

**SHRI VIDHYABHARATHI MAT. HR.SEC.SCHOOL**

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**PUBLIC EXAMINATION - APRIL - 2025**  
**SSLC - SCIENCE – TENTATIVE ANSWER KEY**

**MARKS : 75**

S.NO	ANSWER	MARK
<b>PART-I</b>		
1.	c) $98 \times 10^4$ dyne	1
2.	d) bifocal lens	1
3.	b) 10V	1
4.	b) Irene curie	1
5.	b) Hg	1
6.	b) increases	1
7.	c) $1 \times 10^{-11}$ M	1
8.	b) combustion of ethanol	1
9.	d) endodermis	1
10.	b) Metacentric	1
11.	d) December 1	1
12.	d) Scratch	1
<b>PART-II</b>		
13.	<b><u>One calorie :</u></b> ➤ One calorie is defined as the amount of heat energy required to rise the temperature of 1 gram of water through 1°C.	2
14.	<b><u>Longitudinal wave:</u></b> ➤ It is a wave in which the particles of the medium vibrating along the direction of propagation of the wave.	2
15.	<b><u>The TWO conditions necessary for rusting of iron are.</u></b> ➤ Presence of <b>oxygen</b> ➤ Presence of <b>water droplets</b> in the atmosphere. ➤ Iron is exposed to moist air	1 1
16.	<b><u>Match it:</u></b> 1. Functional Group –OH – Alcohol 2. Heterocyclic Compounds – Furan 3. Unsaturated Componds – Ethene 4. Soap – Potassium Stearate 5. Carbocyclic Compounds – Benzene	2
17.	<b><u>Importance of valves in the heart:</u></b> The valves are muscular flaps that regulate the flow of blood in a single direction. They prevent back flow of blood.	2

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The above relation is called the combined law of gases. If you consider a gas, which contains  $\mu$  moles of the gas, the number of atoms contained will be equal to  $\mu$  times the Avogadro number,  $N_A$ .

$$\text{i. e. } n = \mu N_A \dots \dots \dots (5)$$

Using equation (5), equation (4) can be written as

$$PV / \mu N_A T = \text{constant}$$

The value of the constant in the above equation is taken to be  $k_B$ , which is called as Boltzmann constant ( $1.38 \times 10^{-23} \text{ J K}^{-1}$ ). Hence, we have the following equation:

$$PV / \mu N_A T = k_B$$

$$PV = \mu N_A k_B T$$

$$\text{Here, } \mu N_A k_B = R,$$

which is termed as universal gas constant whose value is  $8.31 \text{ J mol}^{-1} \text{ K}^{-1}$ .

$$PV = RT.$$

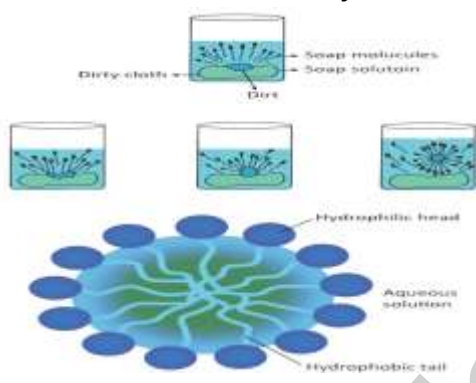
Ideal gas equation is also called as equation of state because it gives the relation between the state variables and it is used to describe the state of any gas.

24.	<b>Myopia</b>	<b>Hypermetropia</b>	
	➤ Myopia, also known as <b>short sightedness</b> .	Hypermetropia, also known as <b>long sightedness</b> .	1/2
	➤ It occurs due to the <b>lengthening</b> of eye ball.	It occurs due to the <b>shortening</b> of eye ball.	1/2
	➤ With this defect, <b>nearby</b> objects can be <b>seen clearly</b>	Distant objects can be seen clearly.	1/2
	➤ <b>Distant objects</b> cannot be seen clearly.	<b>Nearby objects</b> cannot be seen clearly	1/2
	➤ The <b>focal length</b> of eye lens is <b>reduced</b> or the distance between eye lens and <b>retina increases</b> .	The <b>focal length</b> of eye lens is <b>increased</b> or the distance between eye lens and <b>retina decreases</b> .	1/2
	➤ In <b>far point</b> will not be <b>infinity</b> for such eyes and <b>far point</b> has come <b>closer</b> .	The <b>near point</b> will not be at <b>25cm</b> for such eyes and the near point has moved farther.	1/2
	➤ Due to this, the <b>image</b> of distant objects are formed <b>before the retina</b>	Due to this, the <b>image</b> of nearby objects are formed <b>behind the retina</b>	1/2
	➤ This defect can be <b>corrected</b> using a <b>concave lens</b> . The focal length of the concave lens to be used	This defect can be <b>corrected</b> using a <b>convex lens</b> . The focal length of the convex lens to be used is computed	1/2

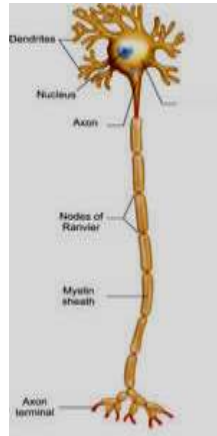
25. **The properties of alpha, beta and gamma radiations:**

Properties	$\alpha$ rays	$\beta$ rays	$\gamma$ rays
What are they?	<b>Helium nucleus</b> ( ${}_2\text{He}^4$ ) consisting of two protons and two neutrons.	They are <b>electrons</b> ( ${}_{-1}\text{e}^0$ ), basic elementary particle in all atoms.	They are <b>electromagnetic waves</b> consisting of photons.
Charge	<b>Positively</b> charged particles. Charge of each alpha particle = <b>+2e</b>	<b>Negatively</b> charged particles. Charge of each beta particle = <b>-e</b>	<b>Neutral</b> particles. Charge of each gamma particle = <b>zero</b>

	Ionising power	<b>100 time</b> greater than $\beta$ rays and <b>10,000</b> times greater than $\gamma$ rays	Comparatively <b>low</b>	<b>Very less</b> ionization power	
	Penetrating power	<b>Low</b> penetrating power (even stopped by a thick paper)	Penetrating power is <b>greater</b> than that of $\alpha$ rays. They can penetrate through a thin metal foil.	They have a <b>very high</b> penetrating power greater than that of $\beta$ rays. They can penetrate through thick metal blocks.	
	Effect of electric and magnetic field	Deflected by <b>both</b> the fields. (in accordance with Fleming's left hand rule)	Deflected by both the fields; but the direction of deflection is <b>opposite</b> to that for alpha rays. (in accordance with Fleming's left hand rule)	They are <b>not deflected</b> by both the fields.	
	Speed	Their speed ranges from <b>1/10</b> to <b>1/20</b> times the speed of light.	Their speed can go up to <b>9/10</b> times the speed of light.	They travel with the <b>speed of light</b> .	
26.	<b><u>The applications of Avogadro's law:</u></b> <ul style="list-style-type: none"> <li>➤ It explains Gay-Lussac's law.</li> <li>➤ It helps in the determination of atomicity of gases.</li> <li>➤ Molecular formula of gases can be derived using Avogadro's law</li> <li>➤ It determines the relation between molecular mass and vapour density.</li> <li>➤ It helps to determine gram molar volume of all gases (i.e, 22.4 litre at S.T.P)</li> </ul>				4
27.	<b>i) Alloy:</b> An alloy is a homogeneous mixture of two or more metals or of one or more metals with certain non-metallic elements.				2
	<b>ii) Reasons for alloying:</b> <ul style="list-style-type: none"> <li>➤ To modify appearance and colour</li> <li>➤ To modify chemical activity.</li> <li>➤ To lower the melting point.</li> <li>➤ To increase hardness and tensile strength.</li> <li>➤ To increase resistance to electricity.</li> </ul>				2
28.	<b><u>The mechanism of cleansing action of soap:</u></b> <ul style="list-style-type: none"> <li>❖ A soap molecule contains two chemically distinct (hydrophilic &amp; hydrophobic) parts that interact differently with water.</li> </ul> <b>Hydrophilic :</b> <ul style="list-style-type: none"> <li>❖ It has one <b>polar end</b>, which is a <b>short head</b> with a carboxylate group (<math>-\text{COONa}</math>) and one <b>non-polar end</b> having the <b>long tail made of the hydrocarbon chain</b>.</li> <li>❖ The polar end is <b>hydrophilic (Water loving)</b> in nature and this end is attracted towards water.</li> </ul>				1

	<p><b>Hydrophobic :</b></p> <ul style="list-style-type: none"> <li>❖ The non-polar end is <b>hydrophobic (Water hating)</b> in nature and it is attracted towards dirt or oil on the cloth, but not attracted towards water.</li> <li>❖ Thus, the hydrophobic part of the soap molecule traps the dirt and the hydrophilic part makes the entire molecule soluble in water.</li> </ul> <p><b>Mechanism :</b></p> <ul style="list-style-type: none"> <li>❖ When a soap or detergent is dissolved in water, the molecules join together as clusters called '<b>micelles</b>'.</li> <li>❖ Their long hydrocarbon chains attach themselves to the oil and dirt.</li> <li>❖ The dirt is thus surrounded by the non-polar end of the soap molecules.</li> <li>❖ The charged carboxylate end of the soap molecules makes the micelles soluble in water.</li> <li>❖ Thus, the dirt is washed away with the soap.</li> </ul> 	1
		2
29.	<p>Locomotion in leech takes place by (i) looping or crawling movement (ii) Swimming movement.</p> <p><b>(i) Looping or Crawling movement</b></p> <p>This type of movement is brought about by the contraction and relaxation of muscles.</p> <p>The two suckers serve for attachment during movement on a substratum.</p> <p><b>(ii) Swimming movement</b></p> <p>Leeches swim very actively and perform undulating movements in water.</p>	2
30.	<p><b><u>Explain the structure of neuron :</u></b></p> <p>A neuron typically consists of three basic parts: Cyton, Dendrites and Axon.</p> <p><b>(i) Cyton:</b></p> <ul style="list-style-type: none"> <li>❖ Cyton is also called cell body or perikaryon.</li> <li>❖ It has a central nucleus with abundant cytoplasm called neuroplasm.</li> <li>❖ The cytoplasm has large granular body called Nissl's granules and the other cell organelles like mitochondria, ribosomes, lysosomes, and endoplasmic reticulum.</li> <li>❖ Neurons do not have the ability to divide. Several neurofibrils are present in the cytoplasm that help in transmission of nerve impulses to and from the cell body.</li> </ul> <p><b>(ii) Dendrites:</b></p> <ul style="list-style-type: none"> <li>❖ These are the numerous branched cytoplasmic processes that project from the surface of the cell body.</li> <li>❖ They conduct nerve impulses towards the cyton. The branched projections increase the surface area for receiving the signals from other nerve cells.</li> </ul> <p><b>(iii) Axon:</b></p> <ul style="list-style-type: none"> <li>❖ The axon is a single, elongated, slender projection.</li> <li>❖ The end of axon terminates as knob like swellings called synaptic knob</li> <li>❖ Axolemma: The plasma membrane of axon</li> <li>❖ Axoplasm: The cytoplasm of axon</li> <li>❖ It carries impulses away from the cyton</li> </ul> <p>Myelin sheath:</p>	1
		1
		1

- ❖ The protective sheath of axon
- ❖ It is covered by a layer of Schwann cells called neurilemma
- ❖ It breaks at intervals by depressions called nodes of ranvier.
- ❖ It acts an insulator and ensures rapid transmission of nerves impulses



1

31.

**Soil erosion causes: (any-4)**

- ❖ The top layers of soil contain humus and mineral salts, which are vital for the growth of plants.
- ❖ Removal of upper layer of soil by wind and water is called soil erosion.
- ❖ Soil erosion causes a significant loss of humus, nutrients and decrease the fertility of soil.
- ❖ The direct and primary effect of soil erosion is soil loss and nutrient leaching resulting in reduction of land productivity.
- ❖ Annual floods cause damages to crops property and lives.
- ❖ Deforested rain forest soil becomes dry and nutrient- deficient as there is no longer vegetation to hold water and nutrient in place
- ❖ Desertification is another possible consequence of erosion.
- ❖ Signification loss of humus and nutrients
- ❖ It decreases the fertility of the soil
- ❖ Vegetation, cover cannot be retained
- ❖ This will affect ground water levels
- ❖ Heavy rains further erode soil and saturate waterways with excess nutrients disrupting the food chains of tropical ecosystems.
- ❖ Eroded sediments can even change the course of rivers which suffer from huge deposits of silt from deforestation.

**ii) Prevent soil erosion: (any-2)**

- ❖ Retain vegetation cover, so that soil is not exposed.
- ❖ Cattle grazing should be controlled.
- ❖ Crop rotation and soil management improve soil organic matter.
- ❖ Runoff water should be stored in the catchment.
- ❖ Reforestation, terracing and contour ploughing.
- ❖ Wind speed can be controlled by planting trees in form of a shelter belt.

2

2



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(OR)

i) **Electric current:**

Electric current is often termed as 'current' and it is represented by the symbol 'I'.

It is defined as the rate of flow of charges in a conductor.

$$I = \frac{Q}{t}$$

ii) **Name and define its unit:**

- SI Unit of electric current is ampere.
- The current flowing through a conductor is said to be one ampere, when a charge of one coulomb flows across any cross-section of a conductor, in one second.

$$\bullet \quad 1 \text{ ampere} = \frac{1 \text{ coulomb}}{1 \text{ second}}$$

iii) **Instrument:** Ammeter.

It is used to measure electric current.

It should be connected in **series** in a circuit

34. a)

i)

Hygroscopic substances	Deliquescence substances
When exposed to the atmosphere at ordinary temperature, they <b>absorb moisture</b> and <b>do not dissolve</b> .	When exposed to the atmospheric air at ordinary temperature, they <b>absorb moisture</b> and <b>dissolve</b> .
Hygroscopic substances <b>do not change</b> its physical state on exposure to air.	Deliquescent substances change its physical state on exposure to air.
Hygroscopic substances may be <b>amorphous solids</b> or liquids.	Deliquescent substances are <b>crystalline solids</b> .
They may be used as <b>drying agents</b> .	They dissolve in water forming <b>saturated solutions</b> .
E.g: quick lime , silica gel	E.g: Caustic soda , caustic potash

ii) **Aquatic animals live more in cold regions because,**

- ❖ More amount of dissolved oxygen is present in the water of cold regions.
- ❖ The solubility of oxygen in water is more at low temperatures.

iii) **Volume percentage:**

The percentage by volume of solute (in ml) present in the given volume of the solution.

$$\text{Volume percentage} = \frac{\text{volume of the solute}}{\text{volume of the solution}} \times 100$$

$$\text{Volume percentage} = \frac{\text{volume of the solute}}{\text{volume of the solute} + \text{volume of the solvent}} \times 100$$

b)

(OR)

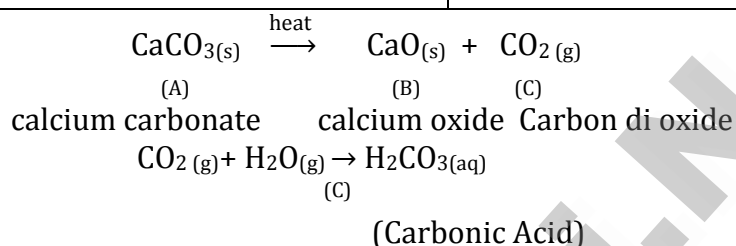


i)

REVERSIBLE REACTION	IRREVERSIBLE REACTION
It <b>can</b> be <b>reversed</b> under suitable condition.	It <b>cannot</b> be <b>reversed</b> .
Both <b>forward and backward</b> reactions take place simultaneously.	It is <b>unidirectional</b> . It proceeds only in forward direction.
It attains equilibrium	Equilibrium is <b>not</b> attained.
The reactants <b>cannot be converted</b> completely into products	The reactants <b>can</b> be completely <b>converted</b> into products
It is relatively <b>slow</b> .	It is <b>fast</b>

3

ii)



1

1

2

Compound	Molecular formula	Name
A	CaCO <sub>3</sub>	calcium carbonate
B	CaO	calcium oxide
C	CO <sub>2</sub>	carbon di oxide

35.

i)

S. N	Tissues	Monocot Root	Dicot Root
1	Number of Xylem	Polyarch	Tetrarch
2	Cambium	Absent	Present (During secondary growth only)
3	Secondary Growth	Absent	Present
4	Pith	Present	Absent
5.	conjunctive tissue	Made up of sclerenchyma	Made up of parenchyma
6.	Eg	Maize	Bean

4

ii) Aerobic and Anaerobic respirationAerobic respiration

- ❖ Aerobic respiration is a type of cellular respiration
- ❖ In which organic food is completely oxidized with the help of oxygen into carbon dioxide, water and energy. It occurs in most plants and animals.
- ❖ Lot of energy is produced
- ❖  $\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O} + \text{ATP}$

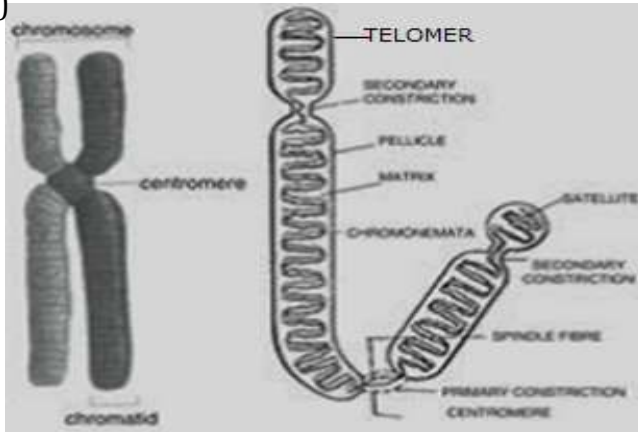
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Anaerobic respiration

- ❖ Anaerobic respiration takes place without oxygen.
- ❖ Glucose is converted into ethanol (Ethanol fermentation by yeast) or lactic acid (lactic acid fermentation by bacteria).
- ❖ Very small quantity of energy is produced  
 $\text{C}_6\text{H}_{12}\text{O}_6 \rightarrow 2\text{CO}_2 + 2\text{C}_2\text{H}_5\text{OH} + \text{Energy (ATP)}$

(OR)

i)



2

- ❖ The chromosomes are thin, long and thread like structures consisting of two identical strands called sister chromatids. They are held together by the centromere.
- ❖ Each **chromatid** is made up of spirally coiled thin structure called **chromonema**.
- ❖ The chromonema has number of bead-like structures along its length which are called **chromomeres**.
- ❖ The chromosomes are made up of DNA, RNA, chromosomal proteins (histones and non-histones) and certain metallic ions.
- ❖ These proteins provide structural support to the chromosome.

**A chromosome consists of the following regions:**

**Primary constriction:** The two arms of a chromosome meet at a point called **primary constriction** or **centromere**.

- ❖ The centromere is the region where spindle fibres attach to the chromosomes during cell division.

3

**Secondary constriction:**

- ❖ Some chromosomes possess secondary constriction **at any point** of the chromosome. They are known as the nuclear zone or **nucleolar organizer** (formation of nucleolus in the nucleus).

**Telomere:**

- ❖ The **end of the chromosome** is called telomere.
- ❖ Each extremity of the chromosome has a polarity and prevents it from joining the adjacent chromosome.
- ❖ It maintains and provides **stability to the chromosomes**.

**Satellite:**

- ❖ Some of the chromosomes have an elongated **knob-like appendage** at one end of the chromosome known as satellite.
- ❖ The chromosomes with satellites are called as the **sat-chromosomes**.

ii)

S.n	Somatic gene therapy	Germ line gene therapy
1.	It is the replacement of defective gene in somatic cell	It is the replacement of defective gene in germ cell (egg and sperm )
2.	It is only beneficial to the patient but not corrected gene may not be carried to the next generation	The gene can be carried to the next generation but till date targeted only somatic cell gene therapy has been targeted
3.	It is take place in body cell	It is take place in egg and sperm cell

2

## ANALYSIS

PART	BOOK BACK		BOOK INSIDE	
	NUMBER OF QUESTIONS	MARKS	NUMBER OF QUESTIONS	MARKS
PART-I	10	10	2	2
PART-II	9	18	1	2
PART-III	7	28	3	12
PART-IV	6	42	-	-

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## மாநிலத்தில் முதலிடம்



2023-2024 ஆம் கல்வியாண்டில் தமிழ் வழியில் பயின்ற மாணவி மாநில அளவில் முதலிடம் பிடித்து சாதனை

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**ZOOLOGY** - 180/180  
**CHEMISTRY** - 162/180  
**PHYSICS** - 147/180



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