

**CLASS 12 CHEMISTRY - KEY FOR CREATED 1 MARK QUESTIONS**Lesson 1

1. b) Zn
2. d) Malachite
3. c) Zone refining
4. c) Copper (I) sulphide ( $\text{Cu}_2\text{S}$ )
5. c) Froth flotation
6. d) Fractional crystallization
7. b) Roasting
8. a)  $2\text{CuCO}_3 \cdot \text{Cu}(\text{OH})_2$
9. a) ZnS
10. c) Sulphur
11. a)  $\Delta G^\circ = -nFE^\circ$
12. d) Flux
13. c) Sulphide
14. c) Smelting
15. a) Carbon blocks
16. a) Collector
17. c) Leaching
18. a) Gravity separation
19. d) Calcination
20. b) Liqutation
21. b) Zinc oxide
22. b) Chrome steel
23. c) Both (a) and (b)
24. b)  $\text{Na}_2[\text{Zn}(\text{CN})_4]$
25. a) NaCN
26. c) Acid leaching

Lesson 2:

1. c) -1
2. d)  $ns^2 np^6$
3. b) Stable electronic configuration
4. b) Stable electronic configuration
5. a) Fluorine
6. b) Allotropism
7. d) 7
8. c) Borax

9. a) Diborane
10. a) Borax
11. a)  $^{10}\text{B}_5$
12. a)  $\text{H}_3\text{BO}_3$
13. d)  $\text{B}_3\text{N}_3\text{H}_6$
14. a) Boron nitride
15. a) (i) and (iii)
16. b) Hexagonal
17. b) Producer gas
18. b)  $\text{CO} + \text{H}_2$
19. c)  $31^\circ\text{C}$
20. c) Neso silicates
21. b) Beryl
22. a) Zeolite
23. c) Quartz

Lesson 3:

1. b) Fractional distillation
2. a) Liquid  $\text{N}_2$
3. d)  $10^{14}$
4. b)  $107^\circ 28'$
5. c) Pyramidal
6. a) Ostwald's process
7. c) Nitronium ion
8. b) AgBr
9. b) Firearms
10. a) Phosphorescence
11. c) Phosphine
12. d)  $sp^3$
13. c) Calcium carbide + Calcium phosphide
14. c) Bent shape
15. b) 2.5 times
16. b) Lead chamber process
17. a)  $\text{V}_2\text{O}_5$
18. a)  $\text{Cl}_2$
19. a)  $\text{CuCl}_2$
20. b)  $\text{Ca}(\text{OCl})_2$
21. a) HF

22. c) HF
23. d)  $\text{HI} > \text{HBr} > \text{HCl} > \text{HF}$
24. b) T-shape
25. d)  $\text{HClO}_4$
26. b) Distorted octahedron
27. a) He
28. a) 3 parts of conc. HCl, one part of conc.  $\text{HNO}_3$
29. b) Pentagonal bipyramidal
30. a) Square pyramidal
31. b)  $sp^3d^2$

Lesson 4:

1. b) Iron
2. a) Cobalt
3. c)  $[\text{Ar}] 3d^5 4s^1$
4. b)  $[\text{Ar}] 3d^{10} 4s$
5. d)  $[\text{Noble gas}] (n - 1) d^{1-10} ns^2$
6. b) Silver
7. c)  $\text{Zn}^{2+}$
8. a)  $\text{Mn}^{2+}, \text{Fe}^{3+}$
9. c)  $\text{TiCl}_4 + \text{Al}(\text{C}_2\text{H}_5)_3$
10. c)  $\text{TiCl}_4 + \text{Al}(\text{C}_2\text{H}_5)_3$
11. b)  $\text{Cr}_2\text{O}_3$
12. c) +6, +6
13. c)  $\text{CrO}_2\text{Cl}_2$
14. b) Tetrahedral
15. a) Cold dilute alkaline  $\text{KMnO}_4$
16. b) Unsaturated organic compound
17. c) Poor shielding effect of 4f sub-shell
18. d) Zr and Hf
19. a)  $\text{La}(\text{OH})_3$

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20. a) Both (A) and (R) are correct and (R) explains (A)  
 21. d) Zinc  
 22. a) Manganese

Lesson 6:

1. a) Glass
2. d) Naphthalene
3. b) Diamond
4. a) Strong electrostatic attractive forces
5. a) London forces
6. b) Polar molecular solids
7. a)
8. b) 4
9. b) 8
10. a) 52.31%
11. b) 68%
12. a) 12
13. b) 4
14. c) 6
15. c) AgBr
16. a) FeO
17. d) Both (a) and (c)
18. a) 74%
19. c) 6
20. a)  $\rho = nM/a^3N_A$
21. b)  $d = n\lambda/2\sin\theta$
22. c) Non-polar molecular solid

Lesson 5:

1. b) 6, 3
2. c) (iii) only
3. b) CO
4. a) II
5. b) 4
6. d)  $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$
7. b)  $\text{K}[\text{PtCl}_3(\text{C}_2\text{H}_4)]$

8. c) Magnus's green salt
9. b) Potassium hexacyanidoferrate (II)
10. a) Tetrammine dichlorido cobalt (III) chloride
11. c)  $[\text{FeF}_6]^{2-}$
12. a) Linkage isomerism
13. b)  $dsp^2$
14. d) Either (a) or (b)
15. b) Octahedral
16. b) EDTA
17. d) Cis-platin
18. d) Potassium trichloro(ethene)platinate(II)
19. a)  $[\text{CoF}_4]^{2-}$
20. c)  $dsp^2$
21. b) Both A and R are correct but R is the correct explanation of A

Lesson 7:

1. c)  $\text{mol L}^{-1} \text{s}^{-1}$
2. a) First order
3. c)  $\text{s}^{-1}$
4. d) 1
5. b) Acid hydrolysis of an ester
6. d) 999 seconds
7. d) Pressure
8. a) First order
9. c) Second order
10. c) First order reaction
11. c)  $\text{mol L}^{-1} \text{s}^{-1}$
12. a) Fractional
13. c) Evaluating rate constants at two different temperatures

14. b) Increase in the number of activated molecules
15. d) Threshold energy
16. a) Both A and R are correct and R is the correct explanation of A
17. b) Both A and R are correct and R is the correct explanation of A
18. a) Both A and R are correct and R is the correct explanation of A

Lesson 8:

1. d) Either (b) or (c)
2. a)  $\text{H}_2\text{O}$
3. d)  $\text{H}_2\text{O} + \text{Cl}^-$
4. b)  $\text{Cr}^{3+}$
5. c)  $1 \times 10^{-14}$
6. a) 3
7. d)  $\text{NaOH} + \text{NaCl}$
8. a)  $\text{OH}^-$  and  $\text{HSO}_4^-$
9. c)  $1.0 \times 10^9$
10. b) 5.09
11. c) 4
12. d) 6.63
13. b) 4.8
14. d) 12
15. a) pH will increase
16. d)  $10^{-9}$
17. a)  $\text{CH}_3\text{COOH} + \text{CH}_3\text{COONa}$
18. c) Decreases
19. d)  $\text{mol}^2 \text{L}^{-2}$
20. a)  $[\text{Ag}^+]^2 [\text{CrO}_4^{2-}]$
21. c)  $\text{NH}_4\text{OH} + \text{NH}_4\text{Cl}$

Lesson 9:

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- (c) Redox reactions
- (b)  $\Omega$  m
- (a) Ohm metre
- (b) Siemen (or) S
- (d)  $\text{Ohm}^{-1} \text{m}^{-1}$
- (a)  $\text{Sm}^2 \text{g}$  equivalent
- (b) Wheatstone bridge
- (a) Kohlrausch's law
- (Not provided)
- (d) All the above
- (b) Agar-Agar gel +  $\text{Na}_2\text{SO}_4$
- (c) Platinum
- (b) 1.107 Volts
- (b) Zero
- (c) Platinum
- (b) - nFE

Lesson 10:

- (a) 40 kJ/mole
- (b) (iii) only
- (a)  $\text{NH}_3$
- (c) Permutit
- (d) Nickel
- (c) Nickel
- (b) Fe, Mg
- (d) Chromatography
- (b) Decomposition of  $\text{H}_2\text{O}_2$  in the presence of Pt catalyst
- (a)  $\text{As}_2\text{O}_3$
- (c)  $\text{H}_2\text{S}$
- (d)  $\text{H}_2\text{S}$
- (a) Ethanol
- (b) Activation energy
- (d) Zymase
- (a)  $\text{Fe}^0/\text{Pd}^0$
- (c) Fog
- (c) Bredig's arc method
- (d) Bredig's arc method
- (d) Peptisation

- (c) Coagulation
- (b) 1 m $\mu$  to 1  $\mu\text{m}$  diameter
- (c) Plate like
- (a)  $\text{As}_2\text{S}_3$
- (c) Scattering of light
- (c) Electrophoresis
- (c) Electro osmosis
- (d) Tollen's test
- (c) Alum containing  $\text{Al}^{3+}$
- (a) Chromium salt
- (b) Fehling's solution
- (d) The measure of protective power of a lyophilic colloid

Lesson 11:

- (c) Ethane-1,2-diol
- (a) Propan-2-ol
- (b) Ethanol
- (b) 2-methyl-propan-2-ol
- (c) Prop-2-en-1-ol
- (c) Cold dilute alkaline  $\text{KMnO}_4$
- (a) Conc. HCl + Anhydrous  $\text{ZnCl}_2$
- (a) Red
- (b)  $\text{SN}_2$  mechanism
- (a)  $1^\circ < 2^\circ < 3^\circ$
- (b) Swern oxidation
- (a) Esterification
- (d) Oxirane
- (a) 1,4-dioxane
- (a) Methanal
- (a) Nitroglycerine
- (b) Glyceric acid and tartaric acid
- (a)  $1^\circ$  alcohol  $>$   $2^\circ$  alcohol  $>$   $3^\circ$  alcohol
- (c) Pyrogallol

- (b) 1,2-dihydroxybenzene
- (c) Dow's process
- (d) Benzene
- (b) Schotten-Baumann reaction
- (d) 2,4,6-trinitrophenol
- (c) Kolbe's Schmitt reaction
- (c) Reimer-Tiemann reaction
- (b) Salicylaldehyde
- (b) p-Hydroxy azo benzene
- (b) Methanal
- (d) Williamson ether synthesis
- (b) p-Bromoanisole
- (b) Ethoxyethane
- (c) Comparatively inert

Lesson 12:

- (a) Prop-2-enal
- (d) Rosenmund reduction
- (b) Catalyst poison
- (c) Gattermann-Koch reaction
- (b) Benzophenone
- (c) Anhydrous  $\text{AlCl}_3$
- (b) Hexamethylene tetramine
- (a) Urotropine
- (c) Nitration of urotropine
- (c) Propane
- (d) Cannizzaro reaction
- (b) Claisen-Schmidt reaction
- (c) Acetone
- (a) Ethanal
- (d) Benzaldehyde

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16. (b) Ethanal
17. (c) Aldol condensation
18. (a) Glucose
19. (c) Guanine
20. (d) Base-Sugar-Phosphate
21. (b) Adenine and thymine; guanine and cytosine
22. (b) Ribose sugar and uracil
23. (b) A globular protein
24. (a) 8
25. (b) 16

Lesson 13:

1. (c) Dopamine
2. (c) Nitro ethane
3. (a) Position isomerism
4. (c) Nitro benzene
5. (b) m-Dinitro benzene
6. (c)  $\text{CCl}_3\text{NO}_2$
7. (d) N-Methyl ethanamine
8. (b) N-Methyl propan-1-amine
9. (b) Hoffmann degradation reaction
10. (d) Gabriel phthalimide synthesis
11. (b) Sabatier-Mailhe method
12. (a) N-Aralkylamine < Arylamine < Ammonia < Alkyl amine < Aralkyl amine
13. (c) Schotten-Baumann reaction
14. (c) Diazotisation
15. (b) Carbylamine reaction
16. (a) Mustard oil reaction
17. (c) Sandmeyer reaction

18. (b) Gomberg reaction
19. (c) Insecticide
20. (c) Bantz-Schiemann reaction
21. (c) Dimethyl amine
22. (d) All the above
23. (d) Both aliphatic and aromatic 1° amine

Lesson 14:

1. (b) 16 isomers
2. (a) 4
3. (a) Glucose
4. (b)  $112^\circ$ ,  $18.7^\circ$
5. (a) Epimers
6. (a) Sorbitol + Mannitol
7. (a) 4
8. (c) Glycosidic bond
9. (b) Invertase
10. (d) Sucrose
11. (c) Iodine solution
12. (c) Glycine
13. (b) Myoglobin
14. (b) Folic acid
15. (c) Guanine
16. (a) Insulin
17. (d) Estrogen
18. (d) Base, Sugar, Phosphate
19. (c) Adenine
20. (d) Base-Sugar-Phosphate
21. (b) Adenine and thymine; guanine and cytosine
22. (b) Ribose sugar and uracil
23. (b) A globular protein
24. (a) 8
25. (b) 16

Lesson 15:

1. (a) Erythromycin
2. (a) Atenolol
3. (b) Sulphanilamide
4. (a) PABA
5. (d) Antagonists
6. (d) All the above
7. (c) Morphine
8. (a) Tranquilizer
9. (b) Diazepam
10. (a) Valium
11. (d) Antipyretic
12. (d) Opioids
13. (c) Zantac
14. (b) Histamine
15. (b) Cimetidine
16. (c) Antiseptic
17. (a) Antiseptic
18. (b) Dettol
19. (c) Iodoform
20. (b) Hydrogen peroxide
21. (b) Bithional
22. (b) Disinfectant
23. (b) Tincture of iodine
24. (d) Paracetamol
25. (b) Aspirin

**All the best...!**