VIVEKANANDA VIDYALAYA MATRIC. HR. SEC. SCHOOL

REVISION EXAMINATION MCQ(B) 2025 60

12th Standard

Chemistry

Date : 23-02-25

Reg.No. : ☐☐☐☐☐☐

Exam Time : 00:30 Hrs Total Marks : 60

Make a proper circle on correct answer

Multiple Choice Question 60 x 1 = 60

1) Roasting of sulphide ore gives the gas (A).(A) is a colourless gas. Aqueous solution of (A) is acidic. The gas (A) is_____.

(a) CO_2 (b) SO_3 (c) SO_2 (d) H_2S

2) The metal oxide which cannot be reduced to metal by carbon is _____.

(a) PbO (b) Al_2O_3 (c) ZnO (d) FeO

3) Wolframite ore is separated from tinstone by the process of_____.

(a) Smelting (b) Calcination (c) Roasting (d) Electromagnetic separation

4) The incorrect statement among the following is_____.

(a) Nickel is refined by Mond's process

(b) Titanium is refined by Van Arkel's process

(c) Zinc blende is concentrated by froth floatation

(d) In the metallurgy of gold, the metal is leached with dilute sodium chloride solution

5) Which of the following metals has the largest abundance in the earth's crust?

(a) Aluminium (b) Calcium (c) Magnesium (d) Sodium

6) Boric acid is an acid because its molecule_____.

(a) contains replaceable H^+ ion (b) gives up a proton

(c) combines with proton to form water molecule

(d) accepts OH^- from water, releasing proton.

7) The element that does not show catenation among the following p-block elements is _____.

(a) Carbon (b) silicon (c) Lead (d) germanium

8) Which of these is not a monomer for a high molecular mass silicone polymer?

(a) Me_3SiCl (b) $PhSiCl_3$ (c) $MeSiCl_3$ (d) Me_2SiCl_2

9) In which of the following, NH_3 is not used?

(a) Nessler's reagent (b) Reagent for the analysis of IV group basic radical

(c) Reagent for the analysis of III group basic radical (d) Tollen's reagent

10) Solid (A) reacts with strong aqueous NaOH liberating a foul smelling gas(B) which spontaneously burn in air giving smoky rings. A and B are respectively_____.

(a) P_4(red) & PH_3 (b) P_4(white) & PH_3 (c) S_8 & H_2S (d) P_4(white) & H_2S

11) Which one of the following compounds is not formed?

(a) $XeOF_4$ (b) XeO_3 (c) XeF_2 (d) NeF_2

12) Which of the following d block element has half filled penultimate d sub shell as well as half filled valence sub shell?

(a) Cr (b) Pd (c) Pt (d) none of these

13) Which one of the following ions has the same number of unpaired electrons as present in V^{3+}?

(a) Ti^{3+} (b) Fe^{3+} (c) Ni^{2+} (d) Cr^{3+}

14) Permanganate ion changes to _____ in acidic medium.

(a) MnO_4^{2-} (b) Mn^{2+} (c) Mn^{3+} (d) MnO_2

15) Which of the following oxidation states is most common among the lanthanoids?

(a) +4 (b) +2 (c) +5 (d) +3

16) An excess of silver nitrate is added to 100ml of a 0.01M solution of Pentaaquachlorochromium (III)chloride. The number of moles of AgCl precipitated would be _____.

(a) 0.02 (b) 0.002 (c) 0.01 (d) 0.2

17) Oxidation state of Iron and the charge on the ligand NO in $[Fe(H_2O)_5NO]SO_4$ are_____.

(a) +2 and 0 respectively (b) +3 and 0 respectively

(c) +3 and -1 respectively (d) +1 and +1 respectively

18) IUPAC name of the complex $K_3[Al(C_2O_4)_3]$ is _____.

(a) potassium trioxalatoaluminium (III)

(b) possium trioxalatoaluminate (II)

(c) potassium trisoxalatoaluminate (III)

(d) potassium trioxalatoaluminate (III)

19) Which type of isomerism is exhibited by $[Pt(NH_3)_2Cl_2]$?

(a) Coordination isomerism (b) Linkage isomerism

(c) Optical isomerism (d) Geometrical isomerism

20) Graphite and diamond are _____.

(a) Covalent and molecular crystals (b) ionic and covalent crystals

(c) both covalent crystals (d) both molecular crystals

21) Solid CO_2 is an example of _____.

(a) Covalent solid (b) metallic solid (c) molecular solid (d) ionic solid

22) In a solid atom M occupies ccp lattice and $\left(\frac{1}{3}\right)$ of tetrahedral voids are occupied by atom N. Find the formula of solid formed by M and N _____.

(a) MN (b) M_3N (c) MN_3 (d) M_3N_2

23) The vacant space in bcc lattice unit cell is _____.

(a) 48% (b) 23% (c) 32% (d) 26%

24) A zero order reaction X $\longrightarrow$ Product, with an initial concentration 0.02M has a half life of 10 min. if one starts with concentration 0.04M, then the half life is

(a) 10 s (b) 5 min (c) 20 min

(d) cannot be predicted using the given information

25) The decomposition of phosphine (PH_3) on tungsten at low pressure is a first order reaction. It is because the _____.

(a) rate is proportional to the surface coverage

(b) rate is inversely proportional to the surface coverage

(c) rate is independent of the surface coverage

(d) rate of decomposition is slow

26) For a first order reaction, the rate constant is 6.909 min^{-1} the time taken for 75% conversion in minutes is _____.

(a) $\left(\frac{3}{2}\right)\log 2$ (b) $\left(\frac{2}{3}\right)\log 2$ (c) $\left(\frac{3}{2}\right)\log\left(\frac{3}{4}\right)$ (d) $\left(\frac{2}{3}\right)\log\left(\frac{4}{3}\right)$

27) During the decomposition of H_2O_2 to give dioxygen, 48 g O_2 is formed per minute at certain point of time. The rate of formation of water at this point is

(a) 0.75 mol min^{-1} (b) 1.5 mol min^{-1} (c) 2.25 mol min^{-1}

(d) 3.0 mol min^{-1}

28) Following solutions were prepared by mixing different volumes of NAOH of HCL different concentrations

1) 60 mL $\frac{M}{10}$ HCl + 40mL $\frac{M}{10}$ NaOH

2) 55 mL $\frac{M}{10}$ HCl + 45mL $\frac{M}{10}$ NaOH

3) 75 mL $\frac{M}{5}$ HCl + 25mL $\frac{M}{5}$ NaOH

4) 100 mL $\frac{M}{10}$ HCl + 100mL $\frac{M}{10}$ NaOH

pH of which one of them will be equal to 1?

(a) (iv) (b) (i) (c) (ii) (d) (iii)

29) Conjugate base for Bronsted acids H_2O and HF are _____.

(a) OH^- and H_2FH^+, respectively (b) H_3O^+ and F^-, respectively

(c) OH^- and F^-, respectively (d) H_3O^+ and H_2F^+, respectively

30) Which of the following fluro compounds is most likely to behave as a Lewis base?

(a) BF_3 (b) PF_3 (c) CF_4 (d) SiF_4

31) The percentage of pyridine (C_5H_5N) that forms pyridinium ion (C_5H_5NH) in a 0.10M aqueous pyridine solution _____.(K_b for $C_5H_5N = 1.7\times10^{-9}$) is

(a) 0.006% (b) 0.013% (c) 0.77% (d) 1.6%

32) The number of electrons that have a total charge of 9650 coulombs is_____.

(a) 6.22×10^{23} (b) 6.022×10^{24} (c) 6.022×10^{22} (d) 6.022×10^{-34}

33) Faraday constant is defined as_____.

(a) charge carried by 1 electron

(b) charge carried by one mole of electrons

(c) charge required to deposit one mole of substance

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34) Zinc can be coated on iron to produce galvanized iron but the reverse is not possible. It is because _____.

(a) Zinc is lighter than iron (b) Zinc has lower melting point than iron

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35) A certain current liberated 0.504gm of hydrogen in 2 hours. How many grams of copper can be liberated by the same current flowing for the same time in a copper sulphate solution _____.

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36) Which of the following is incorrect for physisorption?

(a) reversible (b) increases with increase in temperature

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37) Fog is colloidal solution of _____.

(a) solid in gas (b) gas in gas (c) liquid in gas (d) gas in liquid

38) The most effective electrolyte for the coagulation of As_2S_3Sol is _____.

(a) NaCl (b) $Ba(NO_3)_2$ (c) $K_3[Fe(CN)_6]$ (d) $Al_2(SO_4)_3$

39) The phenomenon observed when a beam of light is passed through a colloidal solution is_____.

(a) Cataphoresis (b) Electrophoresis (c) Coagulation (d) Tyndall effect

40) Which of the following compounds on reaction with methyl magnesium bromide will give tertiary alcohol.

(a) benzaldehyde (b) propanoic acid (c) methyl propanoate

(d) acetaldehyde

41) Which one of the following is the strongest acid.

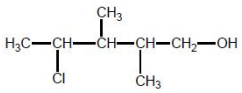
(a) 2 - nitrophenol (b) 4 – chlorophenol (c) 4– nitrophenol

(d) 3 – nitrophenol

42) Carbolic acid is _____.

(a) Phenol (b) Picri cacid (c) benzoic acid (d) phenylacetic acid

43) The correct IUPAC name of the compound



(a) 4 – chloro – 2,3 – dimethyl pentan – 1-ol

(b) 2,3 – dimethyl – 4- chloropentan -1-ol

(c) 2,3,4 – trimethyl – 4- chlorobutan -1-ol

(d) 4– chloro – 2,3,4 – trimethyl pentan – 1-ol

44) Which one of the following reduces tollens reagent

(a) formic acid (b) acetic acid (c) benzophenone (d) none of these

45) Phenyl methanal is reacted with concentrated NaOH to give two products X and Y. X reacts with metallic sodium to liberate hydrogen X and Y are _____.

(a) sodium benzoate and phenol

(b) Sodium benzoate and phenyl methanol

(c) phenyl methanol and sodium benzoate (d) none of these

46) In which of the following reactions new carbon – carbon bond is not formed?

(a) Aldol condensation (b) Friedel craft reaction (c) Kolbe's reaction

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47) Carboxylic acids have higher boiling points than aldehydes, ketones and even alcohols of comparable molecular mass. It is due to their _____.

(a) more extensive association of carboxylic acid via van der Waals force of attraction

(b) formation of carboxylate ion

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48) The product formed by the reaction an aldehyde with a primary amine _____.

(a) carboxylic acid (b) aromatic acid (c) schiff 's base (d) ketone

49) Nitrobenzene on reaction with Con HNO_3 / H_2SO_4 at 80-100^oC forms which one of the following products?

(a) 1,4 – dinitrobenzene (b) 2,4,6 – tirnitrobenzene

(c) 1,2 – dinitrobenzene (d) 1,3 – dinitrobenzene

50) Secondary nitro alkanes react with nitrous acid to form _____.
(a) red solution (b) blue solution (c) green solution (d) yellow solution

51) Which one of the following rotates the plane polarized light towards left?
(a) D(+) Glucose (b) L(+) Glucose (c) D(-) Fructose (d) D(+) Galactose

52) Which one given below is a non-reducing sugar?
(a) Glucose (b) Sucrose (c) maltose (d) Lactose

53) Among the following the achiral amino acid is _____.
(a) 2-ethylalanine (b) 2-methylglycine (c) 2-hydroxymethylserine
(d) Tryptophan

54) The pyrimidine bases present in DNA are _____.
(a) Cytosine and Adenine (b) Cytosine and Guanine
(c) Cytosine and Thiamine (d) Cytosine and Uracil

55) Aspirin is a/an _____.
(a) acetylsalicylic acid (b) benzoyl salicylic acid (c) chlorobenzoic acid
(d) anthranilic acid

56) Nylon is an example of _____.
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57) Non stick cook wares generally have a coating of a polymer, whose monomer is _____.
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58) Regarding cross-linked or network polymers, which of the following statement is incorrect?
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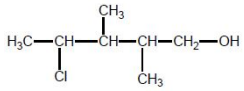

(a) 4 – chloro – 2,3 – dimethyl pentan – 1-ol

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2) Non stick cook wares generally have a coating of a polymer, whose monomer is _____.

(a) ethane (b) prop-2-enenitrile (c) chloroethene

(d) 1,1,2,2-tetrafluoroethane

3) Which one of the following rotates the plane polarized light towards left?

(a) D(+) Glucose (b) L(+) Glucose (c) D(-) Fructose (d) D(+) Galactose

4) Which one of the following reduces tollens reagent

(a) formic acid (b) acetic acid (c) benzophenone (d) none of these

5) The decomposition of phosphine (PH_3) on tungsten at low pressure is a first order reaction. It is because the _____.

(a) rate is proportional to the surface coverage

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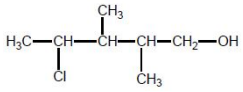

(a) 4 – chloro – 2,3 – dimethyl pentan – 1-ol

(b) 2,3 – dimethyl – 4- chloropentan -1-ol

(c) 2,3,4 – trimethyl – 4- chlorobutan -1-ol

(d) 4– chloro – 2,3,4 – trimethyl pentan – 1-ol

8) Zinc can be coated on iron to produce galvanized iron but the reverse is not possible. It is because _____.

(a) Zinc is lighter than iron (b) Zinc has lower melting point than iron

(c) Zinc has lower negative electrode potential than iron

(d) Zinc has higher negative electrode potential than iron

9) Faraday constant is defined as_____.

(a) charge carried by 1 electron (b) charge carried by one mole of electrons

(c) charge required to deposit one mole of substance

(d) charge carried by 6.22×10^{10} electrons

10) Nitrobenzene on reaction with Con HNO_3 / H_2SO_4 at 80-100°C forms which one of the following products?

(a) 1,4 – dinitrobenzene (b) 2,4,6 – tirnitrobenzene

(c) 1,2 – dinitrobenzene (d) 1,3 – dinitrobenzene

11) Which of the following compounds on reaction with methyl magnesium bromide will give tertiary alcohol.

(a) benzaldehyde (b) propanoic acid (c) methyl propanoate

(d) acetaldehyde

12) Which of the following metals has the largest abundance in the earth's crust?

(a) Aluminium (b) Calcium (c) Magnesium (d) Sodium

13) Carbolic acid is _____.

(a) Phenol (b) Picri cacid (c) benzoic acid (d) phenylacetic acid

14) For a first order reaction, the rate constant is 6.909 min^{-1} the time taken for 75% conversion in minutes is _____.

(a) $\left(\frac{3}{2}\right)\log 2$ (b) $\left(\frac{2}{3}\right)\log 2$ (c) $\left(\frac{3}{2}\right)\log\left(\frac{3}{4}\right)$ (d) $\left(\frac{2}{3}\right)\log\left(\frac{4}{3}\right)$

15) The element that does not show catenation among the following p-block elements is _____.

(a) Carbon (b) silicon (c) Lead (d) germanium

16) The pyrimidine bases present in DNA are _____.

(a) Cytosine and Adenine (b) Cytosine and Guanine

(c) Cytosine and Thiamine (d) Cytosine and Uracil

17) $ZnS + 3O_2 \xrightarrow{\Delta} 2ZnO + 2SO_2 \uparrow$. The above equation is an example for_____.

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22) Permanganate ion changes to _____ in acidic medium.
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(a) Covalent and molecular crystals (b) ionic and covalent crystals
(c) both covalent crystals (d) both molecular crystals

24) The phenomenon observed when a beam of light is passed through a colloidal solution is_____.
(a) Cataphoresis (b) Electrophoresis (c) Coagulation (d) Tyndall effect

25) Conjugate base for Bronsted acids H_2O and HF are _____.
(a) OH^- and H_2FH^+, respectively (b) H_3O^+ and F^-, respectively
(c) OH^- and F^-, respectively (d) H_3O^+ and H_2F^+, respectively

26) Which type of isomerism is exhibited by $[Pt(NH_3)_2Cl_2]$?
(a) Coordination isomerism (b) Linkage isomerism
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(b) Titanium is refined by Van Arkel's process
(c) Zinc blende is concentrated by froth floatation
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32) An excess of silver nitrate is added to 100ml of a 0.01M solution of Pentaaquachlorochromium (III)chloride. The number of moles of AgCl precipitated would be _____.
(a) 0.02 (b) 0.002 (c) 0.01 (d) 0.2

33) The most effective electrolyte for the coagulation of As_2S_3Sol is _____.
(a) NaCl (b) $Ba(NO_3)_2$ (c) $K_3[Fe(CN)_6]$ (d) $Al_2(SO_4)_3$

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(a) carboxylic acid (b) aromatic acid (c) schiff 's base (d) ketone

35) Following solutions were prepared by mixing different volumes of NAOH of HCL different concentrations
1) 60 mL $\frac{M}{10}$ HCl + 40mL $\frac{M}{10}$ NaOH
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3) 75 mL $\frac{M}{5}$ HCl + 25mL $\frac{M}{5}$ NaOH
4) 100 mL $\frac{M}{10}$ HCl + 100mL $\frac{M}{10}$ NaOH
pH of which one of them will be equal to 1?
(a) (iv) (b) (i) (c) (ii) (d) (iii)

36) The number of electrons that have a total charge of 9650 coulombs is_____.
(a) 6.22×10^{23} (b) 6.022×10^{24} (c) 6.022×10^{22} (d) 6.022×10^{-34}

37) Which of the following oxidation states is most common among the lanthanoids?
(a) +4 (b) +2 (c) +5 (d) +3

38) Which is a metalled?
(a) B (b) Be (c) S (d) C

39) Which of the following is incorrect for physisorption?
(a) reversible (b) increases with increase in temperature
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40) Secondary nitro alkanes react with nitrous acid to form _____.
(a) red solution (b) blue solution (c) green solution (d) yellow solution

41) Which one of the following is the strongest acid.
(a) 2 - nitrophenol (b) 4 – chlorophenol (c) 4– nitrophenol
(d) 3 – nitrophenol

42) Roasting of sulphide ore gives the gas (A).(A) is a colourless gas. Aqueous solution of (A) is acidic. The gas (A) is_____.
(a) CO_2 (b) SO_3 (c) SO_2 (d) H_2S

43) Solid (A) reacts with strong aqueous NaOH liberating a foul smelling gas(B) which spontaneously burn in air giving smoky rings. A and B are respectively_____.
(a) P_4(red) & PH_3 (b) P_4(white) & PH_3 (c) S_8 & H_2S (d) P_4(white) & H_2S

44) During the decomposition of H_2O_2 to give dioxygen, 48 g O_2 is formed per minute at certain point of time. The rate of formation of water at this point is
(a) 0.75 mol min^{-1} (b) 1.5 mol min^{-1} (c) 2.25 mol min^{-1}
(d) 3.0 mol min^{-1}

45) The metal oxide which cannot be reduced to metal by carbon is _____.
(a) PbO (b) Al_2O_3 (c) ZnO (d) FeO

46) Fog is colloidal solution of _____.
(a) solid in gas (b) gas in gas (c) liquid in gas (d) gas in liquid

47) Which of these is not a monomer for a high molecular mass silicone polymer?
(a) Me_3SiCl (b) $PhSiCl_3$ (c) $MeSiCl_3$ (d) Me_2SiCl_2

48) IUPAC name of the complex $K_3[Al(C_2O_4)_3]$ is _____.
(a) potassium trioxalatoaluminium (III)
(b) possium trioxalatoaluminate (II)
(c) potassium trisoxalatoaluminate (III)
(d) potassium trioxalatoaluminate (III)

49) In a solid atom M occupies ccp lattice and $\left(\frac{1}{3}\right)$ of tetrahedral voids are occupied by atom N. Find the formula of solid formed by M and N _____.
(a) MN (b) M_3N (c) MN_3 (d) M_3N_2

50) In which of the following, NH_3 is not used?
(a) Nessler's reagent
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51) Oxidation state of Iron and the charge on the ligand NO in $[Fe(H_2O)_5NO]SO_4$ are_____.
(a) +2 and 0 respectively (b) +3 and 0 respectively
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52) Among the following the achiral amino acid is _____.
(a) 2-ethylalanine (b) 2-methylglycine (c) 2-hydroxymethylserine
(d) Tryptophan

53) Which of the following d block element has half filled penultimate d sub shell as well as half filled valence sub shell?
(a) Cr (b) Pd (c) Pt (d) none of these

54) Boric acid is an acid because its molecule_____.
(a) contains replaceable H^+ ion (b) gives up a proton
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55) Phenyl methanal is reacted with concentrated NaOH to give two products X and Y. X reacts with metallic sodium to liberate hydrogen X and Y are _____.

(a) sodium benzoate and phenol

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56) A zero order reaction X ⟶ Product, with an initial concentration 0.02M has a half life of 10 min. if one starts with concentration 0.04M, then the half life is

(a) 10 s (b) 5 min (c) 20 min

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57) Which one of the following ions has the same number of unpaired electrons as present in V^{3+}?

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All the best

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VIVEKANANDA VIDYALAYA MATRIC. HR. SEC. SCHOOL

REVISION EXAMINATION MCQ(B) 2025 60

12th Standard

Chemistry

Date : 23-02-25

Reg.No. : ☐☐☐☐☐☐

Exam Time : 00:30 Hrs Total Marks : 60

Make a proper circle on correct answer

Multiple Choice Question 60 x 1 = 60

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(a) Me_3SiCl (b) $PhSiCl_3$ (c) $MeSiCl_3$ (d) Me_2SiCl_2

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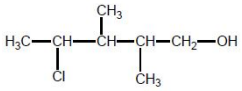

(a) 4 – chloro – 2,3 – dimethyl pentan – 1-ol

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(b) Sodium benzoate and phenyl methanol

(c) phenyl methanol and sodium benzoate (d) none of these

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(a) 1,4 – dinitrobenzene (b) 2,4,6 – tirnitrobenzene

(c) 1,2 – dinitrobenzene (d) 1,3 – dinitrobenzene

All the best

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VIVEKANANDA VIDYALAYA MATRIC. HR. SEC. SCHOOL

REVISION EXAMINATION MCQ(B) 2025 60

12th Standard

Chemistry

Date : 22-02-25

Reg.No. | | | | | | |

Make a proper circle on correct answer

Exam Time : 00:30 Hrs

Total Marks : 60

Multiple Choice Question 60 x 1 = 60

1) (c) SO_2
2) (b) Al_2O_3
3) (d) Electromagnetic separation
4) (d) In the metallurgy of gold, the metal is leached with dilute sodium chloride solution
5) (a) Aluminium
6) (d) accepts OH^- from water, releasing proton.
7) (c) Lead
8) (a) Me_3SiCl
9) (a) Nessler's reagent
10) (b) P_4(white) & PH_3
11) (d) NeF_2
12) (a) Cr
13) (c) Ni^{2+}
14) (b) Mn^{2+}
15) (d) +3
16) (b) 0.002
17) (d) +1 and +1 respectively
18) (d) potassium trioxalatoaluminate (III)
19) (d) Geometrical isomerism
20) (c) both covalent crystals
21) (c) molecular solid
22) (d) M_3N_2
23) (c) 32%
24) (c) 20 min
25) (a) rate is proportional to the surface coverage
26) (b) $\left(\frac{2}{3}\right)\log 2$
27) (d) 3.0 mol min^{-1}
28) (d) (iii)
29) (c) OH^- and F^-, respectively

30) (b) PF_3
31) (b) 0.013%
32) (c) 6.022×10^{22}
33) (b) charge carried by one mole of electrons
34) (d) Zinc has higher negative electrode potential than iron
35) (b) 15.8
36) (b) increases with increase in temperature
37) (c) liquid in gas
38) (d) $Al_2(SO_4)_3$
39) (d) Tyndall effect
40) (c) methyl propanoate
41) (c) 4– nitrophenol
42) (a) Phenol
43) (a) 4 – chloro – 2,3 – dimethyl pentan – 1-ol
44) (a) formic acid
45) (c) phenyl methanol and sodium benzoate
46) (d) Wolf kishner reduction
47) (d) formation of intermolecular H – bonding
48) (c) schiff 's base
49) (d) 1,3 – dinitrobenzene
50) (b) blue solution
51) (c) D(-) Fructose
52) (b) Sucrose
53) (c) 2-hydroxymethylserine
54) (c) Cytosine and Thiamine
55) (a) acetylsalicylic acid
56) (a) polyamide
57) (d) 1,1,2,2-tetrafluoroethane
58) (d) They contain strong covalent bonds in their polymer chain
59) (c) roasting
60) (b) Be