CaBr_2 (d) **CaI_2**

3. Property of the alkaline earth metals that increase with their atomic number?

[BSE AIPMT 2010]

- (a) electronegativity.
(b) **solubility of their hydroxides in water.**
(c) solubility of their sulphates in water.
(d) ionisation energy.

4. Which of one following alkaline earth metal sulphates has hydration enthalpy higher than lattice enthalpy?

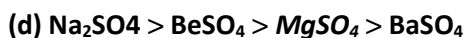
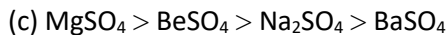
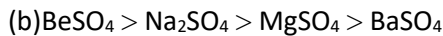
[CBSE AIPMT 2010]

- (a) SrSO_4 (b) CaSO_4
(c) **BeSO_4** (d) BaSO_4

5. The solubilities of Na_2SO_4 , BeSO_4 , MgSO_4 , and BaSO_4 will follow the order

[CBSE AIPMT 2010]

- (a) $\text{BeSO}_4 > \text{MgSO}_4 > \text{Na}_2\text{SO}_4 > \text{BaSO}_4$



6. In aqueous solution, the most stable sulphate is

[CPMT 2010]



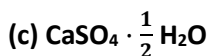
7. The yellow coloured flame is of

[MP PMT 2010]



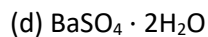
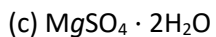
8. Gypsum on heating to 390 K gives

[RPMT 2010]



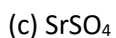
9. Epsom salt is

[RPMT, VMCC 2010]



10. Which of the following alkaline earth metal sulphates is least soluble in water?

[Haryana PET 2010]



11. Assertion Bleaching powder is a mixed salt

Reason In the presence of COCl_2 bleaching powder decomposes to give CaCl_2 and O_2

[AIMS 2009]

(a)Both Assertion and Reason are true and Reason is the correct explanation of Assertion.

(b)Both Assertion and Reason are true but Reason is not the correct explanation of Assertion.

(c)Assertion is true but Reason is false.

(d)Both Assertion and Reason are false.

12. Which one of the following does not react with water even under hot condition?

[RPMT 2009]

(a)Na

(b)Be

(c)Ca

(d)K

(e)Sr

13. Correct order of stability of group II A metal carbonates is

[OJEE 2009]

(a) $\text{MgCO}_3 > \text{CaCO}_3 > \text{SrCO}_3 > \text{BaCO}_3$

(b) $\text{BaCO}_3 > \text{SrCO}_3 > \text{CaCO}_3 > \text{MgCO}_3$

(c) $\text{SrCO}_3 > \text{BaCO}_3 > \text{CaCO}_3 > \text{MgCO}_3$

(d) $\text{CaCO}_3 > \text{MgCO}_3 > \text{BaCO}_3 > \text{SrCO}_3$