COMMON HALF YEARLY EXAMINATION - 2024

*	Standard XII	Reg No
	CHEMISTRY	
Time : 3.00 hrs	Part - A	Marks: 70
How many Faradays of ele	on O ₃ .2H ₂ O c) Al ₂ O ₃ attice unit cell is	d) None of these d) 32% following reaction to occur
MnO ₄ ⁻ →Mn ²⁺ a) 7 F b) 5 I 4. Which of the following reada a) Zn Hg / NaOH b) Zn 5. A complex in which the ox a) K ₄ [Fe(CN) ₆] c) Fe(Co) ₅	gent can be used to convert //NH.Cl c) Sn / HCl	is zero is NH ₃) ₃]
 6. An aqueous solution of bota a) Basic b) Ne 7. Williamson synthesis of pra a) Electrophilic substitution c) SN₂ reaction 	eutral c) Amphoteri reparing dimethyl ether on reaction b) SN, reaction	
9. The transition element whi a) Ni b) Mn 10. Time required for the read	c) Cr	id d) gas in gas ate is d) Sc n one half of its initial value is
called a) Half Life period c) Zero order 11. The aqueous solution of s are respectively	b) First order d) Second or odium formate, anilinium ch	
a) acidic, acidic, acidic c) basic, acidic, basic 12. Assertion: P-N,N-dimeth Reason: The aldehydic a) Both Assertion and Reas b) Both assertion and reas c) Assertion is true, but Re d) Both Assertion and Reas 3. The medicinal value of dru a) Deoxyribose c) Therapeutic index 4. If one strand of the DNA complementary strand wor	c (–CHO) group is meta direction are false son are true eason is false ason are true, but Reason is g is measured in terms of its b) Gold numb d) Equilibrium has the sequence 'ATGC uld be CGAACT c) TCCGAAC	tral, basic ergoes benzoin condensation ecting. s not correct. s per n constant eTTGA' then the sequence of eT d) TACGTACT to aniline

XII Chemistry Part - B II. Answer any 6 questions. (Q.No.24 is compulsory) 16. What is inert pair effect? $6 \times 2 = 12$ 17. Which is more stable? Fe3+ or Fe2+ Why? 18. Give the uses of Borax. Define coordination number. 20. Give examples for the first order reaction. 21. Write the dispersed phase and dispersed medium of butter. 22. Write a note on Electrophoresis. 23. Give the IUPAC names: a) CH₃-C-OH 24. Identify A and B in the following sequence of reaction. CH_3 -Br $\xrightarrow{NaN_3}$ A $\xrightarrow{LiAIH_4}$ B + N_2 III. Answer any 6 questions. (Q.No.33 is compulsory) $6 \times 3 = 18$ 25. What are interhalogen compounds? Give examples. 26. What are the properties of interstitial compounds? 27. Write Arrhenius equation and explain the terms involved. 28. What are the factors that affects electrolytic conductance? 29. State Faraday's I and II law 30. Write any one method of preparation for diethyl ether. 31. How is chloropicrin prepared? 32. State any three advantages of food additives. 33. [Ag(NH₃)₂]^{+ -}Write the following. a) Ligand b) Central metal ion c) IUPAC name Part - D IV. Answer all the questions. $5 \times 5 = 25$ 34. a) Explain zone refining process. (OR) What are catenation? Write any two conditions for catenation. ii) Why HF cannot be stored in glass bottles? 35. a) Write the postulates of Werner's theory. (OR) b) i) How will you prepare bleaching powder? ii) Write a note on Frenkel defect 36. a) What is Lanthanoid contraction and what are the consequences of lanthanoid contraction. (OR) b) Differentiate crystalline solids and amorphous solids. 37. a) Derive an expression of Nernst equation. (OR) b) i) Give any three differences between chemisorption and physisorption. ii) What are the characteristics of catalyst? 38. a) i) Write a note on (1) Carbylamine reaction (2) Gabriel phthalimide synthesis

b) i) What is glycosidic linkage?
ii) Identify A, B, C.

Aromatic amine (A) $\frac{\text{HNO}_2}{\text{NaNO}_2 / \text{Con.HCl}}$ (B) $\frac{\text{C}_6 \text{H}_5 \text{OH}}{\text{273} - 278 \text{ K}}$ (C)