Ex.No.: **Salt name:**

Date:

S. No.	EXPERIMENT	OBSERVATION	INFERENCE
	PRELIMINARY TESTS		
1	**Odour:** Note the Odour of the organic compound.		
2	**Test with litmus paper:** Touch the Moist litmus paper with an organic compound.		
3	**Action with sodium bicarbonate:** Take 2 ml of saturated sodium bi carbonate solution in a test tube. Add 2 or 3 drops (or a pinch of solid) of an organic compound to it.		
4	**Action with Borsche's reagent:** Take a small amount of an organic compound in a test tube. Add 3 ml of Borsche's reagent, 1 ml of Conc HCl to it, then warm the mixture gently and cool it.		
5	**Charring test:** Take a small amount of an organic compound in a dry test tube. Add 2 ml of conc H_2SO_4 to it, and heat the mixture.		
	TESTS FOR ALIPHATIC OR AROMATIC NATURE:		
6	**Ignition test:** Take small amount of the organic compound in a Nickel spatula and burn it in Bunsen flame		
	TESTS FOR AN UNSATURATION:		
7	**Test with bromine water:** Take small amount of the organic compound in a test tube add 2 ml of distilled water to dissolve it. To this solution add few drops of bromine water and shake it well.		
8	**Test with $KMnO_4$ solution:** Take small amount of the organic compound in a test tube add 2 ml of distilled water to dissolve it. To this solution add few drops of very dilute alkaline $KMnO_4$ solution and shake it well.		

S. No.	EXPERIMENT	OBSERVATION	INFERENCE
	TEST FOR SELECTED ORGANIC FUNCTIONAL GROUPS		
	Test for _____		
1			

REPORT:

The given organic compound contains /is

(i) Aromatic / Aliphatic

(ii) Saturated / Unsaturated

(iii) _____ functional group

Ex.No.: **Date:**

Estimation of _____

Aim :

To estimate the amount of _____ dissolved in _____ ml of the given unknown solution volumetrically. For this you are given with a standard solution of _____ of normality _____ N and _____ solution as link solution.

Short procedure:

S. No.	CONTENT	TITRATION-I	TITRATION-II
1	Burette solution	_____ (link)	_____ (link)
2	Pipette solution	20 ml of standard _____	20 ml of unknown _____
3	Acid to be added		
4	Temperature	_____ temperature	_____ temperature
5	Indicator	_____ indicator	_____ indicator
6	End point	Appearance of permanent _____ colour	Appearance of permanent _____ colour
7	Equivalent weight of _____ = _____ g		

Titration –I (Link _____) Vs (Standard _____)

S. No.	Volume of standard _____ (ml)	Burette readings		Concordant value (Volume of _____)(ml)
		Initial (ml)	Final (ml)	
1	20	0		
2	20	0		
3	20	0		

Calculation :

Volume of _____ (link) solution (V_1) = _____ ml

Normality _____ (link) solution (N_1) = _____ N

Volume of standard _____ solution (V_2) = 20 ml

Normality of standard _____ solution (N_2) = _____ N

According to normality equation: $V_1 \times N_1 = V_2 \times N_2$

$$N_1 = \frac{V_2 \times N_2}{V_1} = \frac{20 \times \quad}{\quad} = \_\_\_\_\_ \text{ N}$$

Titration –II (Unknown _____) Vs (Link _____)

S. No.	Volume of unknown _____ (ml)	Burette readings		Concordant value (Volume of _____)(ml)
		Initial (ml)	Final (ml)	
1	20	0		
2	20	0		
3	20	0		

Calculation :

Volume of unknown _____ solution (V_1) = 20 ml

Normality of unknown _____ solution (N_1) = _____ N

Volume of _____ (link) solution (V_2) = _____ ml

Normality _____ (link) solution (N_2) = _____ N

According to normality equation: $V_1 \times N_1 = V_2 \times N_2$

$$N_1 = \frac{V_2 \times N_2}{V_1} = \frac{\qquad\qquad}{20} = \_\_\_\_\_ \text{ N}$$

The normality of unknown _____ solution = _____ N

Weight calculation:

The amount of _____ dissolved in 1 lit of the solution = **(Normality) x (equivalent weight)**

The amount of _____ dissolved

in 750 ml of the solution $= \frac{Normality \times Equivalent\ weight \times \qquad\qquad}{\qquad\qquad}$

$= \frac{\qquad\qquad\qquad\qquad\qquad\qquad\qquad\qquad}{\qquad\qquad\qquad\qquad\qquad\qquad\qquad\qquad}$

= _____ g

Report :

i) The normality of _____ (link) solution = _____ N

ii) The normality of unknown _____ solution = _____ N

iii) The amount of _____ dissolved in ____ ml of the solution = _____ g