SIR CV RAMAN COACHING CENTRE IDAPPADI,SALEM

X SCIENCE UNIT – 7 ,,SLIP TEST

TOTAL MARK : 25 M

DATE : 17.06.2024

CHOOSE THE CORRECT BEST ANSWER (5 X 1= 5M)

1.Which of the following statement is incorrect?

a. 12 gram of C – 12 contains Avogadro's number of atoms.

b. One mole of oxygen gas contains Avogadro's number of molecules.

c. One mole of hydrogen gas contains Avogadro's number of atoms.

d. One mole of electrons stands for 6.023 × 1023 electrons

2. Which of the following is a triatomic molecule?

a. Glucose b. Helium c. Carbon dioxide d. Hydrogen

3. Which of the following has the smallest mass?

a. 6.023×10^{23} atoms of He b. 1 atom of He c. 2 g of He d. 1 mole atoms of He

4. Mass of 1 mole of Nitrogen atom is

a. 28 amu b. 14 amu c. 28 g d. 14 g

5. The volume occupied by 1 mole of a diatomic gas at S.T.P is

a. 11.2 litre b. 5.6 litre c. 22.4 litre d. 44.8 litre

ANSWER ANY FIVE QUESTIONS (5 X 2= 10 M)

6. Define: Relative atomic mass.

7. Define: Atomicity

8. What is Molar volume of a gas?

9. Find the percentage of nitrogen in ammonia.

10. Give any two examples for hetero diatomic molecules.

11. Calculate the % of oxygen in $Al_2(SO_4)_3$. (Atomic mass: Al-27, O-16, S -32)

12. (i)Atomicity of phosphorous is _____

(ii) The number of atoms present in a molecule is called its _____

ANSWER ANY TWO QUESTIONS (2 X 5 = 10 M)

13.Give the salient features of "Modern atomic theory"

14. Derive the relationship between Relative molecular mass and Vapour density

15. How many grams are there in the following?

i. 2 moles of hydrogen molecule, H_2

ii. 3 moles of chlorine molecule, Cl_2

iii. 5 moles of sulphur molecule, S_8

iv. 4 moles of phosphorous molecule, P_4

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